

Write your name here	
Surname	Other names
<b>Pearson</b> <b>Edexcel GCSE</b>	Centre Number
	Candidate Number
<b>Chemistry/Additional Science</b>	
<b>Unit C2: Discovering Chemistry</b>	
<b>Higher Tier</b>	
Tuesday 10 June 2014 – Afternoon <b>Time: 1 hour</b>	Paper Reference <b>5CH2H/01</b>
<b>You must have:</b> Calculator, ruler	Total Marks

### Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided  
– *there may be more space than you need.*

### Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets  
– *use this as a guide as to how much time to spend on each question.*
- Questions labelled with an **asterisk** (\*) are ones where the quality of your written communication will be assessed  
– *you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.*

### Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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# The Periodic Table of the Elements

	1	2		3	4	5	6	7	0								
	<b>7</b> <b>Li</b> lithium 3	<b>9</b> <b>Be</b> beryllium 4	<div style="display: flex; justify-content: center; align-items: center; gap: 20px;"> <div style="border: 1px solid black; padding: 2px 10px;">                     1 <b>H</b> hydrogen 1                 </div> <div style="border: 1px solid black; padding: 5px 10px;">                     Key                       relative atomic mass atomic symbol name atomic (proton) number                 </div> </div>						<b>11</b> <b>B</b> boron 5	<b>12</b> <b>C</b> carbon 6	<b>14</b> <b>N</b> nitrogen 7	<b>16</b> <b>O</b> oxygen 8	<b>19</b> <b>F</b> fluorine 9	<b>20</b> <b>Ne</b> neon 10			
<b>23</b> <b>Na</b> sodium 11	<b>24</b> <b>Mg</b> magnesium 12	<b>27</b> <b>Al</b> aluminium 13							<b>28</b> <b>Si</b> silicon 14	<b>31</b> <b>P</b> phosphorus 15	<b>32</b> <b>S</b> sulfur 16	<b>35.5</b> <b>Cl</b> chlorine 17	<b>40</b> <b>Ar</b> argon 18				
<b>39</b> <b>K</b> potassium 19	<b>40</b> <b>Ca</b> calcium 20	<b>45</b> <b>Sc</b> scandium 21	<b>48</b> <b>Ti</b> titanium 22	<b>51</b> <b>V</b> vanadium 23	<b>52</b> <b>Cr</b> chromium 24	<b>55</b> <b>Mn</b> manganese 25	<b>56</b> <b>Fe</b> iron 26	<b>59</b> <b>Co</b> cobalt 27	<b>59</b> <b>Ni</b> nickel 28	<b>63.5</b> <b>Cu</b> copper 29	<b>65</b> <b>Zn</b> zinc 30	<b>70</b> <b>Ga</b> gallium 31	<b>73</b> <b>Ge</b> germanium 32	<b>75</b> <b>As</b> arsenic 33	<b>79</b> <b>Se</b> selenium 34	<b>80</b> <b>Br</b> bromine 35	<b>84</b> <b>Kr</b> krypton 36
<b>85</b> <b>Rb</b> rubidium 37	<b>88</b> <b>Sr</b> strontium 38	<b>89</b> <b>Y</b> yttrium 39	<b>91</b> <b>Zr</b> zirconium 40	<b>93</b> <b>Nb</b> niobium 41	<b>96</b> <b>Mo</b> molybdenum 42	<b>[98]</b> <b>Tc</b> technetium 43	<b>101</b> <b>Ru</b> ruthenium 44	<b>103</b> <b>Rh</b> rhodium 45	<b>106</b> <b>Pd</b> palladium 46	<b>112</b> <b>Cd</b> cadmium 48	<b>115</b> <b>Ind</b> indium 49	<b>119</b> <b>Sn</b> tin 50	<b>122</b> <b>Sb</b> antimony 51	<b>127</b> <b>I</b> iodine 53	<b>128</b> <b>Te</b> tellurium 52	<b>131</b> <b>Xe</b> xenon 54	
<b>133</b> <b>Cs</b> caesium 55	<b>137</b> <b>Ba</b> barium 56	<b>139</b> <b>La*</b> lanthanum 57	<b>178</b> <b>Hf</b> hafnium 72	<b>181</b> <b>Ta</b> tantalum 73	<b>184</b> <b>W</b> tungsten 74	<b>186</b> <b>Re</b> rhenium 75	<b>190</b> <b>Os</b> osmium 76	<b>192</b> <b>Ir</b> iridium 77	<b>195</b> <b>Pt</b> platinum 78	<b>197</b> <b>Au</b> gold 79	<b>201</b> <b>Hg</b> mercury 80	<b>204</b> <b>Tl</b> thallium 81	<b>207</b> <b>Pb</b> lead 82	<b>209</b> <b>Bi</b> bismuth 83	<b>[209]</b> <b>Po</b> polonium 84	<b>[210]</b> <b>At</b> astatine 85	<b>[222]</b> <b>Rn</b> radon 86
<b>[223]</b> <b>Fr</b> francium 87	<b>[226]</b> <b>Ra</b> radium 88	<b>[227]</b> <b>Ac*</b> actinium 89	<b>[261]</b> <b>Rf</b> rutherfordium 104	<b>[262]</b> <b>Db</b> dubnium 105	<b>[266]</b> <b>Sg</b> seaborgium 106	<b>[264]</b> <b>Bh</b> bohrium 107	<b>[277]</b> <b>Hs</b> hassium 108	<b>[268]</b> <b>Mt</b> meitnerium 109	<b>[271]</b> <b>Ds</b> darmstadtium 110	<b>[272]</b> <b>Rg</b> roentgenium 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated						

\* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.



### Answer ALL questions

Some questions must be answered with a cross . If you change your mind about an answer, put a line through the box  and then mark your new answer with a cross .

### Particles and formulae

- 1 (a) Atoms contain protons, neutrons and electrons.

Complete the table to show the relative mass and relative charge of each particle and its position in an atom.

(3)

	relative mass	relative charge	position in atom
proton		+1	
neutron	1		in nucleus
electron			

- (b) Complete the sentence by putting a cross () in the box next to your answer.

An atom of an element **always** contains

(1)

- A** more protons than neutrons
- B** equal numbers of protons and neutrons
- C** more electrons than protons
- D** equal numbers of protons and electrons

- (c) The symbols for some atoms are given in the box

Ca	Cl	K	N	Ne	O
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From the box, choose the symbol of

- (i) an atom in group 2 of the periodic table

(1)

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- (ii) an atom that readily forms an ion with a charge of 2-

(1)

.....



(d) The formula of aluminium nitrate is  $\text{Al}(\text{NO}_3)_3$

(i) State the total number of atoms in the formula  $\text{Al}(\text{NO}_3)_3$

(1)

(ii) What is the most likely formula of aluminium nitride?

Put a cross (☒) in the box next to your answer.

(1)

**A**  $\text{Al}(\text{NO}_3)_2$

**B**  $\text{AlNO}_3$

**C**  $\text{AlNO}_2$

**D**  $\text{AlN}$

**(Total for Question 1 = 8 marks)**



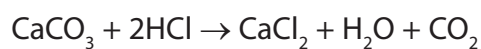
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### Rates of reactions and energy changes

- 2 (a) Marble chips react with hydrochloric acid to produce carbon dioxide.

The equation for the reaction is



Which one of these changes would **decrease** the rate of this reaction?

Put a cross (☒) in the box next to your answer.

(1)

- A use hydrochloric acid which is more dilute
- B use smaller sized marble chips
- C use marble chips which have a larger surface area
- D use a larger volume of the hydrochloric acid

- (b) Explain why increasing the temperature of a reaction increases the rate of the reaction.

(2)

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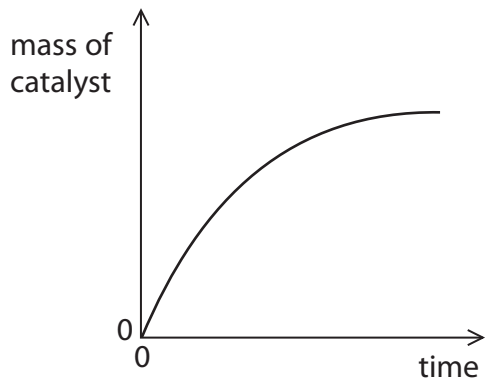
(c) (i) The rate of decomposition of hydrogen peroxide can be increased by adding a catalyst.

Which of these graphs shows the mass of the catalyst during the reaction?

