



Mark Scheme (Results)

Summer 2016

Pearson Edexcel
International Advanced Level
in Chemistry (WCH04) Paper 01
General Principles of Chemistry I

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.
- Mark schemes will indicate within the table where, and which strands of QWC, are being assessed. The strands are as follows:
 - i) ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear
 - ii) select and use a form and style of writing appropriate to purpose and to complex subject matter
 - iii) organise information clearly and coherently, using specialist vocabulary when appropriate

Using the Mark Scheme

Examiners should look for qualities to reward rather than faults to penalise. This does NOT mean giving credit for incorrect or inadequate answers, but it does mean allowing candidates to be rewarded for answers showing correct application of principles and knowledge. Examiners should therefore read carefully and consider every response: even if it is not what is expected it may be worthy of credit.

The mark scheme gives examiners:

- an idea of the types of response expected
- how individual marks are to be awarded
- the total mark for each question
- examples of responses that should NOT receive credit.

/ means that the responses are alternatives and either answer should receive full credit.

() means that a phrase/word is not essential for the award of the mark, but helps the examiner to get the sense of the expected answer.

Phrases/words in **bold** indicate that the meaning of the phrase or the actual word is **essential** to the answer.

TE/TE/cq (error carried forward) means that a wrong answer given in an earlier part of a question is used correctly in answer to a later part of the same question.

Candidates must make their meaning clear to the examiner to gain the mark. Make sure that the answer makes sense. Do not give credit for correct words/phrases which are put together in a meaningless manner. Answers must be in the correct context.

Quality of Written Communication

Questions which involve the writing of continuous prose will expect candidates to:

- write legibly, with accurate use of spelling, grammar and punctuation in order to make the meaning clear
- select and use a form and style of writing appropriate to purpose and to complex subject matter
- organise information clearly and coherently, using specialist vocabulary when appropriate.

Full marks will be awarded if the candidate has demonstrated the above abilities.

Questions where QWC is likely to be particularly important are indicated (QWC) in the mark scheme, but this does not preclude others.

Section A (multiple choice)

Question Number	Correct Answer	Reject	Mark
1	C		(1)

Question Number	Correct Answer	Reject	Mark
2	D		(1)

Question Number	Correct Answer	Reject	Mark
3(a)	B		(1)

Question Number	Correct Answer	Reject	Mark
3(b)	A		(1)

Question Number	Correct Answer	Reject	Mark
4(a)	C		(1)

Question Number	Correct Answer	Reject	Mark
4(b)	C		(1)

Question Number	Correct Answer	Reject	Mark
4(c)	A		(1)

Question Number	Correct Answer	Reject	Mark
4(d)	D		(1)

Question Number	Correct Answer	Reject	Mark
4(e)	D		(1)

Question Number	Correct Answer	Reject	Mark
5	C		(1)

Question Number	Correct Answer	Reject	Mark
6(a)	C		(1)

Question Number	Correct Answer	Reject	Mark
6(b)	A		(1)

Question Number	Correct Answer	Reject	Mark
7(a)	C		(1)

Question Number	Correct Answer	Reject	Mark
7(b)	B		(1)

Question Number	Correct Answer	Reject	Mark
8(a)	D		(1)

Question Number	Correct Answer	Reject	Mark
8(b)	B		(1)

Question Number	Correct Answer	Reject	Mark
9(a)	A		(1)

Question Number	Correct Answer	Reject	Mark
9(b)	D		(1)

Question Number	Correct Answer	Reject	Mark
10(a)	C		(1)

Question Number	Correct Answer	Reject	Mark
10(b)	D		(1)

