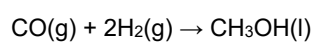


# Enthalpy and Entropy (MCQ)

1. The table below shows standard entropies,  $S^\ominus$

Substance	CO(g)	H <sub>2</sub> (g)	CH <sub>3</sub> OH(l)
$S^\ominus / \text{J mol}^{-1} \text{K}^{-1}$	197.6	130.6	239.7

What is the entropy change,  $\Delta S^\ominus$ , in  $\text{J mol}^{-1} \text{K}^{-1}$ , for the following reaction?



- A -219.1
- B -88.5
- C +88.5
- D +219.1

Your answer

[1]

**END OF QUESTION PAPER**

# Mark scheme – Enthalpy and Entropy (MCQ)

Question			Answer/Indicative content	Marks	Guidance
1			A	1 (AO 1.2)	<b><u>Examiner's Comments</u></b> The correct answer to this was known by the more able candidates. Lower ability candidates struggled.
			<b>Total</b>	<b>1</b>	