Definitions and Concepts for Edexcel Chemistry GCSE

Topic 6 - Groups in the Periodic Table

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*Definitions in **bold** are for higher tier only*

Definitions marked by '' are for separate sciences only*

Definitions have been taken, or modified from the Edexcel Specification for GCSE Chemistry, 1CH0, Issue 3, February 2018

Alkali metals: The elements in Group 1 of the periodic table. They are typically soft and have relatively low melting points.

Displacement: A chemical reaction in which a more reactive element displaces a less reactive element from its compound.

Halides: The ions formed by halogen atoms when they gain an electron. They have a 1- charge. E.g. Cl^-, Br^- and I^-.

Halogens: The elements in Group 7 of the periodic table. The halogens gain an electron to form halide ions with a 1- charge. Down the group the halogens get more reactive and have higher melting and boiling points.

Inert: Unreactive. Noble gases are inert due to their stable electron configuration.

Noble gases: The elements in Group 0 of the periodic table. They have a stable full outer shell of electrons which makes them very unreactive.

Oxidation: A reaction involving the gain of oxygen. **Oxidation is the loss of electrons.**

Reactivity series: A series in which metals are arranged in order of their reactivity. This can be used to predict products from reactions.

Redox reaction: **A reaction in which both oxidation and reduction occur simultaneously.**

Reduction: A reaction involving the loss of oxygen. **Reduction is the gain of electrons.**