



Cambridge International AS & A Level

CHEMISTRY**9701/34**

Paper 3 Advanced Practical Skills 2

October/November 2022

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2022 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

This document consists of **13** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

PUBLISHED**GENERIC MARKING PRINCIPLE 5:**

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- 1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
- 2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
- 3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
- 4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
- 5 'List rule' guidance
 For questions that require ***n*** responses (e.g. State **two** reasons ...):
 - The response should be read as continuous prose, even when numbered answer spaces are provided.
 - Any response marked *ignore* in the mark scheme should not count towards ***n***.
 - Incorrect responses should not be awarded credit but will still count towards ***n***.
 - Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
 - Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Examples of how to apply the list rule			
State three reasons ... [3]			
A	1	Correct	✓
	2	Correct	✓
	3	Wrong	✗
			2
B (4 responses)	1	Correct, Correct	✓, ✓
	2	Correct	✓
	3	Wrong	ignore
			3
C (4 responses)	1	Correct	✓
	2	Correct, Wrong	✓, ✗
	3	Correct	ignore
			2
D (4 responses)	1	Correct	✓
	2	Correct, CON (of 2.)	✗, (discount 2)
	3	Correct	✓
			2
E (4 responses)	1	Correct	✓
	2	Correct	✓
	3	Correct, Wrong	✓
			3
F (4 responses)	1	Correct	✓
	2	Correct	✓
	3	Correct CON (of 3.)	✗ (discount 3)
			2
G (5 responses)	1	Correct	✓
	2	Correct	✓
	3	Correct Correct CON (of 4.)	✓ ignore ignore
			3
H (4 responses)	1	Correct	✓
	2	Correct	✗
	3	CON (of 2.) Correct	(discount 2) ✓
			2
I (4 responses)	1	Correct	✓
	2	Correct	✗
	3	Correct CON (of 2.)	✓ (discount 2)
			2

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Question	Answer	Marks
1(a)	<p>I the following data must be shown: with unambiguous headings and correct units (mass of) container + FB 2 (mass of) container (+ residue) (mass of) FB 2 correctly calculated units: / g, (g) or g by every entry</p> <p>II all thermometer readings recorded to .0 or .5 °C at least one reading should be .0 and at least one reading should be .5</p> <p>III and IV award of accuracy marks check supervisor's $\Delta T (= T_{\max} - T_{t=2}$ from table), choose the applicable range then mark accordingly: ΔT less than 10.0 °C: δ within 0.5 = 2 marks: δ within 1.0 = 1 mark ΔT in range 10.0–19.5 °C: δ within 1.0 = 2 marks: δ within 2.0 = 1 mark ΔT in range 20.0–29.5 °C: δ within 2.0 = 2 marks: δ within 3.5 = 1 mark ΔT in range 30.0–39.5 °C: δ within 3.0 = 2 marks: δ within 5.0 = 1 mark where δ is the difference between the supervisor's and candidate's ΔT</p>	4
1(b)(i)	<p>M1 axes correctly labelled temperature on y-axis and time on the x-axis and correct units</p> <p>M2 appropriate scales linear scales chosen so that the graph occupies, including the extra 5 °C on the y-axis, more than half the available length for both axes</p> <p>M3 all points recorded in table accurately plotted if the point should be on a line, it must be on the line if the point should not be on a line, it must not be on a line the point should be in the correct small square and must be correct to within half a small square</p>	3

