



**Cambridge International Examinations**  
Cambridge International General Certificate of Secondary Education

**CHEMISTRY**

**0620/11**

Paper 1 Multiple Choice (Core)

**October/November 2016**

**45 minutes**

Additional Materials:      Multiple Choice Answer Sheet  
   Soft clean eraser  
   Soft pencil (type B or HB is recommended)



**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **17** printed pages and **3** blank pages.

## 2

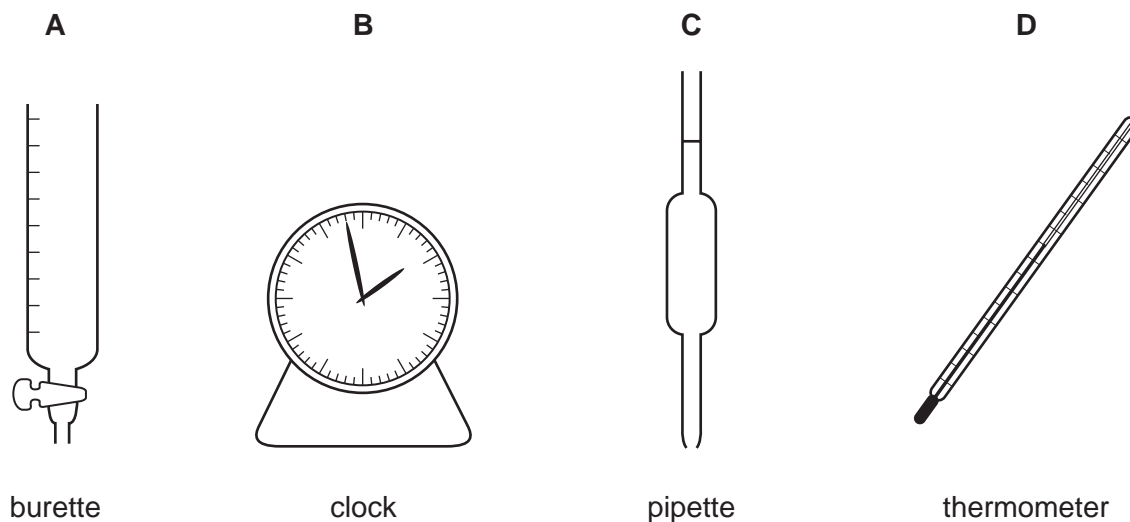
- 1 'Particles moving **very slowly** from an area of higher concentration to an area of lower concentration.'

Which process is being described?

- A a liquid being frozen
  - B a solid melting
  - C a substance diffusing through a liquid
  - D a substance diffusing through the air
- 2 A student mixes  $25\text{ cm}^3$  samples of dilute hydrochloric acid with different volumes of aqueous sodium hydroxide.

In each case, the student measures the change in temperature to test if the reaction is exothermic.

Which piece of apparatus is **not** needed?



## 3

3 Information about the solubility of four solids, P, Q, R and S, is given in the table.

	P	Q	R	S
solubility in water	dissolves	insoluble	insoluble	dissolves

A student attempted to separate mixtures of these solids using the following method.

- 1 Add the mixture to a beaker of water and stir.
- 2 Filter the mixture.
- 3 Crystallise one of the solids from the filtrate.

Which of the following mixtures could **not** be separated by this method?

- A** a mixture of P and R
- B** a mixture of Q and P
- C** a mixture of Q and R
- D** a mixture of R and S

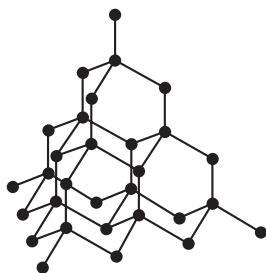
4 The table shows information about atoms of three different elements.

element	proton number	nucleon number	number of protons	number of neutrons	number of electrons
chlorine	17	35	17	W	17
chlorine	17	X	17	19	17
argon	Y	40	18	22	18
potassium	19	39	19	20	Z

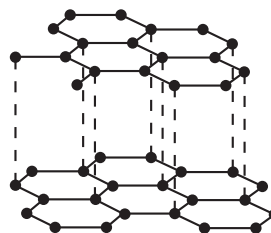
What are the values of W, X, Y and Z?

