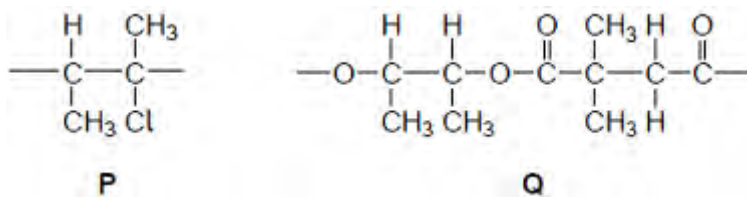


**Q1.** Repeating units of two polymers, **P** and **Q**, are shown in the figure below.



- (a) Draw the structure of the monomer used to form polymer **P**.  
Name the type of polymerisation involved.

Monomer

Type of polymerisation .....

(2)

- (b) Draw the structures of **two** compounds that react together to form polymer **Q**.

Structure of compound 1

Structure of compound 2

(2)

- (c) Suggest an environmental advantage of polymer **Q** over polymer **P**.  
Justify your answer.

Advantage .....

Justification .....

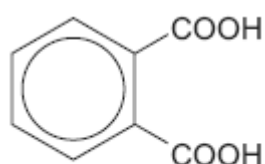
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.....  
.....  
.....

(3)  
(Total 7 marks)

**Q2.** Items softened with plasticisers have become an essential part of our modern society.

Compound **S**, shown below, is commonly known as phthalic acid.

Esters of phthalic acid are called phthalates and are used as plasticisers to soften polymers such as PVC, poly(chloroethene).



**S**

- (a) Give the IUPAC name for phthalic acid.

.....

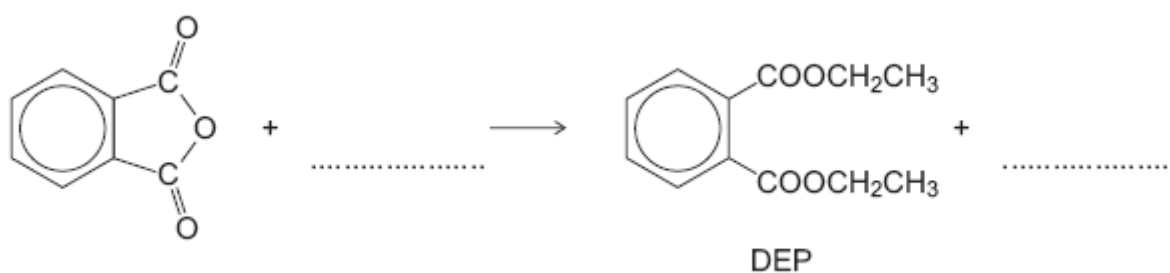
(1)

- (b) Draw the displayed formula of the repeating unit of poly(chloroethene).

(1)

(c) The ester diethyl phthalate (DEP) is used in food packaging and in cosmetics.

(i) Complete the following equation showing the formation of DEP from phthalic anhydride.



(2)

(ii) Deduce the number of peaks in the  $^{13}\text{C}$  n.m.r. spectrum of DEP.

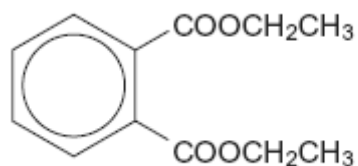
.....

(1)

(iii) One of the peaks in the  $^{13}\text{C}$  n.m.r. spectrum of DEP is at  $\delta = 62$  ppm.

**Table 3** on the Data Sheet can be used to identify a type of carbon atom responsible for this peak.

Draw a circle around **one** carbon atom of this type in the structure below.



(1)

(d) The mass spectrum of DEP includes major peaks at  $m/z = 222$  (the molecular ion) and at  $m/z = 177$

Write an equation to show the fragmentation of the molecular ion to form the fragment that causes the peak at  $m/z = 177$

.....

(2)

- (e) Because of their many uses, phthalates have been tested for possible adverse effects to humans and to the environment.

An organisation that represents the manufacturers of plasticisers asserts that experimental evidence and research findings show that phthalates do not pose a risk to human health because they biodegrade in a short time scale.

According to the organization's research, phthalates do not represent a risk for humans or for the environment and they are biodegradable.

- (i) Hydrolysis of DEP in an excess of water was found to follow first order kinetics.

Write a rate equation for this hydrolysis reaction using DEP to represent the ester.

.....

(1)

- (ii) Suggest what needs to be done so that the public could feel confident that the research discussed above is reliable.

.....

.....

.....

.....

.....

(Extra space) .....

.....

.....

(2)

**Q3.** Lactic acid,  $\text{CH}_3\text{CH}(\text{OH})\text{COOH}$ , is formed in the human body during metabolism and exercise. This acid is also formed by the fermentation of carbohydrates such as sucrose,  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ .

- (a) (i) Give the IUPAC name for lactic acid.

.....