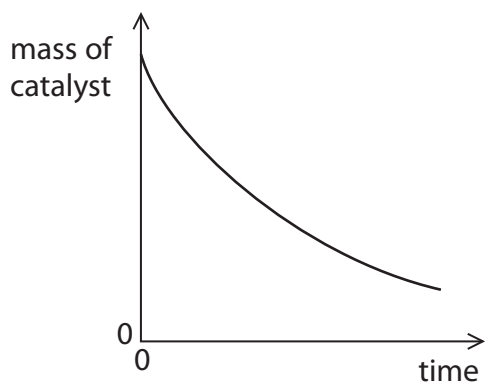
Put a cross (☒) in the box next to your answer.

(1)

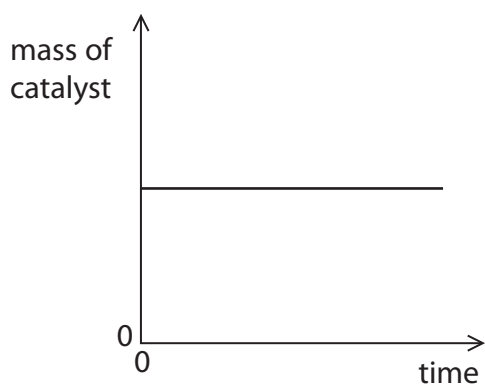
**A**



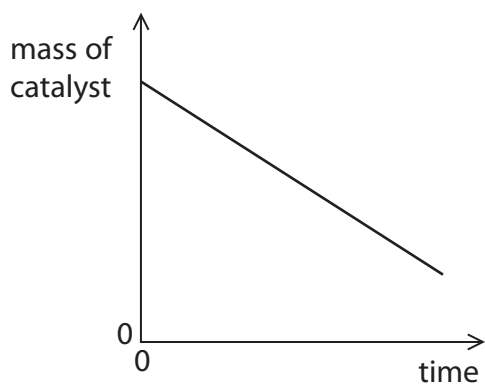
**B**



**C**



**D**





(ii) The decomposition of hydrogen peroxide,  $\text{H}_2\text{O}_2$ , produces oxygen and water.

Give the balanced equation for this reaction.

(2)

(d) Explain, in terms of the energy involved in the breaking of bonds and in the making of bonds, why some reactions are exothermic.

(2)

**(Total for Question 2 = 8 marks)**



## Metals

**3** There are many metallic elements in the periodic table.

(a) Which row of the table correctly shows two metals that are in group 1 and two metals that are transition metals?

Put a cross (☒) in the box next to your answer.

(1)

	group 1	transition metals
<input checked="" type="checkbox"/> <b>A</b>	lithium and zinc	calcium and copper
<input checked="" type="checkbox"/> <b>B</b>	potassium and caesium	copper and iron
<input checked="" type="checkbox"/> <b>C</b>	sodium and potassium	copper and magnesium
<input checked="" type="checkbox"/> <b>D</b>	sodium and magnesium	manganese and nickel

(b) (i) Describe the structure of metals in terms of the particles present in their structures.

(2)

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(ii) Explain how metals conduct electricity.

(2)

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(c) (i) Describe what you would **see** when a small piece of sodium is added to water.

(2)

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(ii) Write the balanced equation for the reaction of sodium with water to form sodium hydroxide and hydrogen.

(3)

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**(Total for Question 3 = 10 marks)**



**Salts and analysis**

- 4 (a) Which of the following pairs of substances contains one substance that is soluble in water and one that is insoluble in water?

Put a cross (☒) in the box next to your answer.

(1)

- A** aluminium nitrate and lead sulfate
- B** ammonium chloride and copper sulfate
- C** copper hydroxide and lead sulfate
- D** sodium hydroxide and potassium nitrate

- (b) Barium chloride is an ionic compound and has a high melting point.

Explain why barium chloride has a high melting point.

(2)

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- (c) Barium chloride solution is used to test for the presence of sulfate ions in a solution.

When sulfate ions are present, insoluble barium sulfate is formed.

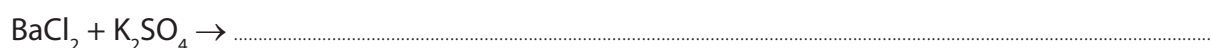
- (i) Describe the appearance of barium sulfate.