	W	X	Y	Z
<b>A</b>	18	35	18	19
<b>B</b>	18	36	18	19
<b>C</b>	19	35	19	18
<b>D</b>	19	36	19	18

- 5 The diagrams show the structures of two forms of the same element.



T



U

What are the reasons for using T in cutting tools and U as a lubricant?

	T	U
<b>A</b>	It is very hard because each atom is held in place by strong covalent bonds.	The layers can slide over each other because the covalent bonds are weak.
<b>B</b>	It is very hard because each atom is held in place by strong covalent bonds.	The layers can slide over each other due to weak forces between the layers.
<b>C</b>	It is very hard because there are no electrons able to move.	The layers can slide over each other because the covalent bonds are weak.
<b>D</b>	It is very hard because there are no electrons able to move.	The layers can slide over each other due to weak forces between the layers.

- 6 Ions are formed by elements losing or gaining electrons.

Which statement is correct?

- A** Metal atoms gain electrons to form positive ions.  
**B** Non-metal atoms lose electrons to form positive ions.  
**C** The charge on an ion is always either +1 or -1.  
**D** Group I ions have the same electronic structure as noble gases.
- 7 A molecule of X contains two carbon atoms, four hydrogen atoms and two oxygen atoms.

What is the formula of X?

- A**  $\text{CH}_2\text{CO}_2\text{H}$       **B**  $\text{CH}_3\text{COH}$       **C**  $\text{CH}_3\text{COOH}$       **D**  $\text{C}_2\text{H}_3\text{COOH}$

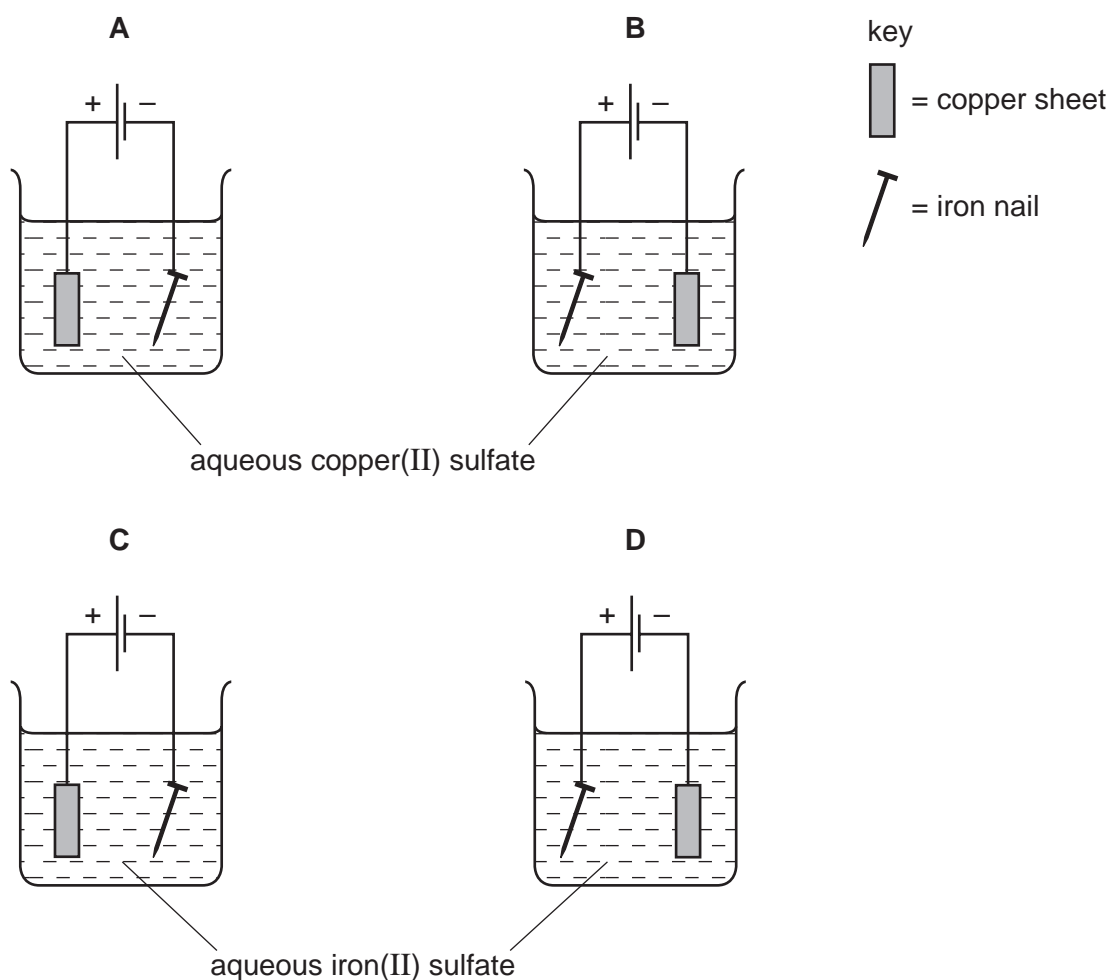
- 8 Concentrated aqueous potassium chloride is electrolysed using platinum electrodes.