TOTAL FOR SECTION A = 20 MARKS

Section B

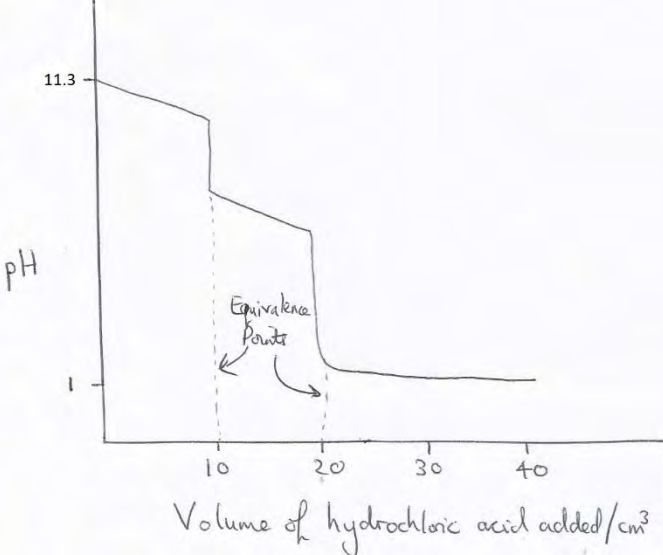
Question Number	Acceptable Answers	Reject	Mark
11(a)	$(K_{a1} =) \frac{[\text{H}_3\text{O}^+(\text{aq})][\text{HCO}_3^-(\text{aq})]}{[\text{H}_2\text{CO}_3(\text{aq})]}$ (1) $(K_{a2} =) \frac{[\text{H}_3\text{O}^+(\text{aq})][\text{CO}_3^{2-}(\text{aq})]}{[\text{HCO}_3^-(\text{aq})]}$ (1) ALLOW $\text{H}^+(\text{aq})$ for $\text{H}_3\text{O}^+(\text{aq})$ IGNORE state symbols, even if incorrect	$[\text{H}_3\text{O}^+]^2$ numerator $[\text{H}_3\text{O}^+]^2$ numerator	(2)

Question Number	Acceptable Answers	Reject	Mark								
11(b)(i)	$\text{H}_2\text{CO}_3(\text{aq}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HCO}_3^-(\text{aq})$ <table style="margin-left: 40px;"> <tr> <td>Initially</td> <td>0.100</td> <td>0</td> <td>0</td> </tr> <tr> <td>At eqm</td> <td>0.100</td> <td>x</td> <td>x</td> </tr> </table> $K_{a1} = \frac{x^2}{0.100}$ <p>M1 $x^2 = 4.17 \times 10^{-8} \text{ (mol}^2 \text{ dm}^{-6}\text{)}$ (1)</p> <p>M2 $(x = 2.0421 \times 10^{-4})$ $[\text{HCO}_3^-] = 2.04 \times 10^{-4} / 0.000204 \text{ (mol dm}^{-3}\text{)}$ (1)</p> <p>Final answer for M2 must be to 3 SF</p> <p>Correct final answer without working scores (2)</p> <p>ALLOW</p> <p>M2 TE on candidate's value for M1, as long as final answer to 3 sf</p> <p>IGNORE units even if incorrect</p>	Initially	0.100	0	0	At eqm	0.100	x	x		(2)
Initially	0.100	0	0								
At eqm	0.100	x	x								

Question Number	Acceptable Answers	Reject	Mark
11(b)(ii)	$(\text{pH} = -\log 2.04 \times 10^{-4} =) 3.69 / 3.7$	$\text{pH} = 4$	(1)

	TE on answer to (b)(i), provided pH < 7 pH = 3.19 / 3.2 from a $[\text{HCO}_3^-]$ value of 6.46×10^{-4} (mol dm ⁻³) ALLOW any SF except 1 SF	pH = 3	
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Question Number	Acceptable Answers	Reject	Mark
*11 (b) (iii)	<p>Max 2 if HA and A⁻ used for H₂CO₃</p> <p>Any THREE from:</p> <p>Assumption 1 $[\text{H}_2\text{CO}_3]_{\text{equilibrium}} = [\text{H}_2\text{CO}_3]_{\text{initial}}$ OR The dissociation of H₂CO₃/the acid is negligible OR 2.04×10^{-4} is (very) small compared to the initial concentration of H₂CO₃/0.100 (hence a valid assumption), or reverse argument</p> <p>Assumption 2 $[\text{H}_3\text{O}^+] = [\text{HCO}_3^-]$ OR $[\text{H}^+] = [\text{HCO}_3^-]$</p> <p>OR Negligible H⁺ from (the dissociation of) water / H⁺ only from H₂CO₃</p> <p>Assumption 3 Negligible dissociation of HCO₃⁻ / HCO₃⁻ doesn't (significantly) dissociate further OR K_{a2} very much smaller than K_{a1}</p> <p>ALLOW</p> <p>Stage 2 does not occur (significantly)</p> <p>Assumption 4 Measurements at 298 K / standard temperature</p> <p>IGNORE</p> <p>References to the concentration of water References just to 'standard conditions'</p>		(3)

