

- (c) The pressure was increased whilst keeping the temperature constant. The mixture was left to reach equilibrium.

The equilibrium position above shifted to the right.

- (i) Explain why the equilibrium position shifted to the right.

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[1]

- (ii) What is the effect, if any, on the value of K_c ?

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[1]

- (d) The temperature was increased whilst keeping the pressure constant. The mixture was left to reach equilibrium.

The value of K_c for the equilibrium above decreased.

- (i) Explain what happened to the equilibrium position in the equilibrium.

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[1]

- (ii) Deduce the sign of the enthalpy change for the forward reaction shown in the equilibrium above.

Explain your reasoning.

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[1]

(e) Methanol can be used as an additive to petrol.

(i) Write an equation for the complete combustion of methanol, CH₃OH.

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[1]

(ii) Suggest why methanol is added to petrol.

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[1]

[Total 13 marks]

2. When heated, phosphorus pentachloride, PCl₅, dissociates.



A chemist placed a mixture of the three gases into a container. The initial concentration of each gas was the same: 0.30 mol dm⁻³. The container was left until equilibrium had been reached.

Under these conditions, $K_c = 0.245 \text{ mol dm}^{-3}$.

(a) Write an expression for K_c for this equilibrium.

[1]

(b) Use the value of K_c for this equilibrium to deduce whether the concentration of each gas increases, decreases or stays the same as the mixture approaches equilibrium.

(i) Show your answer by placing a tick in the appropriate cells in the table below.

	initial concentration / mol dm ⁻³	greater than 0.30 mol dm ⁻³	less than 0.30 mol dm ⁻³	equal to 0.30 mol dm ⁻³
PCl ₅	0.30			
PCl ₃	0.30			
Cl ₂	0.30			

[1]

(ii) Explain your deduction.

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[1]

(c) The chemist compressed the equilibrium mixture at constant temperature and allowed it to reach equilibrium under these new conditions.

(i) Explain what happens to the value of K_c .

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[1]

(ii) Explain what happened to the composition of the equilibrium mixture.

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[2]

(d) The chemist heated the equilibrium mixture and the equilibrium moved to the left.

(i) Explain what happens to the value of K_c .

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[1]

(ii) Explain what additional information this observation reveals about the reaction.

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[2]

[Total 9 marks]

3. Some ammonia plants are run at 200–300 atm and 500 °C, with an iron catalyst.

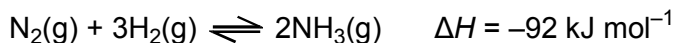
(a) The hydrogen for the plants is obtained by reacting methane with steam.

Construct a possible equation for this reaction.

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[1]

(b) Nitrogen gas and hydrogen gas produce ammonia gas as shown below.



(i) Write the expression for K_c for this equilibrium.

[1]

(ii) At 500 °C, $K_c = 8.00 \times 10^{-2} \text{ dm}^6 \text{ mol}^{-2}$.

At equilibrium, the concentration of N_2 is 1.20 mol dm^{-3} and the concentration of H_2 is 2.00 mol dm^{-3} .

Calculate the equilibrium concentration of ammonia under these conditions.

equilibrium concentration of $\text{NH}_3 = \dots\dots\dots \text{ mol dm}^{-3}$

[3]

(c) In this question one mark is available for the quality of the use and organisation of scientific terms.

Discuss the advantages and disadvantages of running this reaction

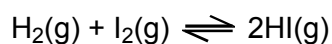
- at a pressure of 200–300 atm;
- at a temperature of 500°C;
- with an iron catalyst.

[6]

Quality of Written Communication [1]

[Total 12 marks]

4. The preparation of hydrogen iodide, HI(g), from hydrogen and iodine gases is a reversible reaction which reaches equilibrium at constant temperature.



- (a) Write the expression for K_c for this equilibrium.