The solution contains the ions  $K^+$ ,  $Cl^-$ ,  $H^+$  and  $OH^-$ .

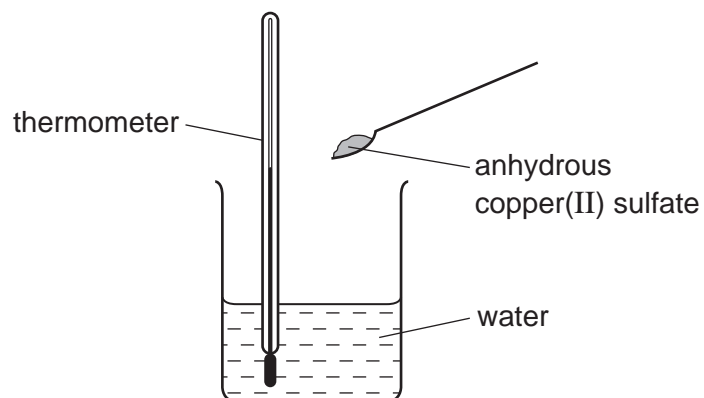
Which electrodes are the ions attracted to during this electrolysis?

	anode	cathode
<b>A</b>	$Cl^-$ and $K^+$	$H^+$ and $OH^-$
<b>B</b>	$Cl^-$ and $OH^-$	$H^+$ and $K^+$
<b>C</b>	$H^+$ and $K^+$	$Cl^-$ and $OH^-$
<b>D</b>	$H^+$ and $OH^-$	$Cl^-$ and $K^+$

- 9 Which apparatus could be used to electroplate an iron nail with copper?



10 When anhydrous copper(II) sulfate is added to water a solution is formed and heat is given out.



Which row shows the temperature change and the type of reaction taking place?

	temperature change	type of reaction
<b>A</b>	decrease	endothermic
<b>B</b>	decrease	exothermic
<b>C</b>	increase	endothermic
<b>D</b>	increase	exothermic

11 The combustion of element X releases large amounts of energy.

What is X?