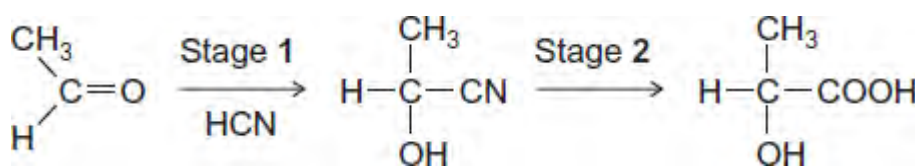
(1)

- (ii) Write an equation for the formation of lactic acid from sucrose and water.

.....

(1)

- (b) A molecule of lactic acid contains an asymmetric carbon atom. The lactic acid in the body occurs as a single enantiomer. A racemic mixture (racemate) of lactic acid can be formed in the following two-stage synthesis.



- (i) Name and outline a mechanism for Stage 1.

Name of mechanism .....

Mechanism

(5)

(ii) Give the meaning of the term *racemic mixture (racemate)*.

.....  
.....  
.....

(1)

(iii) Explain how you could distinguish between a racemic mixture (racemate) of lactic acid and one of the enantiomers of lactic acid.

.....  
.....  
.....  
.....

(2)

(c) A mixture of lactic acid and its salt sodium lactate is used as an acidity regulator in some foods. An acidity regulator makes sure that there is little variation in the pH of food.

(i) Write an equation for the reaction of lactic acid with sodium hydroxide.

.....

(1)

(ii) The acid dissociation constant  $K_a$  for lactic acid has the value  $1.38 \times 10^{-4} \text{ mol dm}^{-3}$  at 298 K.

Calculate the pH of an equimolar solution of lactic acid and sodium lactate.

.....  
.....  
.....  
.....

(2)

- (iii) Suggest an alternative name for the term *acidity regulator*.  
Explain how a mixture of lactic acid and sodium lactate can act as a regulator when natural processes increase the acidity in some foods.

Name .....

Explanation .....

.....

.....

.....

(Extra space) .....

.....

(3)

(d)



© www.biopac.co.uk The cup shown is made from PLA, poly(lactic acid).  
PLA is the condensation polymer formed from lactic acid.

The polymer is described as 100% biodegradable and 100% compostable.

Compostable material breaks down slowly in contact with the moist air in a garden bin.  
This produces compost that can be used to improve soil.

The manufacturers stress that PLA cups differ from traditional plastic cups that are  
neither biodegradable nor compostable.

- (i) Draw a section of PLA that shows **two** repeating units.

(2)

(ii) Name the type of condensation polymer in PLA.

.....

(1)

(iii) An intermediate in the production of PLA is a cyclic compound ( $C_6H_8O_4$ ) that is formed from two PLA molecules.

Draw the structure of this cyclic compound.

(1)

(iv) Traditional non-biodegradable plastic cups can be made from poly(phenylethene), commonly known as *polystyrene*.

Draw the repeating unit of poly(phenylethene).

(1)



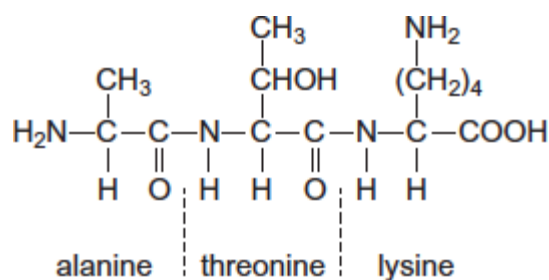
- (v) The manufacturers of PLA claim that the material will break down to compost in just 12 weeks.

Suggest **one** reason why PLA in landfill may take longer than 12 weeks to break down.

.....  
 .....

(1)  
 (Total 22 marks)

- Q4.(a)** The tripeptide shown is formed from the amino acids alanine, threonine and lysine.



- (i) Draw a separate circle around **each** of the asymmetric carbon atoms in the tripeptide.

(1)

- (ii) Draw the zwitterion of alanine.

(1)

- (iii) Give the IUPAC name of threonine.

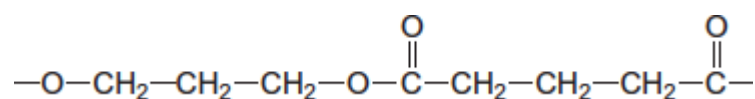
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(1)

- (iv) Draw the species formed by lysine at low pH.

(1)

- (b) The repeating unit shown represents a polyester.



- (i) Name this type of polymer.

.....

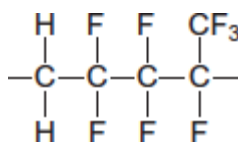
(1)

- (ii) Give the IUPAC name for the alcohol used to prepare this polyester.

.....

(1)

- (c) The repeating unit shown represents a polyalkene co-polymer. This co-polymer is made from two different alkene monomers.



- (i) Name the type of polymerisation occurring in the formation of this co-polymer.

.....

(1)

- (ii) Draw the structure of each alkene monomer.

Alkene monomer 1

Alkene monomer 2

(2)

- (d) One of the three compounds shown in parts (a), (b) and (c) cannot be broken down by hydrolysis.

Write the letter **(a)**, **(b)** or **(c)** to identify this compound and explain why hydrolysis of this compound does **not** occur.

Compound .....

Explanation .....

.....

.....

(2)  
(Total 11 marks)