

THIS IS A NEW SPECIFICATION

**H**

Thursday 20 June 2013 – Afternoon

**GCSE TWENTY FIRST CENTURY SCIENCE
CHEMISTRY A****A173/02** Module C7 (Higher Tier)

Candidates answer on the Question Paper.
A calculator may be used for this paper.

OCR supplied materials:
None

Other materials required:

- Pencil
- Ruler (cm/mm)

Duration: 1 hour

Candidate forename		Candidate surname	
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Centre number						Candidate number				
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INSTRUCTIONS TO CANDIDATES

- Write your name, centre number and candidate number in the boxes above. Please write clearly and in capital letters.
- Use black ink. HB pencil may be used for graphs and diagrams only.
- Answer **all** the questions.
- Read each question carefully. Make sure you know what you have to do before starting your answer.
- Write your answer to each question in the space provided. Additional paper may be used if necessary but you must clearly show your candidate number, centre number and question number(s).
- Do **not** write in the bar codes.

INFORMATION FOR CANDIDATES

- Your quality of written communication is assessed in questions marked with a pencil (✎).
- The number of marks is given in brackets [] at the end of each question or part question.
- The total number of marks for this paper is **60**.
- This document consists of **20** pages. Any blank pages are indicated.
- The Periodic Table is printed on the back page.

2

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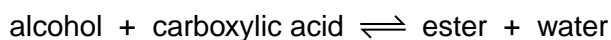
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3

Answer **all** the questions.

- 1 Esters are made by reacting an alcohol with a carboxylic acid.

A few drops of concentrated sulfuric acid are added to the mixture.



- (a) What is the job of the concentrated sulfuric acid?

..... [1]

- (b) The reaction mixture has to be heated for some time.

The reactants are volatile and would escape from an open flask as vapour.

- (i) What piece of apparatus is fitted to the flask to prevent this loss of vapour?

..... [1]

- (ii) What is the name of the technique used to prevent this loss of vapour?

..... [1]

- (c) Propyl pentanoate, $\text{C}_4\text{H}_9\text{COOC}_3\text{H}_7$, is an ester that tastes of pineapples.

It is made by reacting propanol, $\text{C}_3\text{H}_7\text{OH}$ with pentanoic acid $\text{C}_4\text{H}_9\text{COOH}$.

- (i) Write a symbol equation for the reaction that makes propyl pentanoate.

..... \rightleftharpoons [1]

- (ii) Suggest a use for propyl pentanoate.

..... [1]

[Total: 5]

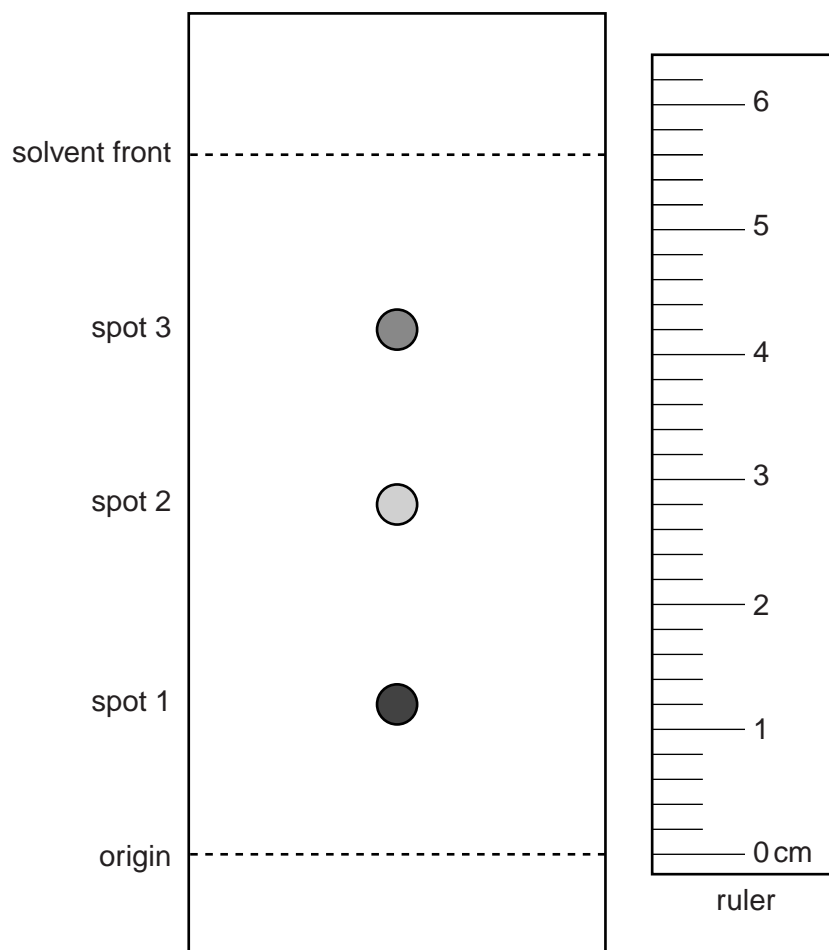
4

2 Peter is testing the food colourings in soft drinks made by different companies.

He wants to find out if any of these drinks contains a banned dye chemical.