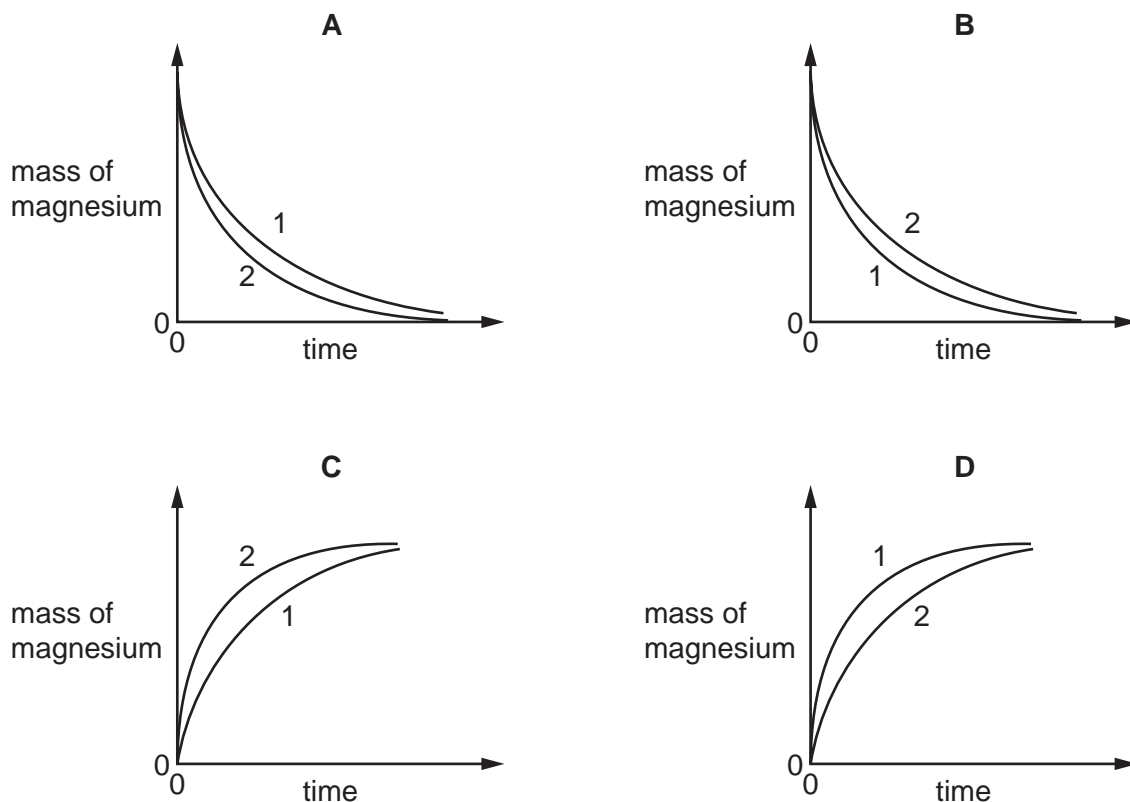
- A** ethanol
- B** hydrogen
- C** methane
- D** uranium

- 12 The rate of reaction between magnesium and excess dilute hydrochloric acid was followed by measuring the mass of magnesium present at regular time intervals.

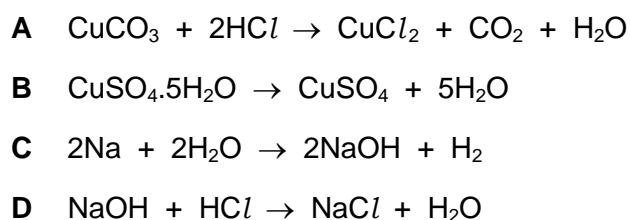
Two experiments were performed.

Both experiments used 0.1 g of magnesium ribbon. The acid in experiment 1 was less concentrated than in experiment 2.

Which graph shows the results of the experiments?



- 13 Which reaction is reversible?



- 14 Tin is formed when tin(II) oxide is heated with carbon.

