
CHEMISTRY**9701/53**

Paper 5 Planning, Analysis and Evaluation

May/June 2018

MARK SCHEME

Maximum Mark: 30

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

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PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

PUBLISHED**GENERIC MARKING PRINCIPLE 5:**

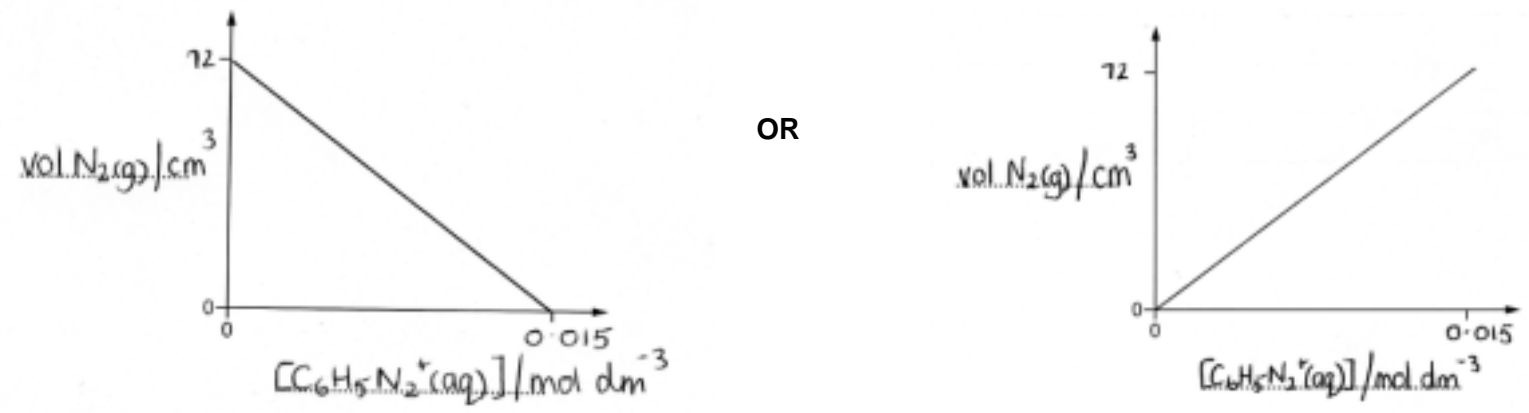
Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

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Question	Answer	Marks
1(a)	Complete circuit with ammeter in series and DC power supply	1
	Anode, cathode and solution labelled	1
1(b)	wear gloves	1
	do not dispose into the water waste / sink OR do not put down drain / sewage OR put in waste bottles	1
1(c)	Mass (of electrode) before and after experiment AND mass unit	1
1(d)	charge = $0.5 \times 30 \times 60 = 900$ C	1
1(e)	$0.282 / 63.5 = 4.44 \times 10^{-3}$ (mol) OR 0.00444	1
1(f)	$(900 / 4.44 \times 10^{-3}) = 202702.7027$ C	1
1(g)	2 moles of electrons are produced / removed / released (so 2 Faradays OR $2 \times 96\,500$)	1
1(h)	(Faraday) value is smaller AND (apparent) mass / moles / amount is more (for same charge passed)	1
1(i)	CuO is formed / oxidation of copper / carbon / soot is formed	1
1(j)	Some copper falls off the electrode during electrolysis / falls to the bottom of the beaker OR Some copper is lost during washing	1

Question	Answer	Marks
2(a)	Water bath/beaker of water containing thermometer around flask	1
	Controlled heat source or heater/temperature regulator	1
2(b)(i)	Moles $N_2 = 72 / 24\,000 = 0.003$ moles (1 mol $C_6H_5N_2^+Cl^- \rightarrow 1$ mol N_2)	1
	Moles $C_6H_5N_2^+$ in 1000 cm^3 solution = $0.003 \times (1000 / 200) = 1.50 \times 10^{-2}$ (mol)	1
2(b)(ii)	 <p style="text-align: center;">OR</p>	
	Axes (label with quantity or correct unit) and values correct	1
	Straight line from axis marks OR from 0,0 over most of the axes	1

