

**UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS**

**GCE Advanced Subsidiary Level and GCE Advanced Level**

**MARK SCHEME for the October/November 2010 question paper  
for the guidance of teachers**

**9701 CHEMISTRY**

**9701/36**

Paper 3 (Advanced Practical Skills), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the October/November 2010 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	36

Question	Sections	Indicative material	Mark
1 (a)	PDO layout	<b>I</b> Volume given for Rough titre <b>and</b> accurate titre details tabulated.	1
	MMO Collection	<b>II</b> Follows instructions - dilutes 45.50 – 46.50 cm <sup>3</sup> <b>FB 1</b> <b>and</b> initial and final burette readings recorded for Rough titre <b>and</b> initial and final burette readings <b>and</b> volume of <b>FB 2</b> added recorded for each accurate titre <i>Headings should match readings.</i> <i>Do <b>not</b> award this mark if:</i> <i>50(.00) is used as an initial burette reading;</i> <i>more than one final burette reading is 50.(00);</i> <i>any burette reading is greater than 50.(00)</i>	1
	MMO Decisions	<b>III</b> Has two uncorrected, accurate titres within 0.1 cm <sup>3</sup> <i>Do not consider the Rough even if ticked.</i> <i>Do <b>not</b> award this mark if having performed two titres within 0.1 cm<sup>3</sup> a further titration is performed which is more than 0.10 cm<sup>3</sup> from the closer of the initial <b>two</b> titres, unless a fourth titration, within 0.1 cm<sup>3</sup> of the third titration (or first two) has also been carried out.</i>	1
	PDO Recording	<b>IV</b> All accurate burette readings (initial and final) recorded to nearest 0.05 cm <sup>3</sup> (Accurate titration & dilution tables) <i>Assess this mark on burette readings only</i> For candidates and Supervisor scale titre for 46.00 cm <sup>3</sup> <b>FB 1</b> diluted. Calculate titre $\times \frac{46.00}{\text{volume of FB1 added}}$ Calculate difference in Supervisor and candidate scaled values and award "quality" marks as below. [If candidate has not recorded a volume diluted, use 46.00 cm <sup>3</sup> ]	1

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	36

	MMO Quality	<p><b>V, VI and VII</b> Round any burette readings to the nearest 0.05 cm<sup>3</sup>. Check and correct subtractions in the titre table. <b>Select the “best” titre using the hierarchy:</b> two identical; titres within 0.05 cm<sup>3</sup>; titres within 0.1 cm<sup>3</sup>; etc.</p> <p>Award <b>V, VI and VII</b> for a difference from Supervisor within 0.20 cm<sup>3</sup></p> <p>Award <b>V and VI only</b> for a difference of 0.20+ cm<sup>3</sup> – 0.30 cm<sup>3</sup></p> <p>Award <b>V only</b> for a difference of 0.30+ – 0.50 cm<sup>3</sup> <i>If the “best” titres are ≥ 0.50 cm<sup>3</sup> apart cancel one of the Q marks.</i></p>	3	[7]
(b)	ACE Interpretation	<p>Calculates the mean, correct to 2 decimal places (third decimal place rounded to the nearest 0.05 cm<sup>3</sup>) from any accurate titres within 0.20 cm<sup>3</sup>. <i>A mean of exactly .x25 or .x75 is allowed but the candidate may round up or down to the nearest 0.05 cm<sup>3</sup>.</i> <i>If ALL burette readings are given to 1 decimal place then the mean can be given to 1 decimal place if numerically correct without rounding.</i> <i>Mean of 24.3 and 24.4 = 24.35 (✓)</i> <i>Mean of 24.3 and 24.4 = 24.4 (✗)</i> <b>Titres to be used in calculating the mean must be clearly shown – in an expression or ticked in the titration table.</b></p>	1	[1]
(c)	ACE Interpretation	<p><b>I</b> Expression correct in step (i) <math display="block">\frac{\text{volume diluted}}{1000} \times 0.125</math></p> <p><b>II</b> Uses answer to (i) <math>\times \frac{25}{250}</math> in step (ii)</p> <p><b>III</b> Uses answer to (ii) <math>\times 2</math> in step (iii) <b>and</b> answer to (iii) <math>\times \frac{1000}{\text{titre}}</math> in step (iv) <i>If an answer, with no working, is given in any section allow if correct.</i></p>	1 1 1	
	PDO Display	<p><b>IV</b> Appropriate working shown in a minimum of 3 sections.</p> <p><b>V</b> 3 to 5 significant figures in final answers to all sections attempted – minimum of 3 final answers required to qualify for the award of this mark.</p>	1 1	[5]

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	36

<b>(d)</b>	ACE Interpretation	<b>(i)</b> For Student A explains that final burette reading was also 0.05 cm <sup>3</sup> greater than the true value ("error" in same direction) <i>Ignore parallax error</i> <i>Not errors cancel – reason needed</i>	1	[2]
		<b>(ii)</b> For Student B explains that final burette reading was 0.05 cm <sup>3</sup> greater than the true value ("error" in opposite direction) <i>Not errors compound each other/add up</i>	1	
<b>(e)</b>	ACE Conclusions	<b>(i)</b> Explains that carbon dioxide is acidic (and its absorption reverses the colour change of the indicator)	1	[2]
	ACE Improvements	<b>(ii)</b> Puts acid/ <b>FB 3</b> in burette and pipettes NaOH/ <b>FB 2</b> into flask <b>or</b> Heat the solution/Use hot solution	1	
			<b>[Total: 17]</b>	