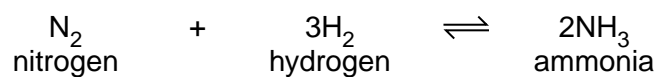
Peter uses paper chromatography to separate and identify the dye chemicals in the soft drinks.

Here is Peter's chromatogram for one of the soft drinks.

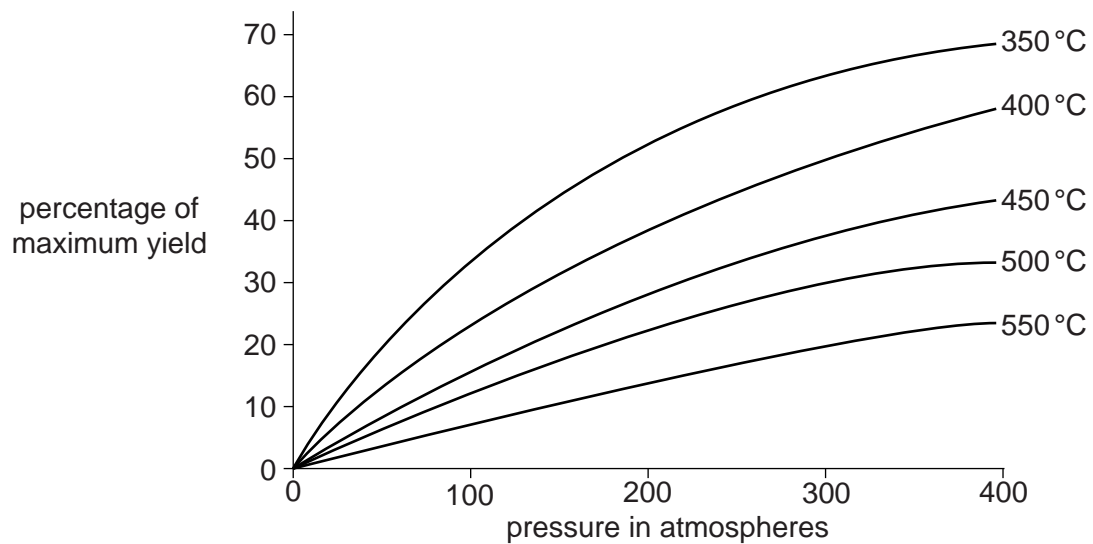


6

3 Ammonia, NH_3 , is made from nitrogen and hydrogen by the Haber process.



(a) The graph shows how the yield of ammonia is related to both the temperature and the pressure used.



8

- (ii) The reaction between nitrogen and hydrogen to form ammonia is reversible.

The reaction mixture can reach a dynamic equilibrium.

Which of these statements describes what is happening at equilibrium?

Put ticks (✓) in the boxes next to the **two** correct statements.

The reaction between nitrogen and hydrogen has stopped.

The forward and reverse reactions happen at the same rate.

All of the nitrogen and hydrogen react to make ammonia.

The concentration of ammonia is increasing.

The concentrations of nitrogen, hydrogen and ammonia are constant.

[2]

- (b) (i) Work out the relative formula mass (RFM) of ammonia, NH_3 .

RFM of ammonia = [1]

- (ii) What is the mass of ammonia that would be made if 1.0 tonne of nitrogen reacted completely with hydrogen?

Show your working.

mass of ammonia = tonne [2]

- (iii) A factory converts 95% of the nitrogen into ammonia.

What mass of ammonia does this factory make from each tonne of nitrogen?