(1)

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- (ii) Complete the balanced equation for the reaction between barium chloride and potassium sulfate.

(2)



(d) Compound **X** is a metal carbonate.

- (i) A flame test was carried out on compound **X**.  
A lilac flame was seen.

Complete the sentence by putting a cross (☒) in the box next to your answer.

(1)

The formula of the metal ion in compound **X** is

- A**  $\text{Ca}^{2+}$   
 **B**  $\text{Cu}^{2+}$   
 **C**  $\text{K}^+$   
 **D**  $\text{Na}^+$

- (ii) Lead carbonate is an insoluble salt.

Describe how a pure, dry sample of solid lead carbonate can be obtained from sodium carbonate solution and lead nitrate solution.

(3)

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**(Total for Question 4 = 10 marks)**



### Chlorine and carbon

- 5 (a) Chlorine has an atomic number of 17.  
Chlorine-35 and chlorine-37 are two isotopes of chlorine.

(i) Complete the table to show the numbers of protons, neutrons and electrons in each of the isotopes.

(2)

	chlorine-35	chlorine-37
number of protons		
number of neutrons		
number of electrons		

(ii) A normal sample of chlorine contains only chlorine-35 and chlorine-37 atoms.

Explain why the relative atomic mass of chlorine is 35.5

(2)

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- (b) Tetrachloromethane is a simple molecular, covalent compound.  
The formula of its molecule is  $\text{CCl}_4$ .

There are four electrons in the outer shell of a carbon atom.  
There are seven electrons in the outer shell of a chlorine atom.

Draw a dot and cross diagram to show the bonding in a molecule of tetrachloromethane,  $\text{CCl}_4$ .

Show outer shell electrons only.

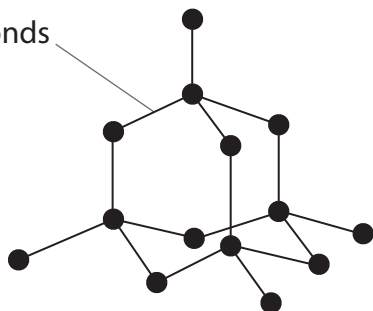
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\*(c) The diagrams show the arrangements of carbon atoms in diamond and in graphite.

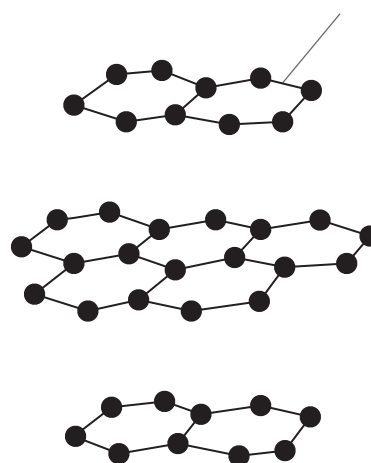
● = carbon atom

strong bonds



diamond

strong bonds



graphite

Compare a use of diamond with a use of graphite, explaining each use in terms of the bonding and structure. In your answer you should use information from the diagrams.

(6)

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(Total for Question 5 = 12 marks)



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### Group 7 elements

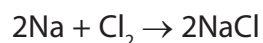
- 6 (a) A compound of iron and chlorine was formed by reacting 2.80 g of iron with 3.55 g of chlorine.

Calculate the empirical formula of the compound.  
(relative atomic masses: Cl = 35.5, Fe = 56.0)

(3)

empirical formula .....

- (b) Sodium reacts with chlorine to form sodium chloride.



Calculate the maximum mass of sodium chloride that could be formed by reacting 9.20 g of sodium with excess chlorine.  
(relative atomic masses: Na = 23.0, Cl = 35.5)

(3)

mass of sodium chloride ..... g



\*(c) Chlorine, bromine and iodine are in group 7 of the periodic table.

The order of reactivity of these three elements can be shown by carrying out displacement experiments.

You are provided with

potassium bromide solution

potassium chloride solution

potassium iodide solution

bromine solution

chlorine solution

iodine solution

Describe how these solutions could be used to carry out experiments to show the order of reactivity of bromine, chlorine and iodine, explaining how the results would show the order of reactivity.

You may use equations if you wish.

(6)

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**(Total for Question 6 = 12 marks)**

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**TOTAL FOR PAPER = 60 MARKS**



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