Question Number	Acceptable Answers	Reject	Mark
<p>11(c)</p>	 <p>M1 General shape of a weak base-strong acid curve with pH decreasing and either one or two vertical sections shown. Penalise vertical section ≥ 8 pH units (1)</p> <p>M2 (Approximately) vertical section at 10 cm³ (1)</p> <p>M3 (Approximately) vertical section at 20 cm³ (1)</p> <p>M4 Any (mid-point of a) vertical section identified as an equivalence point (1)</p> <p>M5 Initial pH = 11-12 and line becomes asymptotic to pH range 1-2 ALLOW a recognisable plateau in the pH range 1 – 2</p> <p>If curve is the wrong way round the starting pH should be pH 1-2 and the asymptote or plateau should be pH 11-12</p> <p>Curve does not need to reach 40 cm³ (1)</p>	<p>pH rising by half a square or more at any point</p>	<p>(5)</p>

	ALLOW Two vertical sections not at 10/20 cm ³ scores (1) if M2 and M3 not awarded		
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(Total for question 11 = 13 Marks)

Question Number	Acceptable Answers	Reject	Mark
12(a)(i)	Effervescence / bubbles / fizzing IGNORE gas evolved / temperature increase		(1)

Question Number	Acceptable Answers	Reject	Mark
*12(a)(ii)	<p>A statement that entropy is positive needs to be made once only and can be used to award M1 and M2</p> <p>Penalise omission of statement that entropy is positive once only</p> <p>M1 Entropy (of the system) positive and solid and liquid reactants form (a solid, a liquid and) a gas</p> <p>ALLOW gas formed / gas is a product (1)</p> <p>M2 Entropy (of the system) positive and EITHER 3 moles → 4 moles OR more moles of products (than reactants)</p> <p>ALLOW 'molecules' for moles</p> <p>OR More ways of distributing energy OR</p>	<p>If entropy of system is negative / decreases scores (0)</p> <p>particles</p>	(2)

	More ways of distributing quanta (1)		
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Question Number	Acceptable Answers	Reject	Mark
12(b)(i)	$\sum S^{\ominus}_{(\text{reactants})} = (31.8) + 3(2 \times 158.6) =$ $+983.4 \text{ J mol}^{-1} \text{ K}^{-1}$ <p>(1)</p> $\Delta S^{\ominus}_{\text{system}} = (291.7 - 983.4 =)$ $-691.7 \text{ J mol}^{-1} \text{ K}^{-1} / -0.6917 \text{ kJ mol}^{-1} \text{ K}^{-1}$ <p>(1)</p> <p>Correct answer no working scores</p> <p>(2)</p> <p>If monoclinic sulfur is used (32.6) final answer = -692.5 scores (1)</p>		(2)

Question Number	Acceptable Answers	Reject	Mark
12(b)(ii)	$\Delta S^{\ominus}_{\text{surroundings}} = (-\Delta H \div T) = - \frac{-1209000 \text{ J mol}^{-1}}{298 \text{ K}}$ <p>(1)</p> $= (4057.04698)$ $= +4057 \text{ J mol}^{-1} \text{ K}^{-1} / +4.057 \text{ kJ mol}^{-1} \text{ K}^{-1}$ <p>(1)</p> <p>Correct answer without working scores 2</p>		(2)

Question Number	Acceptable Answers	Reject	Mark
12(b)(iii)	$\Delta S^{\ominus}_{\text{total}} = \Delta S^{\ominus}_{\text{system}} + \Delta S^{\ominus}_{\text{surroundings}}$ $\Delta S^{\ominus}_{\text{total}} = \text{ans (b)(i)} + \text{ans (b)(ii)}$ $= -691.7 + 4057$ $= +3365.3 \text{ J mol}^{-1} \text{ K}^{-1} / +3.3653 \text{ kJ mol}^{-1} \text{ K}^{-1}$ <p>TE on answers from (b)(i) and (b)(ii)</p>		(1)

