

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

MARK SCHEME for the October/November 2008 question paper

0620 CHEMISTRY

0620/32

Paper 32 (Extended Theory), maximum raw mark 80

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Page 2	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2008	0620	32

- 1 ammonia [1]
 chlorine [1]
 “pop” with a lighted splint **or** burn with a pop **or** goes pop and extinguishes flame [1]
NOT glowing splint
 relights a glowing splint [1]
 turns limewater milky/cloudy/chalky/white [1]
ACCEPT correct formulae

[Total: 5]

- 2 (a) 2Na : 1S correct ratio [1]
 correct charges [1]
 8e around S [1]

if no symbols then must have correct key
 if covalent only mark 1
 ignore electrons around sodium
 if the response includes both a correct and an incorrect answer
 do not select correct one, mark = [0]

- (b) (i) positive ions **or** cations [1]
NOT atoms **or** cores **or** nuclei
 layers **or** lattice **or** regular pattern [1]
 delocalised **or** free **or** mobile electrons **or** sea [1]
- OR** positive ions **or** cations [1]
NOT atoms **or** cores **or** nuclei
 attraction between ions and electrons [1]
 delocalised **or** free **or** mobile electrons **or** sea [1]
 the attraction/electrostatic bonding must be between ions and
 delocalised electrons, between cations and anions does not score
ACCEPT bond if qualified e.g. electrostatic bond, etc.
 if moles or molecular cannot score cation mark
- (ii) delocalised/free/mobile electrons
or electrons can move [1]
- layers **or** ions **or** atoms **or** particles [1]
NB more flexible than 2(b)(i)
 can slip **or** move past each other **or** bonding non-directional [1]

Page 3	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2008	0620	32

- (c) (i) tetrahedral [1]
 1Si : 4O bonded/surrounded, etc. [1]
 1O : 2 Si [1]

NOT molecules of oxygen, etc.

NOT intermolecular forces

ONLY tetrahedral can score for either of the above

Despite what the question states, **ACCEPT** a clear accurate diagram which shows the above three points.

- (ii) hard
 high melting point **or** boiling point
 colourless (**NOT** clear) **or** shiny **or** translucent
 non/poor conductor (of electricity)
 brittle
 insoluble
 any **TWO** [2]
NOT crystalline **or** strong

[Total: 14]

- 3 (a) (i) water **or** moisture **ACCEPT** salty water [1]
 air **or** oxygen [1]

- (ii) galvanising **or** coat with zinc
 tin plate
 chromium plate
 nickel plate
 cobalt plate
 copper plate
 cover with aluminium
 anodic protection **or** sacrificial protection
 cathodic protection
 cover with plastic
 alloying (ignore any named metal)
 any **TWO** [2]
NOT just plate **or** electroplate need electroplate with suitable metal
NOT oil
ACCEPT both galvanising and sacrificial protection

- (b) (i) hydrogen **or** carbon **or** carbon monoxide **or** methane [1]
or more reactive metal **NOT** Group I

- (ii) any correct equation [2]
 only error not balanced [1]

Page 4	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2008	0620	32

- (c) (i) 196 [1]
- (ii) $36/196 \times 100$ [1]
 $= 18(.4)\%$ **ACCEPT** 18 to nearest whole number [1]
 mark **e.c.f.** to (c)(i) provided percentage not greater than 100%
ONLY ACCEPT $36/\text{answer (c)(i)} \times 100$
 otherwise [0]

- (d) (i) forms carbon dioxide/carbon monoxide (which escapes) [1]
- (ii) forms silicon(IV) oxide **or** silicon oxide **or** silica [1]
OR CaO reacts with SiO₂
 to form slag **or** calcium silicate [1]
 ignore an incorrect formula if a correct name given
NOT Si + O₂ + CaO form slag

[Total: 13]

- 4 (a) (i) C₆H₅COOH **or** C₆H₅CO₂H [1]
NOT C₇H₆O₂ / C₆H₆COO
- (ii) sodium hydroxide + benzoic acid = sodium benzoate + water [1]
 correct spelling needed **NOT** benzenoate
ACCEPT correct symbol equation
- (iii) sodium carbonate **or** oxide **or** hydrogencarbonate [2]
 any **TWO**
NOT Na
- (b) (i) 7.7% [1]
- (ii) for any number: equal number ratio [2]
 for example 1:1 or 6:6
- (iii) empirical formula is CH [1]
 molecular formula is C₆H₆ [1]
 no e.c.f., award of marks not dependent on (ii)
- (c) (i) C₆H₈O₆ [1]
- (ii) carbon – carbon double bond **or** alkene [1]
 alcohol **or** hydroxyl **or** hydroxy [1]
NOT hydroxide
 hydroxide and alcohol = 0

