Questions

Q1.

Answer the question with a cross in the box you think is correct ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

When chloride ions are added to a pale blue solution containing copper ions, the mixture turns yellow.

This is a reversible reaction.

pale blue solution + chloride ions ↔ yellow solution + water

What effect does the removal of chloride ions have on the colour of the yellow mixture?

(1)

- ☐ **A** does not change colour
- ☐ **B** turns blue
- ☐ **C** turns colourless
- ☐ **D** turns darker yellow

(Total for question = 1 mark)

Q2.

Metals are extracted from substances naturally occurring in the Earth's crust.

Which of these metals is usually found uncombined in the Earth's crust?

(1)

- ☐ **A** calcium
- ☐ **B** gold
- ☐ **C** iron
- ☐ **D** magnesium

(Total for question = 1 mark)

Q3.

Answer the question with a cross in the box you think is correct ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

Some metals are found in the Earth's crust as uncombined elements.
Reactive metals are found in ores.

In ores, metals are combined with other elements.

Which of these metals is found as the uncombined element in the Earth's crust?

(1)

- ☐ **A** aluminium
- ☐ **B** gold
- ☐ **C** potassium
- ☐ **D** zinc

(Total for question = 1 mark)

Mark Scheme

Q1.

Question number	Answer	Mark
	B turns blue is the only correct answer. A, C and D are incorrect because the position of equilibrium will shift to the left-hand side	(1) AO2

Q2.

Question number	Answer	Mark
	B gold The only correct answer is B. A, C and D are incorrect because calcium, iron and magnesium respectively, are all found chemically combined to other elements in the Earth's crust.	(1)

Q3.

Question number	Answer	Mark
	B gold **A, C, D** – these metals too reactive to exist as uncombined elements	(1)