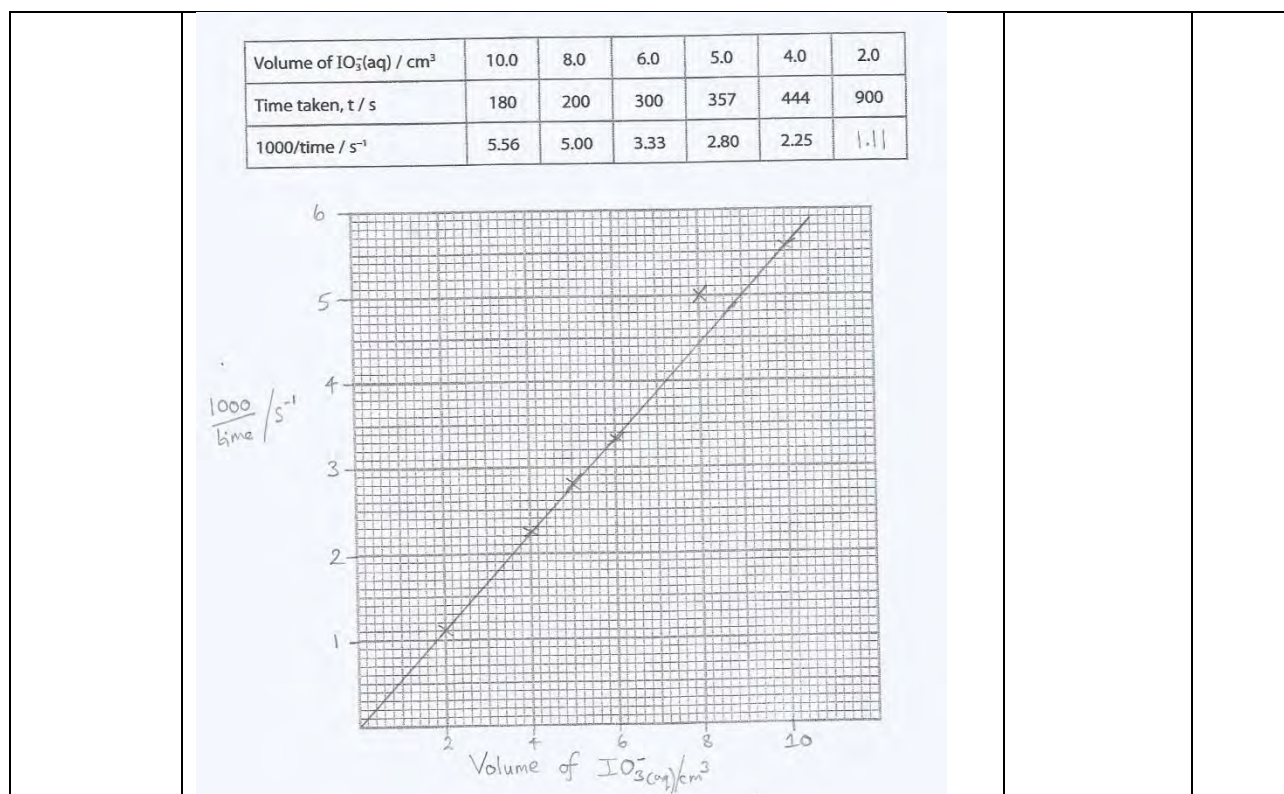
Question Number	Acceptable Answers	Reject	Mark
12(b)(iv)	<p>Marking points may be in any order Mark all 3 points independently</p> <p>M1: $\Delta S^{\ominus}_{\text{surroundings}}$ becomes less positive / smaller (magnitude) / decreases (in magnitude) (because you are dividing $-\Delta H$ by a larger T)</p> <p style="text-align: center;">(1)</p> <p>M2:</p>		(3)

Question Number	Acceptable Answers	Reject	Mark
13(a)	blue-black / blue / black (complex) colour would never form OR no colour change would be seen OR no (excess) iodine would form OR no iodine left to react with starch OR iodine would be reduced back to iodide OR iodine would react with hydrogensulfate as soon as it forms IGNORE Just hydrogensulfate would not get used up	Any other colours	(1)

Question Number	Acceptable Answers	Reject	Mark
13(b)	So the kinetics of reaction 1 can be studied OR iodine complex colour would form too soon / solution would go blue-black too soon IGNORE reference to RDS		(1)

Question Number	Acceptable Answers	Reject	Mark
13(c)(i)	Because temperature affects reaction rate ALLOW Increase in temperature increases reaction rate' or reverse argument for decreasing temperature OR To keep the rate of reaction the same OR So no change in rate constant IGNORE references to validity, reliability or 'fair test' or so temperature is not a variable		(1)

Question Number	Acceptable Answers	Reject	Mark
13(c)(ii)	<p>M1: Completed table with value 1.11 (1)</p> <p>M2: Axes correct with sensible, linear scales so at least half of the graph paper on both axes is covered (1)</p> <p>ALLOW even if graph scales do not start at (0, 0)</p> <p>M3: Axes labels fully correct with units (1)</p> <p>ALLOW 1000 t / s⁻¹ or 1000 / t / s⁻¹ on y-axis ALLOW volume / cm³ on x-axis</p> <p>M4: All points plotted correctly (± 1 small square) (1)</p> <p>Award M4 TE on the table value at 2 cm³ Do not penalise missing crosses/ circles if line is correct</p> <p>M5: Straight line drawn through (0, 0) and through all points except anomalous result at 8 cm³ (1)</p> <p><u>Exemplar:</u></p>	<p>Axes reversed OR decreasing scale</p>	(5)