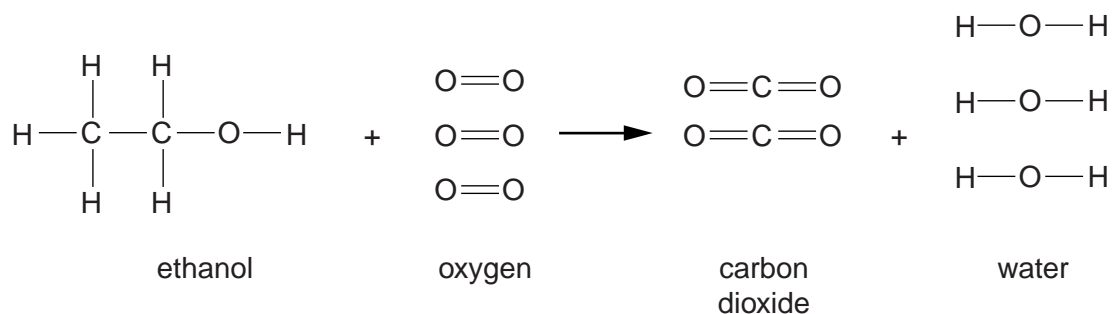
mass of ammonia = tonne [1]

10

4 The table shows the energy involved in the breaking of some bonds.

Bond	Bond energy in kJ/mol
C—H	411
C—O	358
O=O	498
C=O	799
O—H	459
C—C	348

Ethanol burns to make carbon dioxide and water.



(a) The energy needed to break all of the bonds in the oxygen, 3O_2 , is 1494 kJ.

Work out the energy needed to break all of the bonds in ethanol, $\text{C}_2\text{H}_5\text{OH}$.

energy = kJ/mol [2]

(b) The energy given out when new bonds in 2CO_2 are made is 3196 kJ.

Work out the energy given out when new bonds in the water, $3\text{H}_2\text{O}$, are made.

energy = kJ/mol [2]

(c) Use the information given, and your answers from (a) and (b), to complete the table.

	Energy in kJ/mol
energy needed to break all the bonds in ethanol and oxygen	
energy given out when all the bonds in carbon dioxide and water are made	
energy change when ethanol burns	

[2]

(d) The table shows information about some bonds.

Bond	Bond energy in kJ/mol	Bond length in pm
C—C	348	154
C=C	614	134
C≡C	839	120

What conclusions can you make from this data?

.....

.....

.....

.....

..... [2]

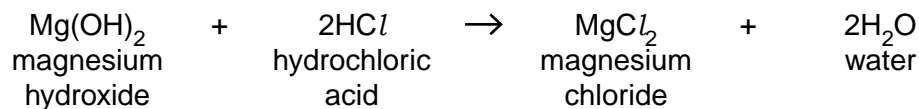
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5 Emma works for a company making indigestion tablets.

Her job is to find the mass of magnesium hydroxide in tablets from each batch.

Emma titrates the magnesium hydroxide in each tablet with hydrochloric acid.



(a) (i) What steps should Emma take to ensure that her titration results are as accurate as possible?

Put ticks (✓) in the boxes next to the **two** correct statements.

Add the acid as quickly as possible.

Use a burette to measure the volume of acid.

Add acid 1 cm³ at a time.

Add acid drop by drop near the end point.

Measure the mass of the flask every 30 seconds.

Allow the product to crystallise.

[2]

(ii) Emma uses an indicator in her titration.

Explain why she uses an indicator.

.....

.....

..... [2]

13

(b) Emma analyses an indigestion tablet to find the mass of magnesium hydroxide.

(i) Work out the relative formula mass (RFM) of magnesium hydroxide, $\text{Mg}(\text{OH})_2$.

Show your working.

(Relative atomic masses: H = 1; Mg = 24; O = 16)

RFM of magnesium hydroxide = [1]

(ii) Emma uses hydrochloric acid with 73.0g of hydrogen chloride in each 1.0dm^3 of the acid solution.

It takes 15.1cm^3 of this hydrochloric acid to neutralise the tablet.

Work out the mass of hydrogen chloride in 15.1cm^3 of the hydrochloric acid.

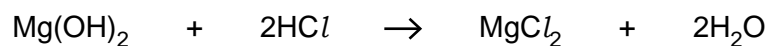
Give your answer to the nearest 0.1 g.

Show your working.

mass of hydrogen chloride = g [2]

(iii) Work out the mass of magnesium hydroxide in the tablet.

Use your answers to (i) and (ii) and this equation to help you.



Show your working.

mass of magnesium hydroxide in the tablet = g [2]

16

(b) How do the reactions of ethanol, ethane and water with sodium compare?

Put ticks (✓) in the correct boxes to show what happens in each reaction.

	Reaction of ethanol with sodium	Reaction of ethane with sodium	Reaction of water with sodium
violent reaction			
steady reaction			
no reaction			
hydrogen made			
sodium ethoxide made			
sodium hydroxide made			

[3]

[Total: 9]

END OF QUESTION PAPER

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