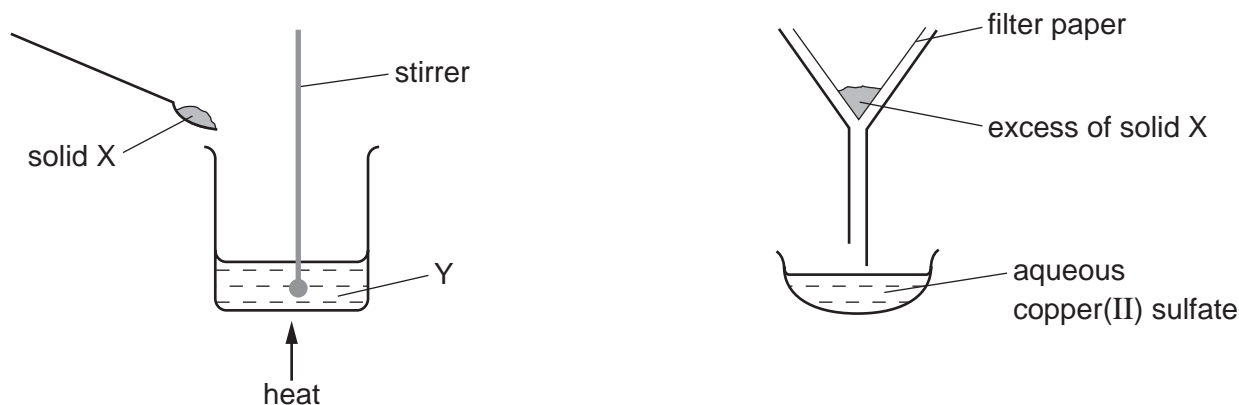
What happens to the tin in the tin(II) oxide in this reaction?

- A It is burnt.  
 B It is electrolysed.  
 C It is oxidised.  
 D It is reduced.





18 The apparatus shown is used to prepare aqueous copper(II) sulfate.



What are X and Y?

	X	Y
<b>A</b>	copper	aqueous iron(II) sulfate
<b>B</b>	copper(II) chloride	sulfuric acid
<b>C</b>	copper(II) oxide	sulfuric acid
<b>D</b>	sulfur	aqueous copper(II) chloride

19 Which statement about trends in the Periodic Table is **not** correct?

- A** Elements in the same period have the same number of electron shells.
- B** The elements change from metals to non-metals from left to right.
- C** The number of protons in an atom of an element increases from left to right.
- D** The oxides of the elements change from acidic to basic from left to right.

20 What is **not** a property of Group I metals?

- A** They are soft and can be cut with a knife.
- B** They react when exposed to oxygen in the air.
- C** They produce an acidic solution when they react with water.
- D** They react rapidly with water producing hydrogen gas.

21 Which statement about the element with proton number 54 is correct?

- A** It burns in the air to form an oxide.
- B** It could be used in balloons because it has a very low density.
- C** It is a gas at room temperature.
- D** It is reactive because it has a full outer shell of electrons.

22 Which element is a transition element?

	colour of chloride	melting point of element / °C
<b>A</b>	orange	113
<b>B</b>	orange	1535
<b>C</b>	white	113
<b>D</b>	white	1535

23 Which row describes the trends in the properties of the Group VII elements as the group is descended?

	colour	density	reactivity with halide ions
<b>A</b>	darkens	decreases	increases
<b>B</b>	darkens	increases	decreases
<b>C</b>	lightens	decreases	increases
<b>D</b>	lightens	increases	decreases

24 Four metals are listed in decreasing order of reactivity.

magnesium

zinc

iron

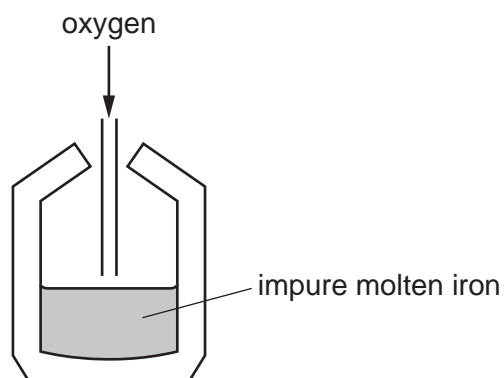
copper

Titanium reacts with acid and cannot be extracted from its ore by heating with carbon.

