



GCE

Chemistry B

H033/02: Chemistry in depth

Advanced Subsidiary GCE

Mark Scheme for June 2019

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.















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H033/02

Mark Scheme

June 2019

Annotations available in RM Assessor

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore

H033/02

Mark Scheme

June 2019

Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

Subject-specific Marking Instructions

INTRODUCTION

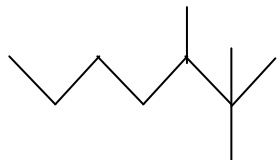
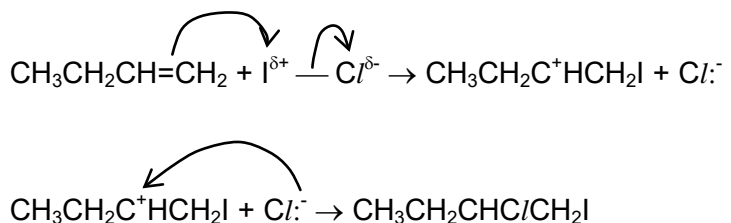
Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

You should ensure that you have copies of these materials.

You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

Please ask for help or guidance whenever you need it. Your first point of contact is your Team Leader.

Question			Answer	Mark	AO element	Guidance
1	(a)	(i)	$C_{10}H_{22} \rightarrow C_4H_8 + C_6H_{14}$ ✓	1	1.2	Molecular formulae of both products required
1	(a)	(ii)	 ✓	1	2.5	IGNORE non-skeletal representation used as working and not rubbed out if a correct skeletal representation is present
1	(b)		Collecting tube must be full of water ✓ (otherwise) gaseous products/gas will not collect in tube ✓ Bung is required to hold delivery tube in heated tube ✓ (otherwise) gaseous products/gas will escape (into atmosphere) ✓	4	3.3 3.4 3.3 3.4	ALLOW answers drawn and/or written onto the diagram
1	(c)		Propene ✓	1	2.1	ALLOW propylene
1	(d)		There is a lack of free rotation about the C=C ✓ The groups on each carbon atom (of the double bond) are different ✓	2	1.1 x 2	ALLOW no rotation DO NOT ALLOW 'movement' unless qualified by 'rotational' IGNORE 'E/Z'
1	(e)		 $CH_3CH_2CH=CH_2 + I^{\delta+} - Cl^{\delta-} \rightarrow CH_3CH_2C^+HCH_2I + Cl:^-$ $CH_3CH_2C^+HCH_2I + Cl:^- \rightarrow CH_3CH_2CHClCH_2I$ ✓ for partial charges and curly arrow on ICl ✓ for full charges ✓ for other two curly arrows	3	1.2 x 3	Curly arrows should start (when projected back if necessary) on the relevant bond (or the minus charge or lone pair of Cl^-) and end (when projected forward if necessary) on the atom concerned or the bond about to be formed. ALLOW $Cl:^-$ with or without lone pair in step 1

H033 02

Mark Scheme

June 2019

Question		Answer	Mark	AO element	Guidance	
1	(f)	<p>FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = 23.3 (dm³) to 3 or more s.f. (23.2642 ...) award 4 marks</p> <p>M_r of C₈H₁₈ = 114 n of C₈H₁₈ = (3.42/114) = 0.03 mol n of CO₂ and H₂O = (17 x 0.03) = 0.51 mol (mark is for 0.51 – working not required) ✓</p> <p>$pV = nRT$ $V = nRT/p$ (mark is for changing subject of equation – can be implied by substitution of numbers, even before T conversion) ✓</p> <p>$V = 0.51 \times 8.314 \times (550+273)/150000$ (mark is for T conversion) ✓</p> <p>$V = 2.33 \times 10^{-2} \text{ m}^3$ $V = 23.3 \text{ dm}^3$ (mark is for V conversion from m³ to dm³) ✓</p>	4		<p>ALLOW 2 or more sf ALLOW ECF</p> <p>2.6</p> <p>2.6</p> <p>2.6</p> <p>2.2</p>	
1	(g)	(i)	Identity of compound: Ethanol/C ₂ H ₅ OH/CH ₃ CH ₂ OH ✓	1	3.1	ALLOW displayed formula IGNORE molecular formula of C ₂ H ₆ O
		(ii)	Evidence from mass spectrum: Molecular ion peak at $m/z = 46$ OR $M_r = 46$ ✓	1	3.2	
		(iii)	Evidence from IR spectrum: O-H at 3350 (cm ⁻¹) ✓	1	3.2	Evidence from IR spectrum must include both bond and wavenumber ALLOW reference to the range 3200 - 3600

