



Cambridge IGCSE™

CO-ORDINATED SCIENCES

0654/12

Paper 1 Multiple Choice (Core)

February/March 2022

45 minutes

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A, B, C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.



2

1 Which processes occur in **both** animals **and** plants?

	excretion	movement	respiration
A	✓	✓	✓
B	✓	✓	x
C	✓	x	✓
D	x	✓	✓

2 Which statements about osmosis are correct?

- 1 Osmosis requires a membrane.
- 2 Water can move out of cells by osmosis.
- 3 Water can move into cells by osmosis.

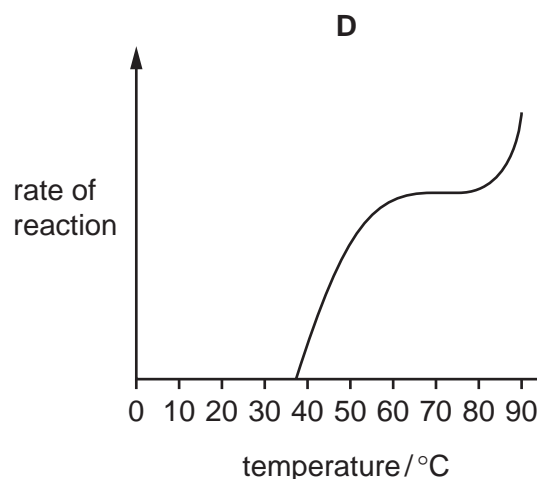
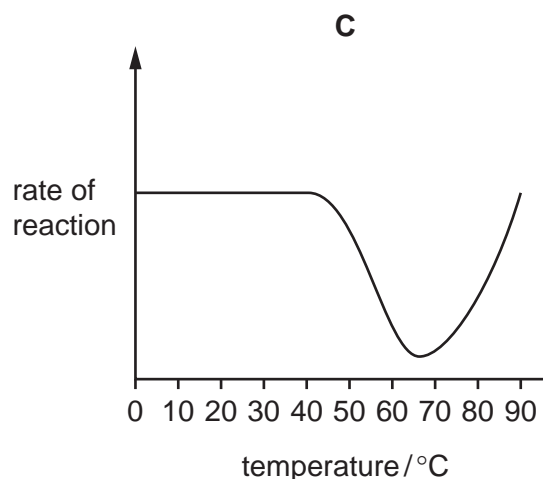
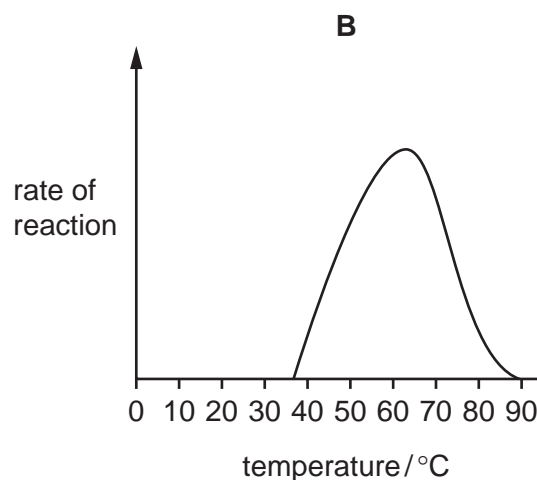
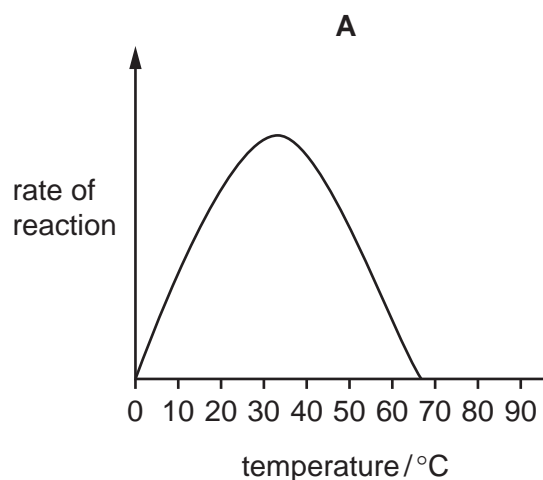
A 1 and 2 only **B** 1 and 3 only **C** 1, 2 and 3 **D** 2 and 3 only

3 Which chemical element is found in all proteins, but **not** in all carbohydrates or fats?

- A** carbon
- B** hydrogen
- C** oxygen
- D** nitrogen

- 4 The Pompeii worm lives in deep-sea hydrothermal vents where **average** temperatures are often as high as 68 °C.

