

Cambridge  
International  
AS & A Level

**Cambridge International Examinations**  
Cambridge International Advanced Subsidiary and Advanced Level

CANDIDATE  
NAME

CENTRE  
NUMBER

--	--	--	--	--

CANDIDATE  
NUMBER

--	--	--	--

**CHEMISTRY**

**9701/22**

Paper 2 Structured Questions AS Core

**May/June 2014**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

Additional Materials: Data Booklet

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

A Data Booklet is provided.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

This document consists of **9** printed pages and **3** blank pages.



Answer **all** the questions in the spaces provided.

- 1 (a) Explain what is meant by the term *nucleon number*.

.....  
 ..... [1]

- (b) Bromine exists naturally as a mixture of two stable isotopes,  $^{79}\text{Br}$  and  $^{81}\text{Br}$ , with relative isotopic masses of 78.92 and 80.92 respectively.

- (i) Define the term *relative isotopic mass*.

.....  
 .....  
 ..... [2]

- (ii) Using the relative atomic mass of bromine, 79.90, calculate the relative isotopic abundances of  $^{79}\text{Br}$  and  $^{81}\text{Br}$ .

[3]

- (c) Bromine reacts with the element **A** to form a compound with empirical formula  $\text{ABr}_3$ . The percentage composition by mass of  $\text{ABr}_3$  is **A**, 4.31; Br, 95.69.

Calculate the relative atomic mass,  $A_r$ , of **A**.  
 Give your answer to **three** significant figures.

$A_r$  of **A** = ..... [3]

(d) The elements in Period 3 of the Periodic Table show different behaviours in their reactions with oxygen.

(i) Describe what you would **see** when separate samples of magnesium and sulfur are reacted with oxygen.

Write an equation for each reaction.

magnesium

.....  
 .....

sulfur

.....  
 .....

[4]

(ii) Write equations for the reactions of aluminium oxide,  $Al_2O_3$ , with sodium hydroxide,

.....

hydrochloric acid.

.....

[2]

(e) Phosphorus reacts with chlorine to form  $PCl_5$ .

State the shape of and two different bond angles in a molecule of  $PCl_5$ .

shape of  $PCl_5$  .....

bond angles in  $PCl_5$  ..... [2]

[Total: 17]

- 2 A 6.30 g sample of hydrated ethanedioic acid,  $\text{H}_2\text{C}_2\text{O}_4 \cdot x\text{H}_2\text{O}$ , was dissolved in water and the solution made up to  $250 \text{ cm}^3$ .

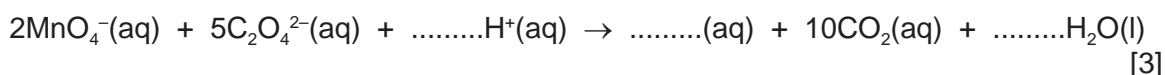
A  $25.0 \text{ cm}^3$  sample of this solution was acidified and titrated with  $0.100 \text{ mol dm}^{-3}$  potassium manganate(VII) solution.  $20.0 \text{ cm}^3$  of this potassium manganate(VII) solution was required to react fully with the ethanedioate ions,  $\text{C}_2\text{O}_4^{2-}$ , present in the sample.

(a) The  $\text{MnO}_4^-$  ions in the potassium manganate(VII) *oxidise* the ethanedioate ions.

- (i) Explain, in terms of electron transfer, the meaning of the term *oxidise* in the sentence above.

.....  
 ..... [1]

- (ii) Complete and balance the ionic equation for the reaction between the manganate(VII) ions and the ethanedioate ions.



- (b) (i) Calculate the number of moles of manganate(VII) used in the titration.

[1]

- (ii) Use the equation in (a)(ii) and your answer to (b)(i) to calculate the number of moles of  $\text{C}_2\text{O}_4^{2-}$  present in the  $25.0 \text{ cm}^3$  sample of solution used.

[1]

- (iii) Calculate the number of moles of  $\text{H}_2\text{C}_2\text{O}_4 \cdot x\text{H}_2\text{O}$  in 6.30 g of the compound.

[1]

- (iv) Calculate the relative formula mass of  $\text{H}_2\text{C}_2\text{O}_4 \cdot x\text{H}_2\text{O}$ .

[1]

- (v) The relative formula mass of anhydrous ethanedioic acid,  $\text{H}_2\text{C}_2\text{O}_4$ , is 90.

Calculate the value of  $x$  in  $\text{H}_2\text{C}_2\text{O}_4 \cdot x\text{H}_2\text{O}$ .

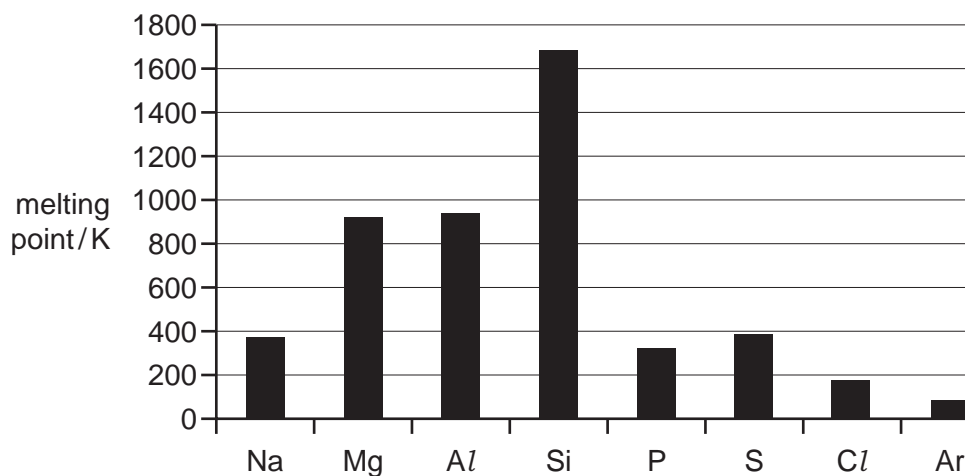
[1]

[Total: 9]

## 5

3 The elements in Period 3 of the Periodic Table show variations in their behaviour across the period.

(a) The bar chart below shows the variation of melting points of the elements across Period 3.



