



Cambridge International AS & A Level

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CENTRE
NUMBER

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CHEMISTRY

9701/23

Paper 2 AS Level Structured Questions

May/June 2020

1 hour 15 minutes

You must answer on the question paper.

You will need: Data booklet

INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working, use appropriate units and use an appropriate number of significant figures.

INFORMATION

- The total mark for this paper is 60.
- The number of marks for each question or part question is shown in brackets [].

This document has **12** pages. Blank pages are indicated.



Answer **all** the questions in the spaces provided.

1 (a) A sample of barium is heated in oxygen.

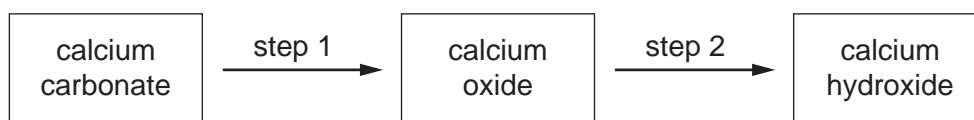
(i) Describe **two** observations for this reaction.

.....
 [2]

(ii) Write an equation for this reaction. Include state symbols.

..... [1]

(b) Calcium carbonate can be converted into calcium hydroxide in a two-step process.



(i) Describe how the two-step process is carried out to convert calcium carbonate into calcium hydroxide. Include relevant equations.

.....

 [3]

(ii) Name the type of reaction occurring when calcium carbonate is converted into calcium oxide.

..... [1]

(iii) State **one** common use for both calcium carbonate and calcium hydroxide.

..... [1]

(c) Gallium is a silver-grey solid. Aluminium and gallium share many similar chemical properties.

- (i) Construct an equation for the reaction of gallium when heated in oxygen to form gallium oxide, Ga_2O_3 .

..... [1]

- (ii) Deduce the oxidation number of gallium in Ga_2O_3 .

..... [1]

- (iii) Complete the table by predicting the formula of each gallium-containing product formed when gallium oxide reacts separately with hot aqueous hydrochloric acid and with hot concentrated sodium hydroxide.

| reagents and conditions | formula of gallium-containing product |
|--|---------------------------------------|
| gallium oxide + hot $\text{HCl}(\text{aq})$ | |
| gallium oxide + hot concentrated $\text{NaOH}(\text{aq})$ | |

[2]

[Total: 12]

- 2 (a) Explain what is meant by the term *relative isotopic mass*.

.....

 [2]

- (b) A sample of copper contains two isotopes, ^{63}Cu and ^{65}Cu . The relative atomic mass of the copper in this sample is 63.55.

Calculate the percentage abundance of each of these isotopes. Show your working.

percentage abundance of ^{63}Cu = %

percentage abundance of ^{65}Cu = %
 [2]

- (c) (i) Name the type of bonding within a sample of solid copper.

..... [1]

- (ii) Draw a labelled diagram to show the bonding within a sample of solid copper.

[2]

- (iii) State the electronic configuration of a copper atom.

$1s^2$ [1]

5

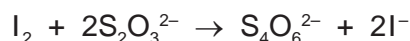
- (d) A student is provided with a sample of hydrated copper(II) sulfate, $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$, and is asked to determine the value of x .

The student dissolves a sample of the hydrated copper(II) sulfate in water and adds it to an excess of aqueous potassium iodide to make a total volume of 250.0 cm^3 of solution.



The amount of iodine produced during this reaction is found by titrating a sample of this solution with sodium thiosulfate solution.

25.0 cm^3 of the iodine-containing solution requires 20.0 cm^3 of 0.10 mol dm^{-3} sodium thiosulfate solution.



- (i) Calculate the amount, in mol, of copper(II) sulfate present in the original sample of hydrated copper(II) sulfate.

Show your working.

amount of copper(II) sulfate = mol [2]

- (ii) A total of 7.98 g of CuSO_4 is present in 10.68 g of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$.

Complete each row of the table to calculate the value of x , where x is an integer.