Which graph represents the activity of enzymes found in the Pompeii worm?



- 5 Which **ion** is important for chlorophyll production in plants?

- A calcium
- B iron
- C magnesium
- D nitrate

- 6 What is assimilation?

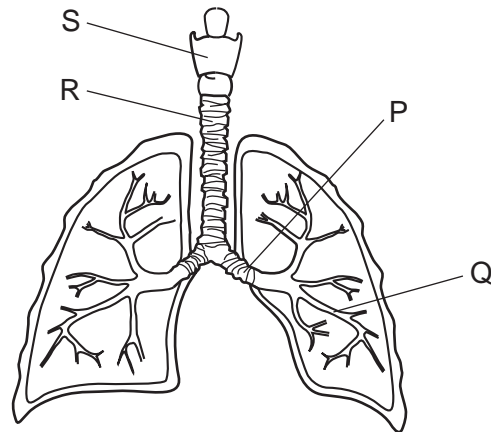
- A the movement of digested food molecules into the cells of the body where they are used, becoming part of the cells
- B the movement of digested food molecules through the wall of the intestine into the blood
- C the passing out of food that has not been digested, as faeces, through the anus
- D the taking of food and drink into the body through the mouth

4

7 In which conditions will the rate of transpiration be greatest?

- A 10°C and high humidity
- B 10°C and low humidity
- C 30°C and high humidity
- D 30°C and low humidity

8 The diagram shows the main structures in the breathing system of humans.

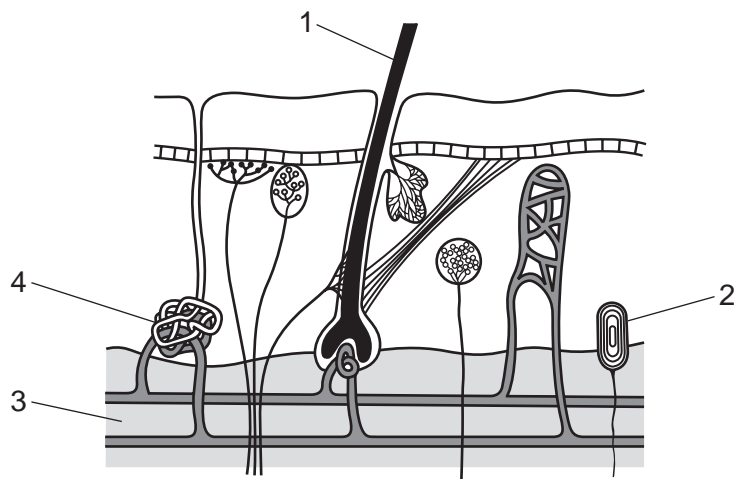


Which row identifies the larynx, bronchus, trachea and bronchioles?

	larynx	bronchus	trachea	bronchioles
A	P	Q	R	S
B	R	P	S	Q
C	S	P	R	Q
D	S	Q	P	R

5

9 The diagram shows a section through the skin.



Which labelled structures help to maintain body temperature in the cold?

- A** 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 3 and 4

10 During the menstrual cycle, an egg is released at ovulation.

The egg passes out of the body if it is not fertilised.

What is the correct order of structures through which the egg passes?

- A** cervix → oviduct → uterus → vagina
B oviduct → uterus → cervix → vagina
C oviduct → vagina → cervix → uterus
D uterus → oviduct → vagina → cervix

11 Which statement is correct?

- A** An allele is a version of a gene.
B DNA is only found in gametes.
C A gene is a length of DNA that codes for fats.
D Cells of human males contain two X chromosomes.