Where should titanium be placed in the list?

- A** below copper
- B** between iron and copper
- C** between magnesium and zinc
- D** between zinc and iron

25 Impure iron from the blast furnace is converted to steel as shown.



Which statement about the process is correct?

- A Acidic oxides are added to remove alkaline impurities.
- B Coke is added as a reducing agent.
- C Oxygen is blown in to oxidise the impure iron.
- D The steel produced contains less carbon than the impure iron.

26 A student added dilute hydrochloric acid to four metals and recorded the results.

Some of the results are **not** correct.

	results	
	metal	gas given off
1	copper	yes
2	iron	yes
3	magnesium	no
4	zinc	yes

Which **two** results are correct?

- A 1 and 3
- B 1 and 4
- C 2 and 3
- D 2 and 4

27 Some properties of three metals, P, Q and R, are shown.

metal	density	resistance to corrosion	electrical conductivity
P	low	high	very good
Q	high	high	very good
R	low	low	good

Which metals would be suitable for use in electrical wiring and aircraft manufacture?

	electrical wiring	aircraft manufacture
<b>A</b>	P	Q
<b>B</b>	Q	P
<b>C</b>	Q	R
<b>D</b>	R	P

28 One sample of sea-water is distilled while another sample of sea-water is filtered.

Which statement about the samples is correct?

- A** The distilled sample boils at exactly 100 °C and contains dissolved salts.
- B** The distilled sample boils at 103 °C and does **not** contain dissolved salts.
- C** The filtered sample boils at 103 °C and contains dissolved salts.
- D** The filtered sample boils at exactly 100 °C and does **not** contain dissolved salts.

29 Air is a mixture of gases.

Which gas is present in the largest amount?

- A** argon
- B** carbon dioxide
- C** nitrogen
- D** oxygen

30 Which information about carbon dioxide and methane is correct?

		carbon dioxide	methane
<b>A</b>	formed when vegetation decomposes	✓	✗
<b>B</b>	greenhouse gas	✓	✓
<b>C</b>	present in unpolluted air	✗	✗
<b>D</b>	produced during respiration	✗	✓

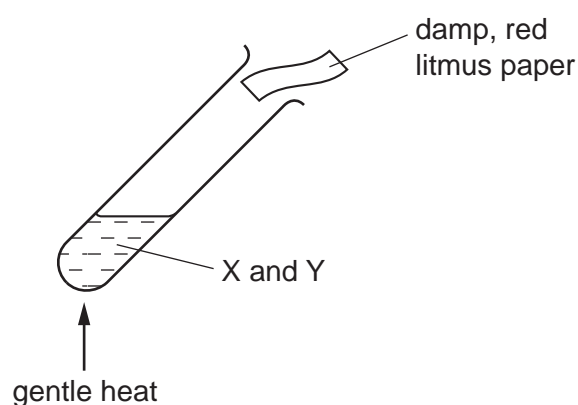
key

✓ = true

✗ = false

31 A mixture of two substances, X and Y, is heated.

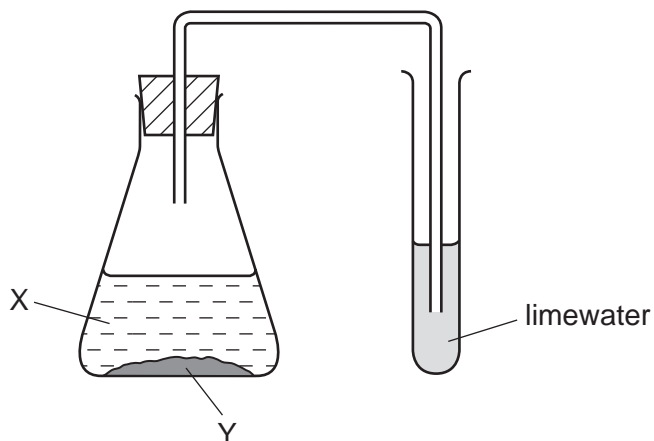
The damp, red litmus paper turns blue.