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Question	Answer	Marks
1(b)(ii)	<p>M1 2 lines of best fit drawn (straight or smoothly curved lines), one before adding FB 2 and one for the cooling of the mixture and extrapolated to 2½ minutes so intersect is at or above the maximum thermometer reading recorded in (a) and vertical line drawn at 2½ minutes</p> <p>M2 correct ΔT from the intersect at 2½ minutes to 1 dp and to within 0.5 °C of examiner value</p>	2
1(c)(i)	<p>correctly calculates amount of $\text{CuSO}_4 = 0.80 \times 0.03 = 0.024(00)$ (mol) and answer to 2–4 significant figures (sf)</p>	1
1(c)(ii)	<p>correctly calculates amount of Zn = $\frac{\text{mass in (a)}}{65.4}$ (mol) and moles Zn > moles Cu^{2+} (owtte)</p>	1
1(c)(iii)	<p>correctly calculates heat energy = $30 \times 4.18 \times \Delta T(\text{b)(ii)}$ (J) and answer to 2–4 sf</p>	1
1(c)(iv)	<p>correct expression $\Delta H = \frac{\text{(c)(iii)}}{\text{(c)(i)}} \times 1000$ (kJ mol^{-1}) and negative sign shown and answer given to 2–4 sf</p>	1

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Question	Answer	Marks
1(d)	<p>conclusion M1 ticks 3rd box and an explanation attempted</p> <p>explanation M2 more (moles of) Cu^{2+} present in (old) FB 1 (so $\Delta T \uparrow$) / solution (for old sample) is more concentrated with respect to Cu^{2+} (ora)</p>	2

Question	Answer	Marks
2(a)	<p>I correct headings and units in Results section</p> <ul style="list-style-type: none"> • mass of crucible (+ lid) • mass of crucible (+ lid) + FB 3 • mass of crucible (+ lid) + residue / contents / CuO • mass of FB 3 • mass of residue <p>units: / g, (g) or g by every entry</p> <p>II readings and subtraction</p> <ul style="list-style-type: none"> • all balance readings are to 2 or 3 dp • mass of FB 3 and residue are correctly calculated • mass of FB 3 added is between 1.50 and 1.80 g (candidate's value) <p>III and IV award of accuracy marks award III mark if $\delta \leq \pm 0.25$ award IV marks if $\delta \leq \pm 0.10$ where δ is the difference between the supervisor's and candidate's mass ratio = mass in FB 3 / mass residue calculated to 2 dp</p>	4

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Question	Answer	Marks
2(b)(i)	<p>correctly calculates amount of CuO = mass of residue from (a) / 79.5 (mol) and answer given to 2–4 sf</p>	1
2(b)(ii)	<p>correctly calculates mass of CO₂ = mass of FB 3 – mass of residue and amount of CO₂ = answer / 44 (mol) and answer given to 2–4 sf</p>	1
2(b)(iii)	<p>either</p> <ul style="list-style-type: none"> • yes, amount / moles of CuO = amount / moles of CO₂ (within experimental error) (so no other decomposition product) <p>or</p> <ul style="list-style-type: none"> • no, amount / moles of CuO does not equal amount / moles of CO₂ (so another decomposition product formed) 	1
2(c)	<p>award for any possible formula of basic copper(II) carbonate which includes CuCO₃ and either one or both of Cu(OH)₂ and CuO e.g. CuCO₃•Cu(OH)₂, CuCO₃•CuO, CuCO₃•CuO•Cu(OH)₂, CuCO₃•CuO•2Cu(OH)₂</p>	1
2(d)(i)	<p>improvement: heat to constant mass (owtte)</p>	1

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Question	Answer	Marks
2(d)(ii)	<p>M1 reagent: selects a named (mineral) acid or gives its correct formula</p> <p>M2 observation and conclusion either no fizzing / no reaction / no change = no carbonate (remains) / complete decomposition or fizzing = some carbonate remains / not fully decomposed or gas / CO₂ gives white ppt with limewater = some carbonate remains / not fully decomposed</p>	2