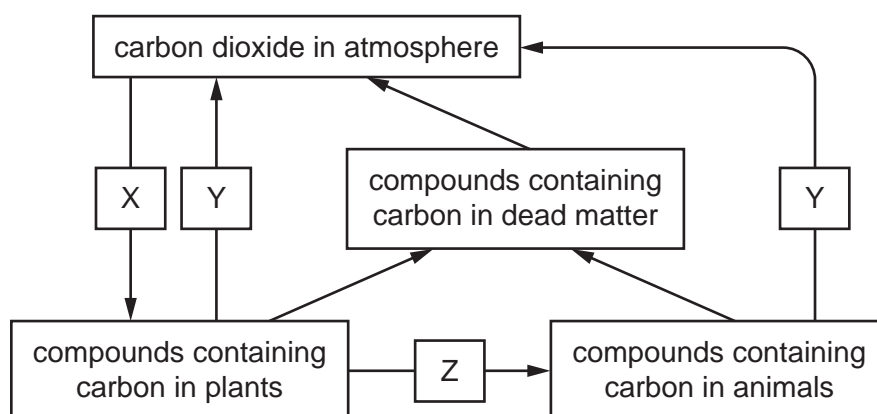
12 The diagram shows a food chain.

beech tree → insect → shrew → owl

Which statement is correct?

- A The beech tree is a consumer.
- B The insect is a producer.
- C The owl is a carnivore.
- D The shrew is a herbivore.

13 The diagram shows part of the carbon cycle.



What are processes X, Y and Z?

	X	Y	Z
A	decomposition	respiration	feeding
B	photosynthesis	respiration	feeding
C	photosynthesis	decomposition	respiration
D	decomposition	photosynthesis	respiration

14 Substance P is separated into different parts using simple physical techniques.

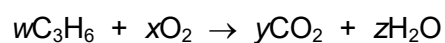
Substance Q is only separated into simpler parts using chemical processes.

Substance R is not separated into simpler parts by either physical or chemical processes.

Which type of substance are P, Q and R?

	P	Q	R
A	compound	mixture	element
B	element	compound	mixture
C	mixture	element	compound
D	mixture	compound	element

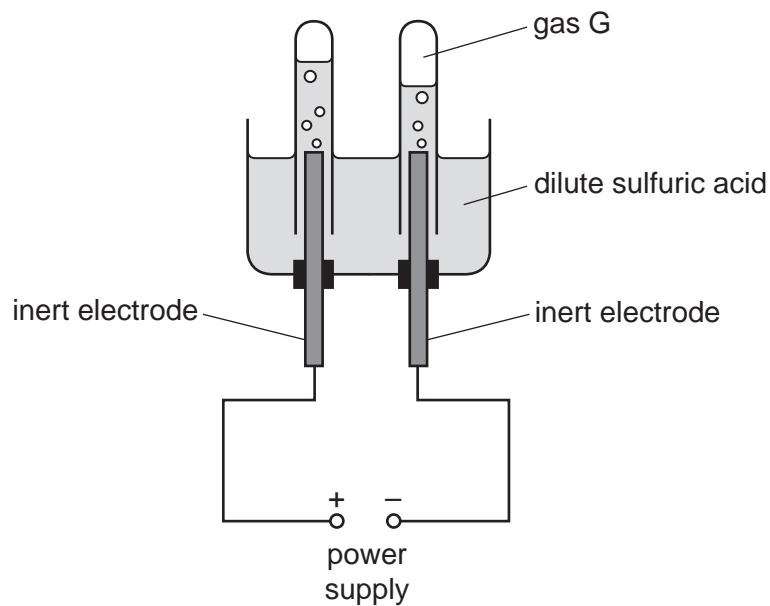
15 Propene, C_3H_6 , burns in excess oxygen to form carbon dioxide and water.



Which values of w , x , y and z balance this equation?

	w	x	y	z
A	1	9	3	3
B	1	5	3	6
C	2	9	6	6
D	2	5	6	3

16 The diagram shows the electrolysis of dilute sulfuric acid.



Gas G ignites with a 'pop' when it is tested with a lighted splint.

What is gas G and at which electrode is it formed?

	gas G	electrode
A	hydrogen	anode
B	hydrogen	cathode
C	oxygen	anode
D	oxygen	cathode

17 Ammonium nitrate is dissolved in a beaker of water.

The temperature of the water decreases by 5 °C.

Which type of reaction occurs?