H033 02

Mark Scheme

June 2019

1	*(h)	<p><i>Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question.</i></p> <p>Level 3 (5-6 marks) Gives a balanced and <u>detailed</u> account of the adv and disadv of biofuels <i>(detailed means more than a simple statement).</i></p> <p><i>There is a well-developed line of reasoning which is clear and logically structured.</i></p> <p>Level 2 (3-4 marks) Gives a balanced account of the advantages and disadvantages of biofuels. OR Gives a detailed account of the advantages and outlines the disadvantages of biofuels, or vice versa.</p> <p><i>There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.</i></p> <p>Level 1 (1-2 marks) Gives some of the advantages of biofuels. OR Gives some of the disadvantages of biofuels.</p> <p><i>There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.</i></p> <p>0 marks No response or no response worthy of credit.</p>	6	1.2 x 3 2.5 x 3	<p><i>Indicative Scientific points include:</i></p> <p>Advantages (compared with fossil fuels)</p> <ul style="list-style-type: none"> • renewable • the CO₂ they produce when burnt is the CO₂ the plants absorbed while growing/they are carbon neutral/less greenhouse gases/less global warming • can be produced from waste (that would otherwise go to landfill) • sustainable • can be used as a straight replacement for diesel • (biodiesel is) biodegradable (if spilled) • (virtually) no sulfur/less SO₂/less acid rain • less particulates/less damage to lungs • less CO/less poisonous emissions • less HC emissions/less photochemical smog/less respiratory problems <p><i>(Ignore references to cost)</i></p> <p>Disadvantages (compared with fossil fuels)</p> <ul style="list-style-type: none"> • not a straight replacement for petrol/car engines need to be modified (to use fuels of high ethanol concentration) • land used to grow crops for fuels can't be used to grow crops for food • not carbon neutral as CO₂ is still produced when producing and transporting the fuel • more NO_x • (NO_x causes, so) more tropospheric ozone/more respiratory problems <p><i>(Ignore references to cost)</i></p>
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H033 02

Mark Scheme

June 2019

Question		Answer	Mark	AO element	Guidance
2	(a)	anode (+) $2Cl^- \rightarrow Cl_2 + 2e^-$ ✓ cathode (-) $2H_2O + 2e^- \rightarrow 2OH^- + H_2$ ✓	2	2.3 x 2	ALLOW multiples or halves of equations ALLOW $2Cl^- - 2e^- \rightarrow Cl_2$ ALLOW $2H^+ + 2e^- \rightarrow H_2$ ALLOW 'e' for 'e ⁻ ' IGNORE state symbols
2	(b)	(i) It is the oxidation state/number of the chlorine/Cl (in chloric(I) acid)/the OS/N of chlorine/Cl is +1 ✓	1	1.2	Both oxidation state and species required If OS/N is given it must be +1 NOT 1+
2	(b)	(ii) It is easier to transport/store/handle/less toxic (ORA) ✓ because it is a solid and chlorine is a gas ✓	2	2.3 x 2	The reason must be linked to the state and compared with the chlorine
2	(c)	(i) FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = 2.48×10^{-5} (mol dm⁻³) to 3 s.f. award 4 marks Amount $S_2O_3^{2-}$ = $(12.4 / 1000 \times 0.000100) = 1.24 \times 10^{-6}$ (mol) ✓ Amount of Cl_2 = $(1.24 \times 10^{-6} / 2) = 6.20 \times 10^{-7}$ (mol) ✓ Concentration of Cl_2 = $(6.2 \times 10^{-7} \times 1000 / 25) = 2.48 \times 10^{-5}$ (mol dm ⁻³) ✓ Appropriate number of s.f. = 3 ✓	4	2.4 x 4	Award fourth mark independently for any number to 3sf that results from a shown calculation. ALLOW ECF

