



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/52

Paper 5 Practical Test

October/November 2016

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2016 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.

© IGCSE is the registered trademark of Cambridge International Examinations.

This syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **4** printed pages.

Page 2	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	52

Question	Answer	Marks
1(a)	table of results for Experiment 1 initial and other temperature boxes completed correctly and results comparable to supervisor's and Experiment 1 maximum greater than Experiment 2 maximum	1 1
1(b)	table of results for Experiment 2 initial and other temperature boxes completed correctly and comparable to supervisor's and temperatures stop rising	1 1
1(c)	all points correctly plotted best-fit smooth line graphs labels	2 1 1
1(d)	value from graph shown clearly	1 1
1(e)	phenolphthalein/litmus/suitable named indicator	1
1(f)	Experiment 1 solution N is a stronger acid/has a higher pH	1 1
1(g)	measured results/temperature changes/results would be smaller OR larger/double volume needed to reach same temperature changes	1
1(h)	polystyrene is an insulator/copper is a (good) conductor	1
1(i)	source of error: heat losses/using a measuring cylinder improvement: lag/insulate/use burette;	1 1

Page 3	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	52

Question	Answer	Mark
2(a)	white crystals	1
2(b)(i)	melts/liquefies/dissolves/bubbles condensation/drops of liquid cobalt(II) chloride paper turns colourless/light pink	1 1 1
2(b)(ii)	no change / colour	1
2(c)(i)	white precipitate dissolves / clears	1 1 1
2(c)(ii)	white precipitate	1
2(c)(iii)	no reaction / no change	1
2(c)(iv)	white precipitate	1 1
2(d)	not a transition element (cation)	1
2(e)	hydrated / contains water	1
2(f)	it is not any named metal that gives a positive flame test	1
2(g)	aluminium sulfate	1 1

Page 4	Mark Scheme	Syllabus	Paper
	Cambridge IGCSE – October/November 2016	0620	52

Question	Answer	Mark
3	<p>method adding Agri Lime to acid add weighed amount/known mass of Agri Lime Q to a known volume of acid with a named indicator added to the acid until the indicator changes colour note the mass of Agri Lime Q added repeat with Agri Lime R conclusion, e.g. 'the experiment using the smaller amount of Agri Lime is better'</p> <p>OR</p> <p>method adding acid to Agri Lime use weighed amount/known mass of Agri Lime Q add acid to it gradually/from a burette with a named indicator added to the acid until the indicator changes colour note volume of acid added repeat with Agri Lime R conclusion, e.g. 'the experiment using the larger volume of acid is better'</p>	6