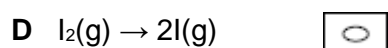
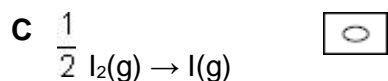
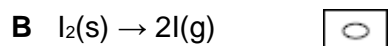
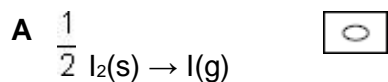


**Q1.**

Which equation represents the process that occurs when the standard enthalpy of atomisation of iodine is measured?



(Total 1 mark)

**Q2.**

Lattice enthalpy values can be obtained from Born–Haber cycles and by calculations based on a perfect ionic model.

Which compound shows the greatest percentage difference between these two values?



(Total 1 mark)

**Q3.**

A reaction is exothermic and has a negative entropy change.

Which statement is correct?

A The reaction is always feasible

B The reaction is feasible above a certain temperature

C The reaction is feasible below a certain temperature

D The reaction is never feasible

(Total 1 mark)