Question Number	Acceptable Answers	Reject	Mark
13(c)(iii)	burette / (graduated) pipette	measuring cylinder teat pipette volumetric flask syringe	(1)

Question Number	Acceptable Answers	Reject	Mark
13(c)(iv)	volume (of iodate(V) ions) and concentration are (directly) proportional IGNORE Concentration varies as volume varies volume is proportional to rate volume is proportional to number of moles		(1)

Question Number	Acceptable Answers	Reject	Mark
13(c)(v)	<p>M1 First order (1) Note: this mark is independent of the graph drawn</p> <p>M2 because the graph is a straight line (through the origin) OR rate is proportional to $[\text{IO}_3^-]$ / rate is proportional to volume of IO_3^- OR as concentration/volume increases by (factor of) 2, rate increases by 2 (or any other numbers, including 'x') OR rate increases linearly (with concentration) (1)</p> <p>ALLOW Gradient of line is constant</p> <p>M2 dependent on M1</p>	<p>Just 'graph is a best fit line'</p> <p>References to constant half-life</p>	(2)

Question Number	Acceptable Answers	Reject	Mark
13(c)(vi)	<p>(repeat the experiment with) double the concentration of HSO_3^- and the rate doubles (keeping the iodate(V) concentration constant) OR Any other ratio i.e. any change to the concentration having the same effect on the rate</p> <p>ALLOW</p> <p>Vary the concentration and the effect on the rate is the same OR Methods involving plotting concentration/time graph and measuring constant half-life</p>	refs to the <u>gradient</u> doubling	(1)

Question Number	Acceptable Answers	Reject	Mark
13(c)(vii)	<p>M1 rate = $k [\text{IO}_3^-][\text{HSO}_3^-]$</p> <p>ALLOW $r = k [\text{IO}_3^-][\text{HSO}_3^-]$ (1)</p> <p>TE on order wrt IO_3^- given in part (v)</p> <p>M2 $\text{dm}^3 \text{mol}^{-1} \text{s}^{-1}$</p> <p>ALLOW the units in any order (1)</p> <p>TE on candidate's stated rate equation in M1</p> <p>e.g. if rate = $k [\text{HSO}_3^-]$, then award M2 as TE for units of s^{-1}</p>	Round brackets	(2)

Question Number	Acceptable Answers	Reject	Mark
13(d)(i)	<p>(measure the) time taken (for the blue-black colour to appear) and temperature</p> <p>ALLOW</p> <p>measure the rate and temperature</p> <p>IGNORE references to $\ln k$ and $1/T$</p>		(1)

Question Number	Acceptable Answers	Reject	Mark
13(d)(ii)	<p>M1 Temperature converted to kelvin</p> <p>ALLOW Kelvin given in (i) (1)</p> <p>COMMENT Only M1 can be transferred from (i) to (ii). Nothing can be credited from (ii) to (i)</p> <p>M2 The vertical axis should be $\ln \text{rate} / \ln 1/t$ ALLOW $\ln k$ (1)</p> <p>M3 The horizontal axis should be $1/T$ (1)</p> <p>M4 Straight line (with a negative gradient) (1)</p> <p>ALLOW M1, M2, M3, M4 shown on a sketch graph</p> <p>M5 Any mention of gradient (of the line) (1)</p> <p>M6 States that: $E_a = -\text{gradient} \times R$ (1)</p> <p>NB Negative sign must be shown or mentioned specifically</p> <p>NOTE: Plot "$\ln \text{rate}$ against/vs $1/T$" scores M2 and M3</p> <p>Plot "$1/T$ against/vs $\ln \text{rate}$" does not score either M2 or M3</p> <p>If axes clearly the wrong way round max (4) ie only marks M1, M4, M5 and M6 are possible</p>	<p>1/T</p> <p>1/t 1/time</p>	(6)

(Total for question 13 =22 Marks)