- A** endothermic
- B** exothermic
- C** oxidation
- D** reduction

18 Which reaction is **not** a redox reaction?

- A iron oxide + carbon \rightarrow iron + carbon dioxide
- B silver nitrate + sodium chloride \rightarrow silver chloride + sodium nitrate
- C copper oxide + hydrogen \rightarrow copper + water
- D magnesium + oxygen \rightarrow magnesium oxide

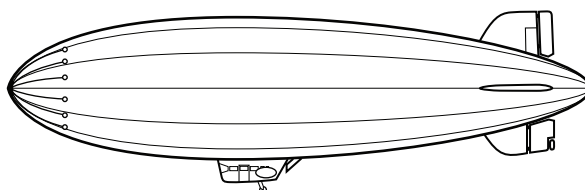
19 Which test and its result identifies aqueous bromide ions?

- A adding acidified aqueous silver nitrate forming a cream precipitate
- B adding acidified aqueous silver nitrate forming a white precipitate
- C adding aluminium foil and heating with sodium hydroxide forming a gas that turns red litmus paper blue
- D adding dilute acid forming a gas that produces a white precipitate when bubbled through limewater

20 Which row about the trends in the elements going down Group I of the Periodic Table is correct?

	reactivity	melting point
A	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

21 An airship containing an unreactive gas floats in air, as shown.



Which gas is used to fill the airship?

- A carbon dioxide
- B helium
- C hydrogen
- D nitrogen

22 Duralumin and magnalium are alloys used in the manufacture of aircraft.

They both contain aluminium and another metallic element.

The alloys are made up of1..... of each element.

They are used because they are2..... than the pure metals.

Which words complete gaps 1 and 2?

	1	2
A	atoms	harder
B	atoms	softer
C	molecules	harder
D	molecules	softer

23 Which metal is extracted from its ore by heating with carbon?

- A** copper
- B** magnesium
- C** potassium
- D** sodium

24 Anhydrous cobalt(II) chloride changes colour when water is added.

Which row shows the colour before and after water is added?

	before	after
A	blue	pink
B	blue	white
C	white	blue
D	white	pink

25 Which substances neutralise acids?

- 1 lime
- 2 limestone
- 3 calcium hydroxide

- A** 1 and 2 only
- B** 1 and 3 only
- C** 2 and 3 only
- D** 1, 2 and 3

26 Butane is a hydrocarbon.

What is the word equation for the complete combustion of butane?

- A butane + oxygen \rightarrow carbon + water
- B butane + oxygen \rightarrow carbon dioxide + carbon monoxide + water
- C butane + oxygen \rightarrow carbon dioxide + water
- D butane + oxygen \rightarrow carbon monoxide + water

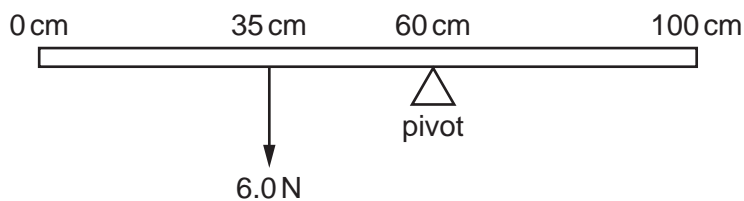
27 Which statement about poly(ethene) is correct?

- A It is an alkene.
- B It is formed in a reaction called cracking.
- C It is obtained from the bitumen fraction of petroleum.
- D It is made by an addition reaction.

28 Which expression is the definition of density?

- A $\frac{\text{mass}}{\text{volume}}$
- B $\frac{\text{volume}}{\text{mass}}$
- C area \times mass
- D mass \times volume

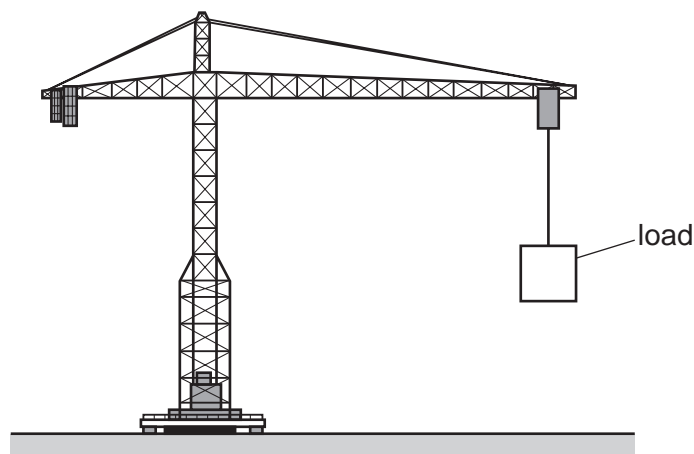
29 The diagram shows a metre rule with a pivot at the 60 cm mark. A force of 6.0 N is applied at the 35 cm mark in the direction shown.