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Question	Answer				Marks																																												
2(c)	<table border="1"> <thead> <tr> <th data-bbox="595 220 824 268">A</th> <th data-bbox="824 220 1048 268">B</th> <th data-bbox="1048 220 1272 268">C</th> <th data-bbox="1272 220 1677 268">D</th> </tr> <tr> <th data-bbox="595 268 824 387">Time / min</th> <th data-bbox="824 268 1048 387">volume of nitrogen, V / cm³</th> <th data-bbox="1048 268 1272 387">V / V_{FINAL}</th> <th data-bbox="1272 268 1677 387">[C₆H₅N₂⁺Cl⁻(aq)] / mol dm⁻³</th> </tr> </thead> <tbody> <tr> <td data-bbox="595 387 824 435">0.0</td> <td data-bbox="824 387 1048 435">0</td> <td data-bbox="1048 387 1272 435">0.000</td> <td data-bbox="1272 387 1677 435">0.0150</td> </tr> <tr> <td data-bbox="595 435 824 483">2.0</td> <td data-bbox="824 435 1048 483">9</td> <td data-bbox="1048 435 1272 483">0.125</td> <td data-bbox="1272 435 1677 483">0.0131</td> </tr> <tr> <td data-bbox="595 483 824 531">4.0</td> <td data-bbox="824 483 1048 531">17</td> <td data-bbox="1048 483 1272 531">0.236</td> <td data-bbox="1272 483 1677 531">0.0115</td> </tr> <tr> <td data-bbox="595 531 824 579">6.0</td> <td data-bbox="824 531 1048 579">24</td> <td data-bbox="1048 531 1272 579">0.333</td> <td data-bbox="1272 531 1677 579">0.0100</td> </tr> <tr> <td data-bbox="595 579 824 627">8.0</td> <td data-bbox="824 579 1048 627">30</td> <td data-bbox="1048 579 1272 627">0.417</td> <td data-bbox="1272 579 1677 627">0.00875</td> </tr> <tr> <td data-bbox="595 627 824 675">10.0</td> <td data-bbox="824 627 1048 675">35</td> <td data-bbox="1048 627 1272 675">0.486</td> <td data-bbox="1272 627 1677 675">0.00771</td> </tr> <tr> <td data-bbox="595 675 824 722">12.0</td> <td data-bbox="824 675 1048 722">40</td> <td data-bbox="1048 675 1272 722">0.556</td> <td data-bbox="1272 675 1677 722">0.00666</td> </tr> <tr> <td data-bbox="595 722 824 770">14.0</td> <td data-bbox="824 722 1048 770">44</td> <td data-bbox="1048 722 1272 770">0.611</td> <td data-bbox="1272 722 1677 770">0.00584</td> </tr> <tr> <td data-bbox="595 770 824 842">16.0</td> <td data-bbox="824 770 1048 842">48</td> <td data-bbox="1048 770 1272 842">0.667</td> <td data-bbox="1272 770 1677 842">0.00500</td> </tr> </tbody> </table>				A	B	C	D	Time / min	volume of nitrogen, V / cm ³	V / V _{FINAL}	[C ₆ H ₅ N ₂ ⁺ Cl ⁻ (aq)] / mol dm ⁻³	0.0	0	0.000	0.0150	2.0	9	0.125	0.0131	4.0	17	0.236	0.0115	6.0	24	0.333	0.0100	8.0	30	0.417	0.00875	10.0	35	0.486	0.00771	12.0	40	0.556	0.00666	14.0	44	0.611	0.00584	16.0	48	0.667	0.00500	
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Column values for D correctly calculated				1																																													
3 sf in C and D				1																																													
2(d)	Candidate's calculated points correctly plotted from table in 2(c)				1																																												
	Smooth curve of best fit				1																																												
2(e)	Tangent drawn at time zero				1																																												
	2 sets of co-ordinates shown				1																																												
	calculation of gradient of tangent				1																																												
	mol dm ⁻³ minute(s) ⁻¹				1																																												

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Question	Answer					Marks															
2(f)	<table border="1"> <thead> <tr> <th data-bbox="640 220 875 268">concentration 1</th> <th data-bbox="875 220 1010 268">Time 1</th> <th data-bbox="1010 220 1256 268">concentration 2</th> <th data-bbox="1256 220 1458 268">time 2</th> <th data-bbox="1458 220 1637 268">$t_{1/2}$</th> </tr> </thead> <tbody> <tr> <td data-bbox="640 268 875 316">(0.0120)</td> <td data-bbox="875 268 1010 316">3</td> <td data-bbox="1010 268 1256 316">(0.0060)</td> <td data-bbox="1256 268 1458 316">13.4</td> <td data-bbox="1458 268 1637 316">10.4</td> </tr> <tr> <td data-bbox="640 316 875 368">0.010</td> <td data-bbox="875 316 1010 368">6</td> <td data-bbox="1010 316 1256 368">0.005</td> <td data-bbox="1256 316 1458 368">16.0</td> <td data-bbox="1458 316 1637 368">10.0</td> </tr> </tbody> </table>					concentration 1	Time 1	concentration 2	time 2	$t_{1/2}$	(0.0120)	3	(0.0060)	13.4	10.4	0.010	6	0.005	16.0	10.0	
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Columns 1 and 3					1																
Columns 2 and 4					1																
Half-lives correctly calculated.					1																
2(g)	First order AND because half-lives are constant/equal					1															