

Question Number	Acceptable Answer	Additional Guidance	Mark								
* contd	<p>The following table shows how the marks should be awarded for structure and lines of reasoning.</p> <table border="1"> <thead> <tr> <th></th> <th>Number of marks awarded for structure of answer and sustained line of reasoning</th> </tr> </thead> <tbody> <tr> <td>Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.</td> <td>2</td> </tr> <tr> <td>Answer is partially structured with some linkages and lines of reasoning.</td> <td>1</td> </tr> <tr> <td>Answer has no linkages between points and is unstructured.</td> <td>0</td> </tr> </tbody> </table>		Number of marks awarded for structure of answer and sustained line of reasoning	Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2	Answer is partially structured with some linkages and lines of reasoning.	1	Answer has no linkages between points and is unstructured.	0	<p>In general it would be expected that 5 or 6 indicative points would get 2 reasoning marks, and 3 or 4 indicative points would get 1 mark for reasoning, and 0, 1 or 2 indicative points would score zero marks for reasoning.</p> <p>Reasoning marks may be reduced for extra incorrect chemistry</p>	(6)
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	<p>Indicative content:</p> <ul style="list-style-type: none"> aluminium and chlorine electronegativity difference 1.5 AND aluminium and fluorine electronegativity difference 2.5 aluminium chloride (mostly) covalent / (small) molecule aluminium fluoride (bonds) more polar aluminium chloride molecular so weak(er) intermolecular forces / London forces aluminium fluoride is a giant structure/ strong electrostatic forces of attraction between the ions more energy needed to break the stronger bonds to cause sublimation in aluminium fluoride 	<p>Allow all 3 electronegativity values / difference between F and Cl is 1 / difference between differences is 1/ F is 4, CL is 3 and this is a significant difference</p> <p>Allow mostly/more ionic</p> <p>Allow weak(er) dipole-dipole interactions Do not award any suggestion of breaking covalent bonds</p> <p>Allow stronger dipole-dipole attractions</p> <p>Allow (dative) covalent bonds breaking (to form small molecule / AlF_3)</p>	
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