[Total: 12]

Page 5	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2008	0620	32

- 5 (a) (i) $2\text{H}^+ + 2\text{e} \rightarrow \text{H}_2$ [1]
- (ii) $2\text{Cl}^- - 2\text{e} \rightarrow \text{Cl}_2$ or $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}$ [1]
- (iii) Na^+ and OH^- are left [1]
OR Cl^- removed OH^- left
NB ions by name or formula essential
NOT any reaction of Na or Na^+
NOT Na^+ and OH^- combine
- (b) (i) sterilise/disinfect water or kill microbes/germs bacteria, etc. [1]
NOT just to make it safe to drink or purify it or clean it
 treat above as neutral they do not negate a correct response
- (ii) ammonia or methanol or hydrogen chloride or margarine [1]
NOT nylon
- (iii) ester or triester or lipid [1]
 hydrolysis or saponification [1]

[Total: 7]

6 (a) (i)

aqueous solution	tin Sn	manganese Mn	silver Ag	zinc Zn
tin(II) nitrate		R	NR	R
manganese(II) nitrate	NR		NR	NR
silver(I) nitrate	R	R		R
zinc nitrate	NR	R	NR	

[1] for each row [3]
 ignore anything written in blank space

- (ii) $\text{Zn} + 2\text{AgNO}_3 \rightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$ [2]
 all species correct [1]
 accept correct ionic equation
 $\text{Zn} + 2\text{Ag}^+ \rightarrow \text{Zn}^{2+} + 2\text{Ag}$ [2]
- (iii) Sn^{2+} must be made clear that the oxidant is Sn^{2+} not Sn [1]
 it gains electrons or oxidation number decreases or it is reduced [1]
 reason must relate to an oxidant
NB not dependent on identifying Sn^{2+}
- (iv) covered with oxide layer [1]
 makes it unreactive or protects or aluminium oxide unreactive [1]

Page 6	Mark Scheme	Syllabus	Paper
	IGCSE – October/November 2008	0620	32

- (b) (i) potassium has one valency electron [1]
or loses one electron
 calcium has two valency electrons
or loses two electrons [1]
- (ii) potassium hydroxide → no reaction [1]
 calcium hydroxide → calcium oxide and water [1]
ACCEPT metal oxide
- (iii) $2\text{KNO}_3 \rightarrow 2\text{KNO}_2 + \text{O}_2$ [2]
 [1] for **formula** of either product
- $2\text{Ca}(\text{NO}_3)_2 \rightarrow 2\text{CaO} + 4\text{NO}_2 + \text{O}_2$ [2]
 [1] for **formulae** of any **TWO** products

[Total: 17]

- 7 (a) (i) 20 cm^3 [1]
 80 cm^3 [1]
- (ii) forms carbon monoxide [1]
 poisonous **or** toxic **or** lethal **or** prevents blood carrying oxygen
or effect on haemoglobin [1]
NOT just harmful, etc.
- (b) (i) chlorobutane **or** butyl chloride [1]
 number not required but if given must be 1, it must be in correct position
- (ii) light **or** UV **or** $200\text{ }^\circ\text{C}$ **or** lead tetraethyl [1]
- (iii) any correct equation for example 2-chlorobutane
or dichlorobutane
 must include HCl [1]
- (c) (i) correct repeat unit [1]
COND continuation [1]
 $-(\text{CH}(\text{CH}_3)-\text{CH}_2)-$
- (ii) propan-1-ol **or** propan-2-ol **or** propanol [1]
 if number given then formula must correspond for second mark.
 number must be in correct position
 structural formula of above [1]
 $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{OH}$ **or** $\text{CH}_3-\text{CH}(\text{OH})-\text{CH}_3$
NOT $\text{C}_3\text{H}_7\text{OH}$
 if first mark not awarded then either formula will gain mark [1].
accept either formula for "propanol" in (i)
NB On scoris both marks entered together not as [1] and [1] separately
- (iii) $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{Cl}$ **or** $\text{CH}_3-\text{CH}_2-\text{CH}(\text{Cl})-\text{CH}_3$ [1]
NOT $\text{C}_4\text{H}_9\text{Cl}$
 if equation given look at RHS only
 response must not include HCl

[Total: 12]