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Question	Answer	Marks
FB 4 is $\text{CuSO}_4(\text{aq})$; FB 5 is $\text{C}_6\text{H}_{12}\text{O}_6(\text{aq})$		
3(a)(i)	<p>observations with FB 4</p> <p>Test 1</p> <p>+ $\text{NH}_3(\text{aq})$:</p> <ul style="list-style-type: none"> • (pale) blue ppt • dark / deep blue (solution) forms (with excess) <p>+ $\text{EDTA}(\text{aq})$:</p> <ul style="list-style-type: none"> • paler / lighter (blue) <p>Test 2</p> <p>+ $\text{EDTA}(\text{aq})$:</p> <ul style="list-style-type: none"> • (slightly) deeper / darker (blue) <p>+ $\text{NaOH}(\text{aq})$:</p> <ul style="list-style-type: none"> • no change / no ppt / remains blue (solution) <p>+ FB 5 and warm:</p> <ul style="list-style-type: none"> • green / yellow / brown (colour) • ppt / solid <p>Test 3</p> <p>+ Mg:</p> <ul style="list-style-type: none"> • bubbles (on ribbon) • solution colour fades / paler blue / colourless • ribbon darkens or pink / red-brown / brown / black and solid / ppt / coating / layer on Mg 	5

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Question	Answer	Marks
3(a)(ii)	<p>M1: 'Table' with headings: 'test / experiment / reagents' and 'observations' (owtte) and at least 1 appropriate reagent and at least 1 observation shown</p> <p>M2 and M3: marks are awarded using 2 • = 1 mark</p> <p>reagent:</p> <ul style="list-style-type: none"> • selects a named (aqueous) barium salt / correct formula <p>observation:</p> <ul style="list-style-type: none"> • White ppt <p>either</p> <p>reagent:</p> <ul style="list-style-type: none"> • hydrochloric / nitric acid / HCl / HNO₃ (add to ppt) <p>observation:</p> <ul style="list-style-type: none"> • ppt insoluble <p>or</p> <p>reagent:</p> <ul style="list-style-type: none"> • (acidified aqueous) potassium manganate(VII) / KMnO₄ (add to fresh sample of solution or to mixture after addition of (aqueous) barium salt) <p>observation:</p> <ul style="list-style-type: none"> • (solution / KMnO₄) remains purple / no change 	3
3(a)(iii)	identifies CuSO ₄ with evidence of sulfate from results	1
3(a)(iv)	Mg(s) + Cu ²⁺ (aq) → Mg ²⁺ (aq) + Cu(s) or Mg(s) + 2H ⁺ (aq) → Mg ²⁺ (aq) + H ₂ (g)	1

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Question	Answer	Marks												
3(b)(i)	observations with FB 5 <table border="1" data-bbox="353 248 1921 544"> <thead> <tr> <th data-bbox="353 248 488 312"><i>mark</i></th> <th data-bbox="488 248 712 312"><i>test</i></th> <th data-bbox="712 248 1921 312"><i>observation</i></th> </tr> </thead> <tbody> <tr> <td data-bbox="353 312 488 376">M1</td> <td data-bbox="488 312 712 376">+ Tollens'</td> <td data-bbox="712 312 1921 376">(silver) mirror / black ppt on warming</td> </tr> <tr> <td data-bbox="353 376 488 480">M2</td> <td data-bbox="488 376 712 480">+ (acidified) KMnO₄(aq)</td> <td data-bbox="712 376 1921 480">pink / purple to colourless (solution) OR KMnO₄ decolourises (on warming)</td> </tr> <tr> <td data-bbox="353 480 488 544">M3</td> <td data-bbox="488 480 712 544">+ Na₂CO₃(aq)</td> <td data-bbox="712 480 1921 544">no change / no bubbling</td> </tr> </tbody> </table>	<i>mark</i>	<i>test</i>	<i>observation</i>	M1	+ Tollens'	(silver) mirror / black ppt on warming	M2	+ (acidified) KMnO ₄ (aq)	pink / purple to colourless (solution) OR KMnO ₄ decolourises (on warming)	M3	+ Na ₂ CO ₃ (aq)	no change / no bubbling	3
<i>mark</i>	<i>test</i>	<i>observation</i>												
M1	+ Tollens'	(silver) mirror / black ppt on warming												
M2	+ (acidified) KMnO ₄ (aq)	pink / purple to colourless (solution) OR KMnO ₄ decolourises (on warming)												
M3	+ Na ₂ CO ₃ (aq)	no change / no bubbling												
3(b)(ii)	identifies aldehyde / RCHO / –CHO and states positive result with Tollens' (owtte)	1												