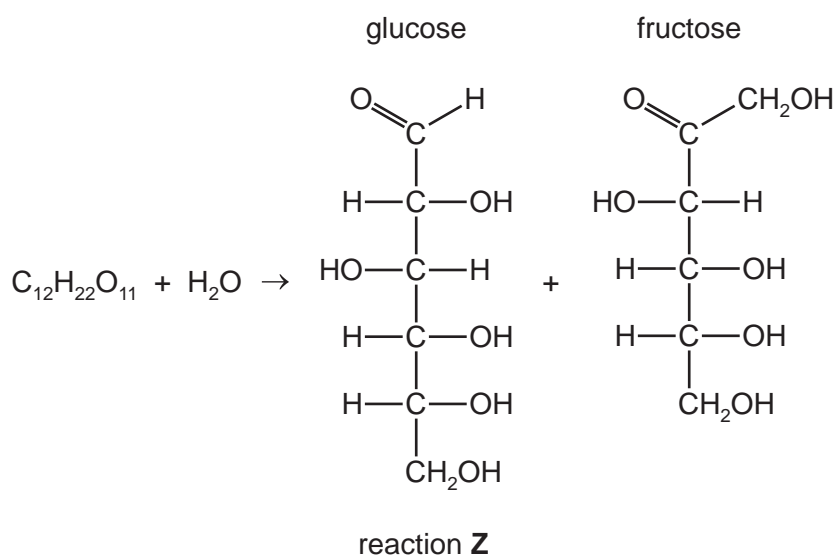
$[M_r: \text{CuSO}_4, 159.6]$

| | |
|---|-----------------------|
| amount of CuSO_4 in 10.68 g of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ | mol |
| amount of H_2O in 10.68 g of $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ | mol |
| value of x | $x = \dots\dots\dots$ |

[3]

[Total: 13]

- 3 Sucrose is a white crystalline solid, $C_{12}H_{22}O_{11}$. In reaction **Z**, sucrose reacts with water in the presence of a catalyst, aqueous hydrochloric acid, to form glucose and fructose.

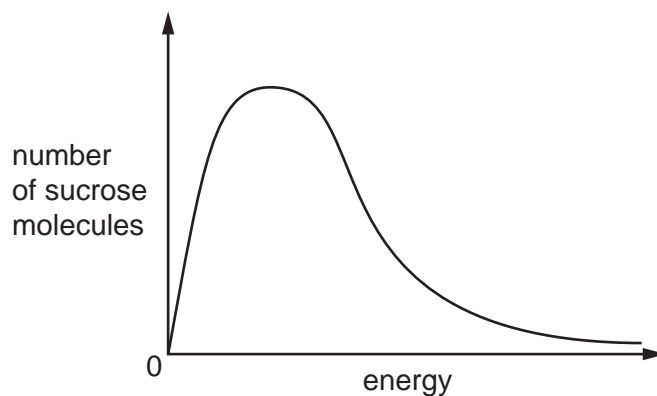


- (a) (i) Suggest a name for the reaction that occurs when sucrose reacts with water to form glucose and fructose.

..... [1]

- (ii) If no catalyst is added in reaction **Z**, the reaction is very slow.

Label the Boltzmann distribution to show the effect of adding a catalyst to the sample of sucrose and water molecules at constant temperature.



Explain your labelled diagram.

.....

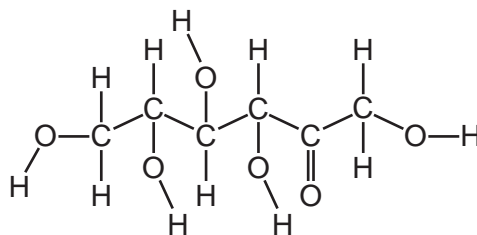
[3]

- (b) Both fructose and glucose contain chiral centres.

- (i) Explain what is meant by the term *chiral centre*.

..... [1]

- (ii) On the diagram of the fructose molecule, label all the chiral centres with an asterisk (*).



[1]

- (iii) Determine the empirical formula of fructose.

..... [1]

- (c) (i) Explain what is meant by the term *enthalpy change of combustion*.

.....

..... [2]

- (ii) Write the equation for the complete combustion of sucrose.

..... [1]

The enthalpy change of reaction **Z**, ΔH_r , can be calculated using the enthalpy change of combustion data given in the table.

| substance | enthalpy change of combustion, $\Delta H_c / \text{kJ mol}^{-1}$ |
|-----------|--|
| sucrose | -5643 |
| glucose | -2805 |
| fructose | -2810 |

- (iii) Use the data in the table to calculate the enthalpy change for the reaction occurring when sucrose reacts with water, ΔH_r . You should draw a labelled Hess' cycle to show your working.

$\Delta H_r = \dots\dots\dots \text{kJ mol}^{-1}$ [2]

[Total: 12]

- 4 Hexane, C_6H_{14} , is a colourless liquid. Two test-tubes contain equal amounts of hexane. 1 cm^3 of bromine, $Br_2(aq)$, is added to both test-tubes. One test-tube is kept in the dark and the other is exposed to sunlight.

