


## Mark scheme – Identifying the Products of Chemical Reactions (H)

Question			Answer/Indicative content	Marks	Guidance
1			D ✓	1(AO 1.1)	<p><b><u>Examiner's Comments</u></b></p> <p>Candidates found this question challenging but there was no pattern of incorrect responses.</p>
			<b>Total</b>	<b>1</b>	
2			C ✓	1(AO 1.2)	<p><b><u>Examiner's Comments</u></b></p> <p> <b>Misconception</b></p> <p>A was a common misconception in this question.</p>
			<b>Total</b>	<b>1</b>	
3			D ✓	1(AO 1.1)	
			<b>Total</b>	<b>1</b>	
4			<p>Copper(II) ions – add aqueous sodium hydroxide (1)</p> <p>Gives a blue precipitate (1)</p> <p>Bromide ion – add aqueous silver nitrate followed by dilute nitric acid (1)</p> <p>Gives a cream precipitate (1)</p>	4	<p><b>ALLOW</b> any soluble metal hydroxide / aqueous ammonia</p> <p><b>ALLOW</b> blue solid / blue solid that redissolves into dark blue solution if ammonia is used</p>
			<b>Total</b>	<b>4</b>	
5			C	1	
			<b>Total</b>	<b>1</b>	