TOTAL FOR SECTION B = 49 MARKS

Section C

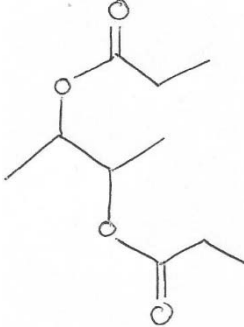
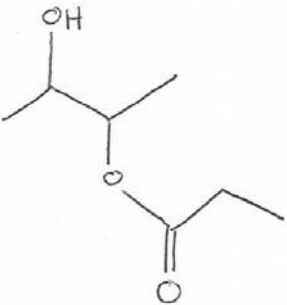
Question Number	Acceptable Answers	Reject	Mark
14(a)	<p>M1 LiAlH₄/lithium aluminium hydride/ lithium tetrahydridoaluminate((III))/ NaBH₄/sodium borohydride/ sodium tetrahydridoborate((III)) (1)</p> <p>M2 4 ([H]) (1)</p> <p>M3 CH₃CHOHCHOHCH₃ OR correct displayed (or skeletal) formula (1)</p>	If another product e.g. water is given in the equation	(3)

Question Number	Acceptable Answers	Reject	Mark
14(b)	<p>(turns from yellow-green to) colourless / yellow-green colour disappears/fades</p> <p>IGNORE bubbles</p>	<p>just "colour change"</p> <p>colour change with incorrect starting colour</p>	(1)

Question Number	Acceptable Answers	Reject	Mark
14(c)(i)	<p>butane-2,3-diol and because it has hydrogen bonds (between the molecules)</p> <p>If other intermolecular forces listed then it must be clear that only butane-2,3-diol has hydrogen bonds</p> <p>Ignore References to intramolecular hydrogen bonding</p>	hydrogen bonding to water	(1)

Question Number	Acceptable Answers	Reject	Mark
14(c)(ii)	Both molecules can form hydrogen bonds with water		(1)

Question Number	Acceptable Answers	Reject	Mark
14(d)	Mark M2 and M1 independently M1 optical (isomerism) (1) IGNORE stereoisomers M2 (molecule contains) two chiral carbon atoms OR a chiral carbon / a carbon with four different groups attached / chiral centre OR molecule exists as non-superimposable mirror images OR exists as a pair of enantiomers ALLOW chiral molecule (1)	Geometric isomerism scores (0) ...four different atoms / molecules attached.	(2)

Question Number	Acceptable Answers	Reject	Mark
14(e)(ii)	<p>IGNORE bond angles and bond lengths in all diagrams Structural / displayed formulae unless no skeletal formula</p> <p>Correct diagram with two ester groups = (2)</p>  <p>M1 for both ester groups shown</p> <p>M2 for the rest of the molecule correct</p> <p>ALLOW 1 mark for a fully-correct structure with only 1 ester bond show i.e.</p>  <p>ALLOW 1 mark for a fully-correct structure using displayed / structural formula only</p>		(2)

Question Number	Acceptable Answers				Reject	Mark
14(f)	Molecule	Peak / (cm⁻¹)	Bond		Individual values Additional wavenumbers or ranges	(2)
	butanedione	1700-1680	C=O	(1)		
	butane-2,3-diol	3750-3200	O-H	(1)		
	ALLOW (if neither mark awarded) 1 mark for wavenumbers identified with correct molecules					

Question Number	Acceptable Answers	Reject	Mark
* 14(g)	<p>IGNORE TMS Peak at Chemical shift $\delta = 0$ ppm</p> <p>M1 Three (different) proton / hydrogen environments OR Three sets of peaks shown on the spectrum (1)</p> <p>M2 One singlet and one triplet and one quartet only OR shown on diagram (1)</p> <p>M3 "n+1" rule correctly applied to at least one peak e.g. quartet formed because 3 adjacent protons/hydrogens. (1)</p> <p>M4 (Area ratios of peaks) is 3:2:1 and related to CH₃:CH₂:COOH OR shown on molecular structure Note that the word 'ratio' or the mathematical symbol as above is required (1)</p> <p>M5 (Chemical shift values, δ ppm) COOH = 10.0 - 12.0; CH₂ = 1.8 - 3.0; CH₃ = 0.1 - 1.9 OR shown on diagram as any peaks centred at these chemical shifts</p> <p>ALLOW</p>		(5)

	individual chemical shift values within the ranges (1)		
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Question Number	Acceptable Answers	Reject	Mark
14 (h)	Radio waves IGNORE electromagnetic radiation	In combination with any other radiation	1

Total for question 14 = 21 Marks)

TOTAL FOR SECTION C = 21 MARKS

TOTAL FOR PAPER = 90 MARKS