What are X and Y?

	X	Y
<b>A</b>	aluminium nitrate	hydrochloric acid
<b>B</b>	aluminium nitrate	sodium hydroxide solution
<b>C</b>	ammonium chloride	hydrochloric acid
<b>D</b>	ammonium chloride	sodium hydroxide solution

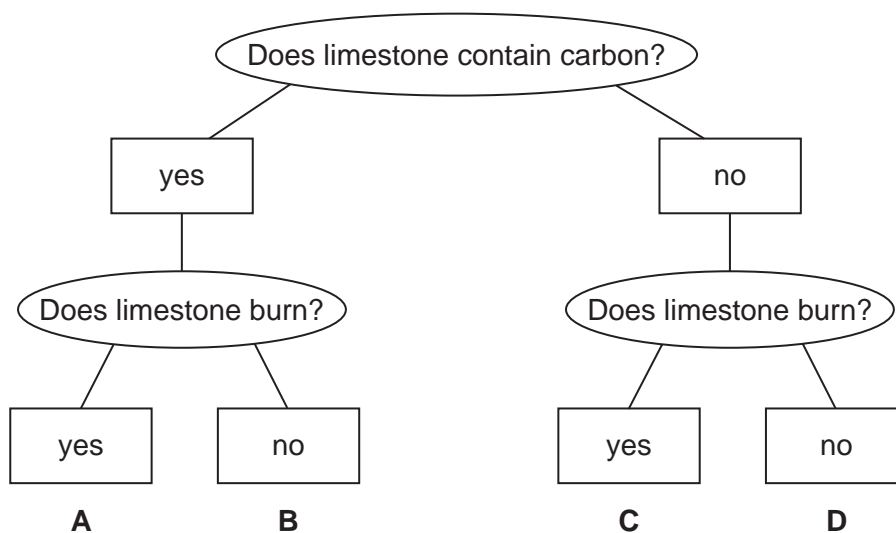
32 In the experiment shown, a white precipitate forms in the limewater.



What are X and Y?

	X	Y
<b>A</b>	aqueous sodium hydroxide	zinc
<b>B</b>	aqueous sodium hydroxide	zinc carbonate
<b>C</b>	dilute sulfuric acid	zinc
<b>D</b>	dilute sulfuric acid	zinc carbonate

33 Which box corresponds to limestone?

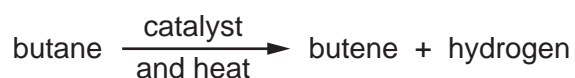


34 Petroleum is an important fossil fuel.

Which row correctly describes petroleum?

	type of substance	composition
<b>A</b>	compound	mainly hydrocarbons
<b>B</b>	compound	only hydrogen and carbon
<b>C</b>	mixture	mainly hydrocarbons
<b>D</b>	mixture	only hydrogen and carbon

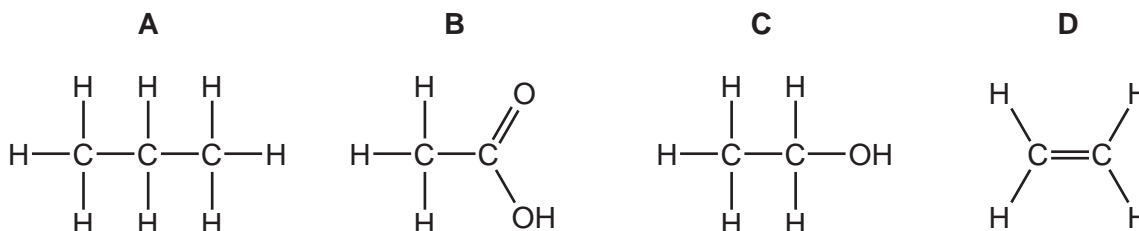
35 Butane reacts as shown.