H033 02

Mark Scheme

June 2019

2	(c)	(ii)	<p>Any three from:</p> <p>(in 1) a 25 cm³ measuring cylinder – use a measuring/graduated/volumetric pipette (instead) ✓</p> <p>(in 2) rinse out a burette with de-ionised water – use the sodium thiosulfate/standard solution (instead) ✓</p> <p>(in 3) record the initial burette reading to the nearest 0.1 cm³ – should be to the nearest 0.05/±0.05 (instead) ✓</p> <p>(in 5) add a few drops of starch indicator – (this should not be added before titrating –) add when addition of sodium thiosulfate has caused iodine/brown colour to become (pale) yellow/straw-coloured (instead) ✓</p>	3	3.3x 2 3.4	ALLOW just 'pipette' (in 1)
2	(d)		<p>Reducing agent: SO₂</p> <p>Explanation: the oxidation state (number) of the sulfur/S (in SO₂) increases (goes) from +4 to +6 (in SO₄²⁻) ✓</p>	1	2.6	<p>ALLOW sulfur dioxide</p> <p>ALLOW the oxidation state (number) of the bromine (Br) (in Br₂) decreases (goes) from 0 to -1 (in Br⁻)</p> <p>Both reducing agent and explanation required for mark</p>
2	(e)		(dark) red/brown liquid ✓	1	1.2	ALLOW red/brown/orange but no other colour
2	(f)	(i)	cream(-coloured)/off-white precipitate/ppt/solid ✓	1	1.2	<p>IGNORE initial colours and changes of colour on standing</p> <p>No other colour(s) should be mentioned</p>

H033 02

Mark Scheme

June 2019

2	(f)	(ii)	<p>FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = rounds to 97(%) to 2 or more s.f. award 3 marks</p> <p>amount of Br⁻ in 25.0 cm³ sample = (25.0 / 1000 x 0.0260) = 6.50 x 10⁻⁴ mol amount of AgBr produced = 6.50 x 10⁻⁴ mol ✓</p> <p>M(AgBr) = (107.9 + 79.9) = 187.8 g mol⁻¹ maximum mass of AgBr = (6.50 x 10⁻⁴ x 187.8) = 0.122(07) (g) ✓</p> <p>% yield = [(0.118 / 0.122) x 100] = 97% ✓</p>	3	2.7 x 3	ALLOW answer to 2 or more s.f. (96.7213...) ALLOW ECF
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Question		Answer	Mark	AO element	Guidance
3	(a)	${}^2_1\text{H} + {}^2_1\text{H} \rightarrow {}^4_2\text{He} \checkmark$	1	2.6	ALLOW atomic/mass numbers (only if correct) to be written to rhs of both H/He
3	(b)	The student is correct AW (because) The lines are in exactly the same positions/frequencies/ match up \checkmark The energy transitions are the same \checkmark Each element has unique energy levels AW \checkmark	3	3.1 3.2 3.2	Must say 'student is correct' AW to score three marks. ALLOW energy is quantized
3	(c)	FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = % ${}^{113}\text{In} = 4\%$ ${}^{115}\text{In} = 96\%$ award 2 marks [112.90x] + [114.90(100-x)] / 100 = 114.82 \checkmark 112.90x + 11490 – 114.90x = 11482 -2x = -8 x = 4 % ${}^{113}\text{In} = 4$ % ${}^{115}\text{In} = 96 \checkmark$	2	2.6 x 2	
3	(d)	(i) $5p^1 \checkmark$	1	1.1	
3	(d)	(ii) Indium has/there are delocalised electrons \checkmark	1	1.1	IGNORE it has metallic bonding
3	(d)	(iii) <u>Strong</u> attraction between <u>oppositely</u> charged ions (in a giant structure) \checkmark	1	1.1	ALLOW ' <u>strong</u> attractions in <u>ionic</u> bonding'
3	(e)	(Student is incorrect about InH_3) InH_3 has (only) <u>3 bond pairs</u> (of electrons) \checkmark (equal repulsion between b-p) gives <u>trigonal planar shape/bond angles 120°</u> \checkmark NH_3 has <u>3 b-p and 1 l-p</u> \checkmark	3	3.2 3.2 1.2	