What is the moment of the 6.0 N force about the pivot?

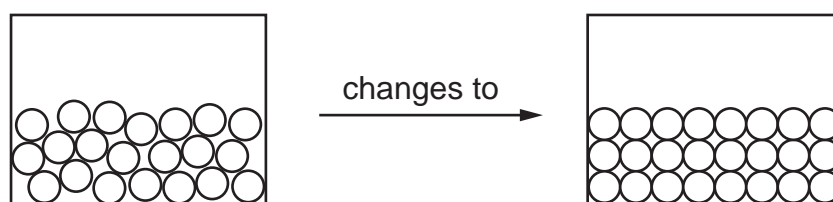
- A 150 N cm
- B 210 N cm
- C 360 N cm
- D 390 N cm

30 A crane does work on a load by lifting it vertically upwards.



Which action decreases the work done on the load?

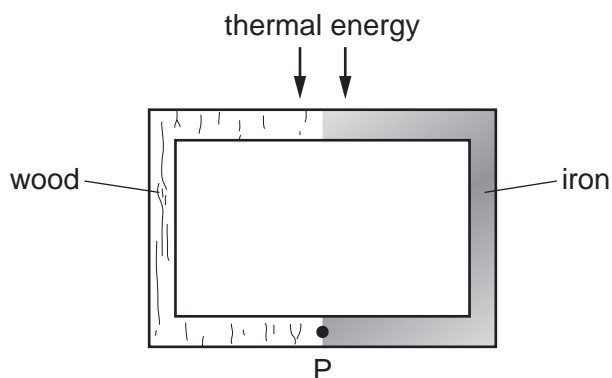
- A lifting the load higher
 - B lifting the load more slowly
 - C reducing the mass of the load
 - D using a more powerful crane
- 31 Which source of energy is non-renewable?
- A hydroelectric
 - B nuclear fission
 - C tides
 - D waves
- 32 The diagram shows the change in the arrangement of the atoms in a substance that is changing state.



What is the change of state?

- A boiling
- B condensation
- C melting
- D solidification

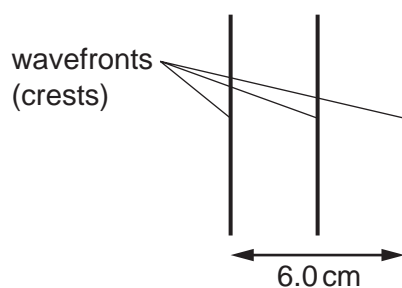
- 33 The diagram shows an object made partly of wood and partly of iron. Thermal energy is supplied in the position shown. Point P is marked at the bottom of the object.



How does most thermal energy reach point P?

- A by conduction through the iron
 - B by conduction through the wood
 - C by convection through the iron
 - D by convection through the wood
- 34 The diagram shows a water wave seen from above.

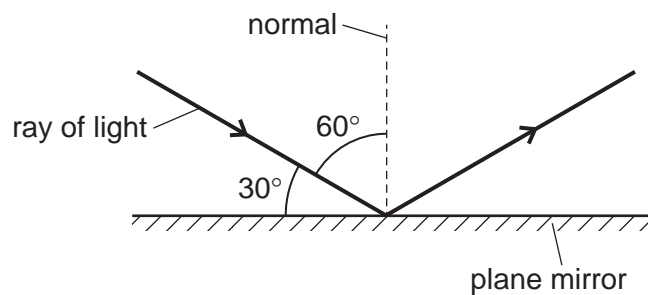
One wavefront (crest) is made every 0.50 s.



What is the frequency of the wave and what is its wavelength?

	frequency / Hz	wavelength / cm
A	0.50	3.0
B	0.50	6.0
C	2.0	3.0
D	2.0	6.0

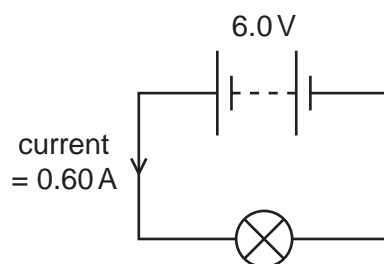
35 The diagram shows a ray of light which is reflected from a plane mirror.