[1]

- (b) A student mixed together 0.30 mol $\text{H}_2(\text{g})$ with 0.20 mol $\text{I}_2(\text{g})$ and the mixture was allowed to reach equilibrium. At equilibrium, 0.14 mol $\text{H}_2(\text{g})$ was present.

- (i) Complete the table below to show the amount of each component in the equilibrium mixture.

component	$\text{H}_2(\text{g})$	$\text{I}_2(\text{g})$	$\text{HI}(\text{g})$
initial amount / mol	0.30	0.20	0
equilibrium amount / mol			

[2]

- (ii) Calculate K_c to an appropriate number of significant figures. State the units, if any.

$K_c = \dots\dots\dots$

units, if any $\dots\dots\dots$

[3]

- (c) The student compressed the equilibrium mixture so that its volume was reduced. The temperature was kept constant.

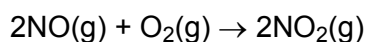
Comment on the value of K_c **and** the composition of the equilibrium mixture under these new conditions.

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[2]

[Total 8 marks]

5. Nitrogen dioxide is one of the major pollutants in air, formed by reaction of nitrogen monoxide with oxygen.



- (a) What is meant by the *rate of reaction*?

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[1]

- (b) A series of experiments was carried out to investigate the kinetics of this reaction. The results are shown in the table below.

Experiment	$[\text{O}_2]$ / mol dm ⁻³	$[\text{NO}]$ / mol dm ⁻³	initial rate / mol dm ⁻³ s ⁻¹
1	0.00100	0.00100	7.10
2	0.00400	0.00100	28.4
3	0.00400	0.00300	256

(i) For each reactant, deduce the order of reaction. Show your reasoning.

O₂(g)

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NO(g)

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[4]

(ii) Deduce the rate equation for this reaction.

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[1]

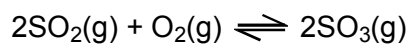
(iii) Calculate the rate constant, k , for this reaction. State the units for k .

$k =$ units

[2]

[Total 8 marks]

6. In the UK, almost all the sulphuric acid, H_2SO_4 , is manufactured by the Contact process.
One stage in the Contact process involves the reaction between sulphur dioxide and oxygen.



The table below shows values of the equilibrium constant, K_p , for this equilibrium at different temperatures.

temperature / °C	K_p / kPa^{-1}
25	4.0×10^{22}
200	2.5×10^8
800	1.3×10^{-3}

- (a) Write an expression for the equilibrium constant, K_p , of this reaction.

[2]

(c) An equilibrium is set up for the SO_2 , O_2 , SO_3 equilibrium at $400\text{ }^\circ\text{C}$.

At this temperature

- the equilibrium partial pressure of SO_2 is 10 kPa
- the equilibrium partial pressure of O_2 is 50 kPa
- $K_p = 3.0 \times 10^2 \text{ kPa}^{-1}$.

Calculate the equilibrium partial pressure of SO_3 at $400\text{ }^\circ\text{C}$. Hence determine the percentage of SO_3 in the equilibrium mixture at this temperature.

answer%

[3]

(d) In the UK, almost all the sulphuric acid manufactured uses sulphur as a starting material for SO_2 production. In some countries, metal ores such as zinc sulphide, ZnS , are used instead to form SO_2 by heating with air.

(i) Construct a balanced equation to show the reaction that takes place when zinc sulphide is heated in air.

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[2]

(ii) Suggest why countries may find it more economic to manufacture sulphuric acid from zinc sulphide.

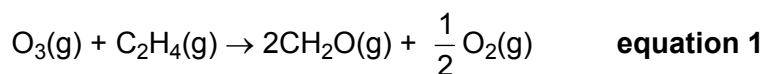
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[1]

[Total 15 marks]

7. One cause of low-level smog is the reaction of ozone, O₃, with ethene. The smog contains methanal, CH₂O(g), and the equation for its production is shown below.



- (a) The rate of the reaction doubles when the initial concentration of either O₃(g) or C₂H₄(g) is doubled.

- (i) What is the order of reaction with respect to

O₃

C₂H₄?