What is this type of reaction?

- A** combustion
- B** cracking
- C** polymerisation
- D** reduction

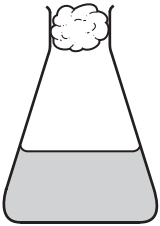
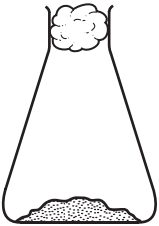
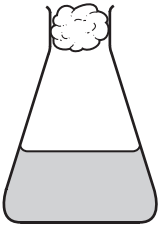
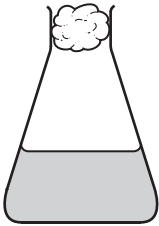
36 Which substance is in the same homologous series as methanol?



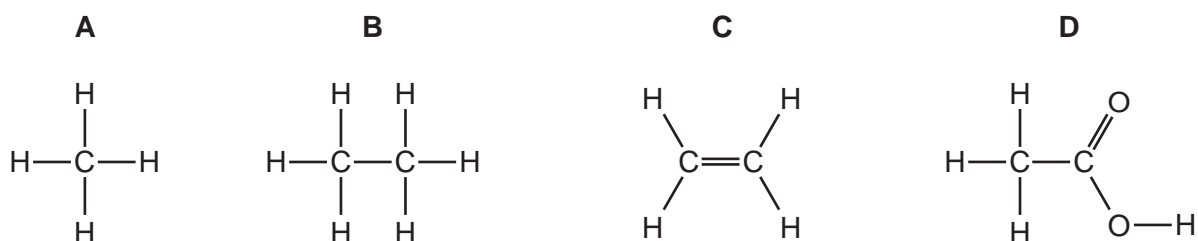
37 Which statement could **not** be correct for an alkane?

- A** It burns readily in a plentiful supply of air to form only carbon dioxide and water.
- B** It decolourises aqueous bromine.
- C** It has a boiling point of  $-42^\circ\text{C}$ .
- D** The carbon and hydrogen atoms in the molecule are joined by sharing pairs of electrons.

38 In which conical flask will ethanol be produced?

<b>A</b>	<b>B</b>	<b>C</b>	<b>D</b>
			
water and sugar	sugar and yeast	water, sugar and yeast	water and yeast

39 Which molecule can be polymerised?



40 Which row describes what happens when ethanol burns in air?

	a white powder is left	heat energy is given out	carbon dioxide is formed	water is formed
<b>A</b>	✓	x	✓	✓
<b>B</b>	x	✓	✓	✓
<b>C</b>	x	✓	✓	x
<b>D</b>	x	✓	x	✓



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## The Periodic Table of Elements

Group																																														
I	II	III						IV	V	VI	VII	VIII																																		
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	<div style="border: 1px solid black; padding: 5px; display: inline-block;"> <b>Key</b>            atomic number            atomic symbol            name            relative atomic mass         </div>										2 <b>He</b> helium 4																																		
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24											5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20																													
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	13 <b>Al</b> aluminium 27	14 <b>Si</b> silicon 28	15 <b>P</b> phosphorus 31	16 <b>S</b> sulfur 32	17 <b>Cl</b> chlorine 35.5	18 <b>Ar</b> argon 40	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84																																			
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	54 <b>Xe</b> xenon 131																																		
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	29 <b>Cu</b> copper 64	28 <b>Ni</b> nickel 59	27 <b>Co</b> cobalt 59	26 <b>Fe</b> iron 56	25 <b>Mn</b> manganese 55	24 <b>Cr</b> chromium 52	30 <b>Zn</b> zinc 65	47 <b>Cu</b> copper 64	46 <b>Ni</b> nickel 59	45 <b>Co</b> cobalt 59	44 <b>Fe</b> iron 56	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —																											
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84	37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131											
89–103 actinoids	88 <b>Ra</b> radium —	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —	87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —												
57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —	87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

lanthanoids	57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
actinoids	89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.)