What is the angle of incidence and what is the angle of reflection?

	angle of incidence / °	angle of reflection / °
A	30	30
B	30	60
C	60	30
D	60	60

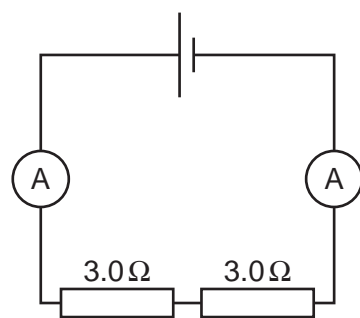
36 A 6.0 V battery is connected to a lamp. The current in the circuit is 0.60 A.



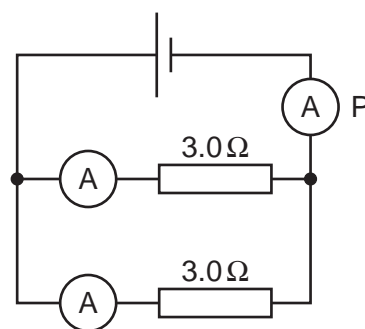
What is the resistance of the lamp?

- A** 0.10 Ω **B** 3.6 Ω **C** 10 Ω **D** 36 Ω

- 37 The diagrams show a series circuit and a parallel circuit. One ammeter in the parallel circuit is labelled P.



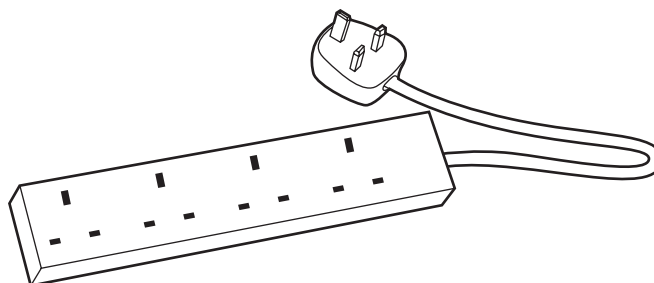
series circuit



parallel circuit

Which statement is correct?

- A The total resistance of the series circuit is $3.0\ \Omega$.
 - B The total resistance of the parallel circuit is $6.0\ \Omega$.
 - C In the series circuit, the readings on the ammeters are the same.
 - D In the parallel circuit, the reading on ammeter P is less than the reading on either of the other two ammeters.
- 38 An electrical extension block has four sockets, a cable which can safely take a current of 6 A and a plug. It is protected by a fuse rated at 5 A.



The extension block is used with four appliances and the 5 A fuse blows. The owner replaces the 5 A fuse with a 13 A fuse.

Why is the extension block now dangerous?

- A The appliances may not receive enough current.
- B The cable may overheat before the fuse blows.
- C The sockets may burn out before the fuse blows.
- D The 13 A fuse may blow too soon.

- 39 How do the ionising effect and the penetrating ability of alpha-emissions compare with those of beta-emissions?

	ionising effect	penetrating ability
A	alpha more ionising than beta	alpha more penetrating than beta
B	alpha more ionising than beta	alpha less penetrating than beta
C	alpha less ionising than beta	alpha more penetrating than beta
D	alpha less ionising than beta	alpha less penetrating than beta

- 40 A radioactive isotope has a half-life of 18 years. A sample contains 80 million atoms of this isotope.

How long does it take for the number of atoms of this isotope in the sample to decrease to 10 million?

- A** 2.25 years
- B** 6.0 years
- C** 54 years
- D** 180 years

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The Periodic Table of Elements

Group																																			
I	II	III										IV	V	VI	VII	VIII																			
3 Li lithium 7	4 Be beryllium 9	<div style="border: 1px solid black; padding: 5px; width: fit-content; margin: auto;"> Key atomic number atomic symbol name relative atomic mass </div>																2 He helium 4																	
11 Na sodium 23	12 Mg magnesium 24																	5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20	13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40	19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131	55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	113 Nh nihonium —	114 Fl flerovium —	115 Mc moscovium —	116 Lv livermorium —	117 Ts tennessine —	118 Og oganesson —																		

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).