In each of the following parts of this question you should clearly identify the interactions involved and, where appropriate, explain their relative magnitudes.

(i) Explain the general increase in melting point from Na to Al.

.....

.....

.....

..... [3]

(ii) Explain the variation of melting points from P to Ar.

.....

.....

.....

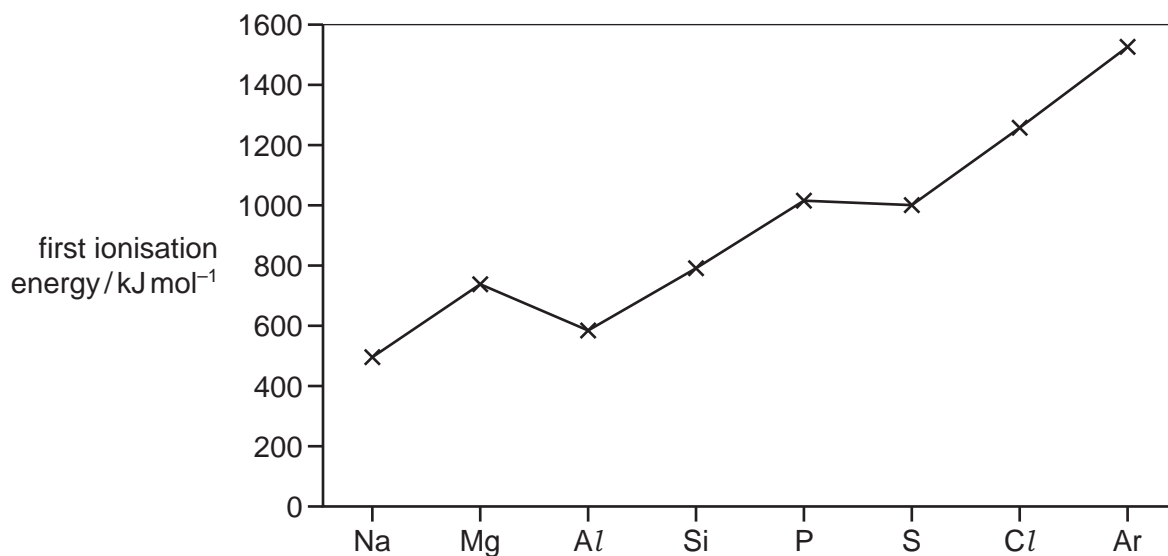
..... [3]

(iii) Explain why Si has a much higher melting point than any of the other elements in the period.

.....

..... [1]

(b) The graph below shows the variation of the first ionisation energies across Period 3.



(i) Explain why the first ionisation energy of Ar is greater than that of Cl.

.....  
 ..... [1]

(ii) Explain why the first ionisation energy of Al is less than that of Mg.

.....  
 ..... [1]

(iii) Explain why the first ionisation energy of S is less than that of P.

.....  
 ..... [1]

[Total: 10]

4 Crude oil is processed to give a wide variety of hydrocarbons.

(a) Give the names of one physical process and one chemical process carried out during the processing of crude oil.

physical process .....

chemical process .....

[2]

(b) Alkanes and alkenes can both be obtained from crude oil.

(i) Explain why alkanes are unreactive.

.....

..... [2]

(ii) State the bond angles in a molecule of

ethane, .....

ethene. ....

[1]

(iii) State the shape of each molecule in terms of the arrangement of the atoms bonded to each carbon atom.

ethane ..... ethene ..... [1]

(iv) Explain why these molecules have different shapes in terms of the carbon-carbon bonds present.

.....

..... [1]

(c) (i) Use a series of equations to describe the mechanism of the reaction of ethane with chlorine to form chloroethane. Name the steps in this reaction.

.....

.....

.....

.....

..... [5]

(ii) Write an equation to show how butane could be produced as a by-product of this reaction.

..... [1]

[Total: 13]

5 A hydrocarbon, **P**, with the formula  $C_6H_{12}$  readily decolourises bromine.

On reaction with hot, concentrated, acidified potassium manganate(VII) solution a single organic product, **Q**, is obtained.

**Q** gives an orange precipitate when reacted with 2,4-dinitrophenylhydrazine, 2,4-DNPH reagent, but has no reaction with Tollens' reagent.

(a) (i) Explain these observations.

.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....  
.....  
..... [4]

(ii) Draw the skeletal formula of **P** and give its name.

name of **P** ..... [2]

(iii) Draw the skeletal formula of **Q** and give its name.

name of **Q** ..... [2]



- (b) There are several structural isomers of **P** that also decolourise bromine, but only four of these structural isomers exhibit geometrical (cis-trans) isomerism.

Give the structures of any **three** structural isomers of **P** that exhibit geometrical (cis-trans) isomerism.

[3]

[Total: 11]





**BLANK PAGE**

---

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included the publisher will be pleased to make amends at the earliest possible opportunity.

Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.