[1]

- (ii) What is the overall order of the reaction?

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[1]

- (iii) Write the rate equation for this reaction.

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[1]

- (b) For an initial concentration of ozone of $0.50 \times 10^{-7} \text{ mol dm}^{-3}$ and one of ethene of $1.0 \times 10^{-8} \text{ mol dm}^{-3}$, the initial rate of methanal formation was $1.0 \times 10^{-12} \text{ mol dm}^{-3} \text{ s}^{-1}$.

- (i) How could the **initial** rate of methanal formation be measured from a concentration/time graph?

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[2]

- (ii) Calculate the value of the rate constant and state the units.

rate constant = units.....

[3]

- (iii) The initial rate of methanal formation is different from that of oxygen formation in **equation 1**.

Explain why.

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[1]

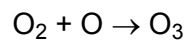
- (iv) The experiment was repeated but at a higher temperature. What would be the effect of this change on the rate and the rate constant of the reaction?

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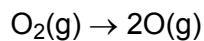
[2]

[Total 11 marks]

8. In the stratosphere, ozone forms when oxygen free radicals react with oxygen molecules.



The oxygen free radicals are initially formed as diradicals when oxygen gas, O_2 , is dissociated by strong ultraviolet radiation,



- (i) Suggest why oxygen free radicals, O, are often called **diradicals**.

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[1]

- (ii) Draw a 'dot-and-cross' diagram of an ozone molecule. Show outer electrons only.

[2]

- (iii) Chlorine free radicals formed from CFCs deplete the ozone layer in a chain reaction.

Typically, 1 g of chlorine free radicals destroys 150 kg of ozone during the atmospheric lifetime of the chlorine free radical (one to two years).

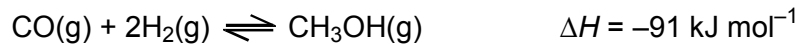
Calculate how many ozone molecules are destroyed in this chain reaction by a single chlorine free radical before the free radical is destroyed.

answer.....

[3]

[Total 6 marks]

9. Methanol, CH₃OH(g), is manufactured from carbon monoxide and hydrogen in an equilibrium reaction.



(a) In this question, one mark is available for the quality of use and organisation of scientific terms.

Explain the advantages and disadvantages of running this reaction

- at a high pressure,
- at a high temperature.

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[5]

Quality of Written Communication [1]

(b) This equilibrium reaction is normally carried out at 10MPa pressure and 550 K, and starting with a 1 : 2 CO : H₂ mixture. At equilibrium, only 10% of the CO has reacted.

(i) Deduce the equilibrium amounts, mole fractions and partial pressures of CO, H₂ and CH₃OH present at equilibrium. Write your answers in the table below.

Assume that you have started with a mixture of 1.0 mol CO and 2.0 mol H₂.

	CO	H ₂	CH ₃ OH
initial amount /mol	1.0	2.0	0.0
equilibrium amount /mol	0.9		
mole fraction at equilibrium			
partial pressure at equilibrium /MPa			

[4]

(ii) Write the expression for K_p for this equilibrium.

[2]

- (iii) The CO : H₂ ratio in the starting mixture was changed from 1 : 2 to 1 : 3 and the mixture was allowed to reach equilibrium at the same temperature and pressure.

Explain, in terms of K_p , the effect of this change on the equilibrium yield of CH₃OH.

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[3]

- (iv) In another experiment, the equilibrium partial pressures were:

CO, 3.70 MPa; H₂, 5.10 MPa; CH₃OH, 0.261 MPa.

Calculate the value of K_p for this equilibrium. Express your answer to an appropriate number of significant figures. State the units of K_p .

$K_p = \dots\dots\dots$ units.....

[2]

- (c) In the UK, the annual production of methanol is 500 000 tonnes. Methanol has many uses in fuels as a reliable and low pollution form of energy.

Suggest an equation for the combustion of methanol.

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[1]

[Total 18 marks]