

WJEC GCSE Flashcards

Component 2: Specific Practical 6B Electrolysis of Aqueous Solutions

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How is a current able to flow in the solution?



How is a current able to flow in the solution?

When in solution, metallic ions dissociate so that they can flow around the solution.

This enables them to act in a similar way to electrons, carrying a flow of charge through the solution from one electrode to the other.



What is the anode?



What is the anode?

The positively charged electrode that loses electrons during electrolysis reactions.



What is the cathode?



What is the cathode?

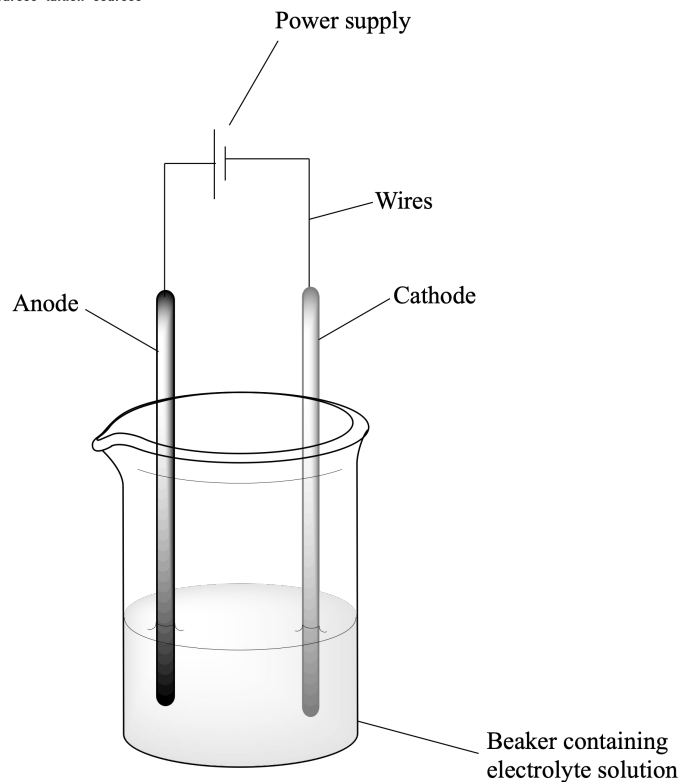
The negatively charged electrode that gains electrons during electrolysis reactions.



Draw a diagram of
electrolysis apparatus.



Draw a diagram of electrolysis equipment.



Why are the electrodes commonly made of carbon?



Why are the electrodes commonly made of carbon?

Carbon electrodes are inert, meaning they won't react and interfere with the electrolysis reaction taking place.



What is the electrolyte?



What is the electrolyte?

The solution between the electrodes that conducts electricity.