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	36

2 (a)	PDO Recording	I	Records results in a single table for both experiments. No repetition of headings	1	
	MMO Quality	II	Titre for either Flask A or B within 0.50 cm <sup>3</sup> of Supervisor	1	
		III	Titre for either Flask A or B within 0.30 cm <sup>3</sup> of Supervisor	1	
		IV	Titre for both Flask A and B within 0.30 cm <sup>3</sup> of Supervisor	1	[4]
(b)	ACE Interpretation	(i)	Calculates a volume of 200 cm <sup>3</sup> in step (i)	1	
		(ii)	Correctly calculates titre x 5 for each flask	1	[2]
(c)	ACE Conclusions		Mark consequentially to practical results Chooses expt with lower titre – less remains (or converse argument) <b>or</b> higher value in (b)(iii) <i>Allow ecf</i>	1	[1]
(d)	ACE Conclusions		Comparison of candidate's $K_c$ values Judgement on constancy or otherwise Supports/does not support equilibrium	1	[1]
					<b>[Total: 8]</b>

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	36

FB 7 is $\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2(\text{aq})$ ; FB 8 is $\text{Na}_2\text{SO}_4(\text{aq})$ ; FB 9 is $\text{CaCl}_2(\text{aq})$					
3 (a)	MMO Decisions	I	Selects sodium hydroxide as reagent ( <u>Not</u> if + Al) and describes (warming the solution and) <u>testing any gas evolved</u> with red litmus/pH paper	1	[5]
	MMO Collection	II	Records positive test for ammonia gas with <b>FB 7</b> only <i>Must link gas/<math>\text{NH}_3</math> with positive test (Allow even if Al mentioned in I)</i>	1	
	MMO Decisions	III	Selects barium chloride or nitrate together with HCl or $\text{HNO}_3$ <i>Do not accept <math>\text{Ba}^{2+}</math> as a reagent Accept <math>\text{Ba}^{2+}(\text{aq})</math> or a solution containing <math>\text{Ba}^{2+}</math> ions</i>	1	
	MMO Collection	IV	White ppt, persisting in acid with <b>FB 7</b> and with <b>FB 8</b> <i>Allow from unspecified strong acid provided there is no ppt with <b>FB 9</b>.</i>	1	
	MMO Conclusions	V	Mark consequentially to observations for solutions containing $\text{NH}_4^+$ and $\text{SO}_4^{2-}$ <b>ecf allowed here but not with other identities</b> <i>Allow from strong acid or from <math>\text{H}_2\text{SO}_4</math> if clearly added after <math>\text{Ba}^{2+}(\text{aq})</math></i>	1	
(b)	PDO Layout	I	(Tabulates) observations clearly, showing: observation when each reagent is first added and observation when reagent added to <u>excess</u> if there is a ppt	1	[4]
	MMO Collection	II, III and IV	1 mark for correct observations in <b>each</b> of the columns or rows representing <b>FB 7</b> , <b>FB 8</b> and <b>FB 9</b> <b>or</b> 1 mark for correct observations in the row or column representing a reagent added (initial and excess count as one row/column)	3	

Minimum observations

Solution	FB 7	FB 8	FB 9
NaOH	Green ppt insoluble (in excess)	no reaction/no change/no ppt <i>Not “–” words needed (Only penalise once)</i>	White ppt insoluble (in excess)
$\text{NH}_3$	Green ppt insoluble (in excess)	colourless <u>soln</u> /no reaction/no change/no ppt	No reaction/no change/no ppt

Page 7	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2010	9701	36

(c)	ACE Conclusions	<p><b>One mark for FB 7 and FB 9</b> containing <math>\text{Fe}^{2+}</math> and <math>\text{Ca}^{2+}</math> respectively provided no CON obs in (a) or (b)  <i>No ecf</i>  <i>Ignore FB 8, ignore supporting evidence</i></p>	1	[1]
<b>FB 10</b> is $\text{CuCO}_3(\text{s})$				
(d) (i)	MMO Collection	<p><b>I</b> observes the <u>solid turning black</u> in step (i)</p> <p><b>II</b> observes fluidity in solid layer in step (i)  <i>Allow description of fluidised solid as "liquid"</i></p>	1	[5]
	MMO Decisions	<p><b>III</b> describes an appropriate test for any of the following <u>gases</u>: <math>\text{O}_2</math>, <math>\text{CO}_2</math>, <math>\text{NH}_3</math> or <math>\text{SO}_2</math> (gas or <math>\text{O}_2</math>/etc needed)</p>	1	
	MMO Collection	<p><b>IV</b> lime water turns milky/cloudy/chalky  <i>Gas or <math>\text{CO}_2</math> turns limewater milky. scores III and IV</i></p>	1	
(ii)		<p><b>V</b> on adding acid to residue from <b>FB 10</b>, observes green solution (on warming)  <i>Ignore any residual solid</i>  <i>Allow blue-green or bluish green</i>  <i>Allow if (qualified) green solution turns blue on cooling</i>            May award <b>either III or IV</b> here but only for gas tests for <math>\text{CO}_2</math> or <math>\text{SO}_2</math> or limewater observations</p>	1	
<b>[Total: 15]</b>				