H033 02

Mark Scheme

June 2019

Question			Answer	Mark	AO element	Guidance
4	(a)	(i)	C–F is most polar ✓ (because) F is (the) most <u>electronegative</u> ✓	2	3.1 3.2	ALLOW order of bond polarities C-F>C-C/>C-Br for MP1 ALLOW order of electronegativity F>C/>Br for MP2
	(a)	(ii)	C–Br is broken ✓ (because the C–Br is the) <u>weakest bond</u> AW ✓	2	3.2 x 2	ALLOW order of bond enthalpies/strengths C-F>C-C/>C-Br for MP2
4	(b)		FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = C₂Br₂F₄ award 2 marks molar ratio of C:Br:F 9.2 / 12.0 : 61.5 / 79.9 : 29.3 / 19.0 0.77 0.77 1.54 1 : 1 : 2 OR empirical formula = CBrF ₂ ✓ empirical mass = 129.9 M _r = 259.8 (259.8 / 129.9) = 2 molecular formula = C ₂ Br ₂ F ₄ ✓	2	2.4 2.8	
4	(c)		FIRST CHECK THE ANSWER ON THE ANSWER LINE If answer = (+) 346 (kJ mol⁻¹) (3 or more s.f.) award 3 marks Energy needed to break one C-Cl bond (E = hν) = 6.63 x 10 ⁻³⁴ x 8.67 x 10 ¹⁴ = 5.75 x 10 ⁻¹⁹ (J) ✓ Energy needed to break one mole of C-Br bonds = 5.75 x 10 ⁻¹⁹ x 6.02 x 10 ²³ = 346042 (J) ✓ = (+) 346 kJ mol ⁻¹ ✓	3	2.4	The working for an incorrect answer MUST be checked in detail. Do be aware that candidates may well multiply/divide the numbers in a different order (or even combine steps) to that shown in the answer column so the order of/the numbers in this method of working may not necessarily be seen. However, candidates should show evidence (explicit or implicit) of using E = hν (✓), and multiplying by both the Planck constant and the Avogadro constant (✓) and converting between J and kJ (✓).

Question		Answer	Mark	AO element	Guidance
4	*(d)	<p><i>Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question.</i></p> <p>Level 3 (5-6 marks) Gives <u>details</u> of adv and disadv of O₃ <i>(details means correctly specifies tropos and stratos)</i> AND Describes the <u>catalytic</u> role of NO₂ in O₃ destruction supported by reference to equations AND Refers to NO₂ and O₃ reacting/forming photochemical smog in the tropos</p> <p><i>Well-developed reasoning, clear, logically structured.</i></p> <p>Level 2 (3-4 marks) Gives the adv and disadv of O₃ AND Describes the <u>catalytic</u> role of NO₂ in O₃ destruction supported by some reference to equation(s)</p> <p style="text-align: center;">OR</p> <p>Gives <u>details</u> of the adv and disadv of O₃ AND Describes the disadv of NO₂ in the tropos and/or stratos.</p> <p><i>Reasoning with some structure. Information is relevant and supported by some evidence.</i></p> <p>Level 1 (1-2 marks) Gives a reference to either the adv or disadv of O₃</p>	6	1.2 (x4) 2.7 (x2)	<p><i>Indicative Scientific points include:</i></p> <p>Ozone/O₃ Disadvantages <u>In troposphere</u></p> <ul style="list-style-type: none"> • it is toxic • (it can lead to) photochemical smog/causes respiratory problems <p>Advantages <u>In stratosphere</u></p> <ul style="list-style-type: none"> • it protects against/absorbs (harmful) ultraviolet radiation <i>(relevant equation)</i> $O_3 \rightarrow O_2 + O$ • UV causes sunburn and skin cancer/DNA damage/ damages eyes/affects crops. <p>Effect of nitrogen dioxide/NO₂ (on ozone) <u>In stratosphere:</u></p> <ul style="list-style-type: none"> • It catalyses the breakdown of ozone • it is split into radicals <i>(relevant equation)</i> $NO_2 + hv \rightarrow NO + O$ • (NO) radicals react with O₃ <i>(relevant equations)</i> $NO + O_3 \rightarrow NO_2 + O_2$ $NO_2 + O \rightarrow NO + O_2$ $(O_3 + O \rightarrow 2O_2)$ overall • radical(s) is/are regenerated/catalysts <p><u>In troposphere:</u></p> <ul style="list-style-type: none"> • NO₂ reacts with O₃ to form photochemical smog

H033 02

Mark Scheme

June 2019

Question		Answer	Mark	AO element	Guidance
		<p style="text-align: center;">OR</p> <p>The disadv of NO₂ in the tropos and stratos</p> <p style="text-align: center;">OR</p> <p>A description of the catalytic role of NO₂ in ozone destruction.</p> <p><i>Attempt at logical structure/line of reasoning. Information is in the most part relevant.</i></p> <p>0 marks No response or no response worthy of credit</p>			

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