The table describes the appearance of each test-tube after one hour.

| test-tube conditions | observations |
|----------------------|--|
| in the dark | no change, mixture remains orange |
| in sunlight | colour of mixture fades to pale yellow |

- (a) The test-tube in the dark is kept cool and is not exposed to ultraviolet light.

Explain the observations for the test-tube kept in the dark.

.....
 [2]

- (b) In sunlight, bromine reacts with hexane by a mechanism which occurs via a series of steps.

- (i) State the name of the mechanism of the reaction that occurs.

..... [1]

- (ii) Give an equation which shows a propagation step in this reaction in which hexane produces $\bullet C_6H_{13}$.

..... [1]

- (iii) Give an equation which shows a propagation step in this reaction that produces 1-bromohexane.

..... [1]

- (iv) Give an equation which shows a termination step in this reaction that produces 1-bromohexane.

..... [1]

(c) **A** and **B** are different straight chain alkenes with molecular formula, C_6H_{12} .

A does not show stereoisomerism.

A reacts with potassium manganate(VII) to form hexane-1,2-diol.

(i) Draw the structural formula of **A**.

[1]

(ii) State the conditions needed for this reaction of **A**.

..... [2]

(d) **B** reacts with hydrogen gas in the presence of a platinum catalyst to produce hexane.

(i) Name the type of reaction occurring.

..... [1]

(ii) In terms of σ and π bonds, describe any similarities and differences in the type of carbon-carbon bonds in **B** and the type of carbon-carbon bonds in hexane.

.....

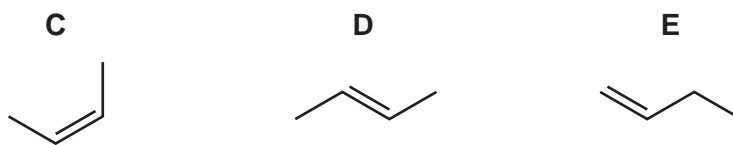
.....

..... [2]

[Total: 12]

10

- 5 **C**, **D** and **E** are isomers of each other.
They are made by passing an alcohol vapour over an aluminium oxide catalyst.



- (a) (i) Name the type of reaction occurring.

..... [1]

- (ii) Draw the displayed formula of the alcohol used in this reaction.

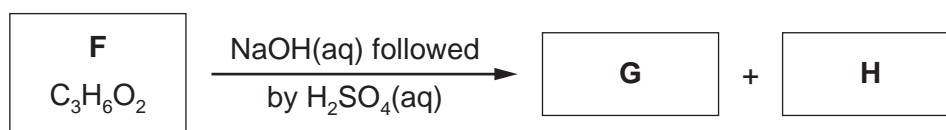
[2]

- (iii) Name the isomers **C**, **D** and **E**.

| isomer | name |
|----------|------|
| C | |
| D | |
| E | |

[2]

- (b) **F** is an organic molecule which has the molecular formula $C_3H_6O_2$.
When **F** is heated with $NaOH(aq)$ followed by $H_2SO_4(aq)$ the products **G** and **H** are made.



Separate samples of **G** and **H** are added to

- $Na_2CO_3(aq)$
- sodium metal
- alkaline aqueous iodine.

The observations are described in the table.

| reagent(s) | G | H |
|-------------------------|------------------------------------|------------------------------------|
| $Na_2CO_3(aq)$ | colourless bubbles of gas produced | no visible reaction |
| Na(s) | colourless bubbles of gas produced | colourless bubbles of gas produced |
| alkaline aqueous iodine | no visible reaction | yellow precipitate forms |

- (i) Complete the table to identify the functional groups present in **F**, **G** and **H**.

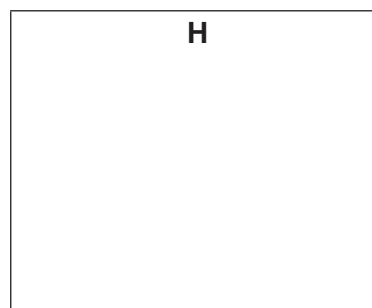
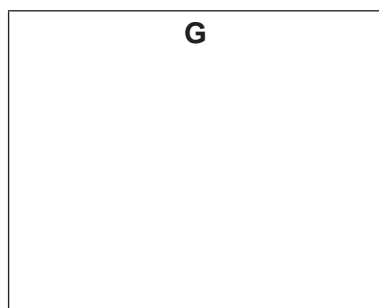
| | functional group |
|----------|------------------|
| F | |
| G | |
| H | |

[3]

- (ii) Name the yellow precipitate formed when alkaline aqueous iodine reacts with **H**.

..... [1]

- (iii) Draw the structures of **G** and **H**.



[2]

[Total: 11]

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