

**CAMBRIDGE**  
INTERNATIONAL EXAMINATIONS

**June 2003**

**INTERNATIONAL GCSE**

**MARKING SCHEME**

**MAXIMUM MARK: 40**

**SYLLABUS/COMPONENT: 0620/05**

**CHEMISTRY**

**(Practical)**



Page 1	Mark Scheme	Syllabus	Paper
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1	Table of results			
	Experiment 1	Initial and final readings recorded		[1]
		to 1 decimal place		[1]
	Experiment 2	Initial and final readings recorded		[1]
		to 1 decimal place		[1]
		Results comparable to Supervisor's results $\pm 1\text{cm}^3$		[2]
	(a)	red/burgundy/brown		[1]
	(b)	yellow (1) to blue/black (1)	see Supervisor	[2]
		IGNORE green		
	(c) (i)	Experiment 1		[1]
	(ii)	$\Delta$ 2 x, double volume (1) in Experiment 1 (1) <u>not</u> just more		[2]
	(iii)	potassium iodate less concentrated solution <b>C</b> than <b>B</b> or vice versa		[1]
		<u>not</u> different concentrations		
(iv)	2 x volume from table for Experiment 1 (1) unit (1)		[2]	
	2 x iodine formed		[1]	
(d)	Indicator (1) reference to accuracy (1)/end-point/see more clearly		[2]	
	<u>not</u> test for $\text{I}_2/\text{I}^-$			
				<b>[Question total: 18]</b>
2	(a)	bubbles/condensation/goes black	max 2	[2]
	(b)	filtrate - colourless <u>not</u> clear		[1]
		residue - green		[1]
	(c) (i)	effervescence/fizz/bubbles		[1]
		limewater $\rightarrow$ milky		[1]
		solution is blue		[1]
	(ii)	blue (1) precipitate (1)		[2]
		royal/deep blue (1) solution (1)		[2]
	(d) (i)	white (1) precipitate (1) dissolves in excess (1)		[3]
	(ii)	white (1) precipitate (1) dissolves (1)		[3]
	(iii)	white precipitate (1)		[1]
	(e)	zinc (1) sulphate (1)	reversed = 0	[2]
(f)	copper (1) carbonate (1)	reversed = 0	[2]	
	hydrated (1)		max 2	
				<b>[Question total: 22]</b>
				<b>[Total for paper: 40]</b>

Results obtained for Question 1/cm<sup>3</sup>

	1 <sup>st</sup>	2 <sup>nd</sup>
Experiment 1	16.5	16.3
Experiment 2	8.3	8.2