

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**  
GCE Advanced Subsidiary Level and GCE Advanced Level

## **MARK SCHEME for the May/June 2014 series**

### **9701 CHEMISTRY**

**9701/22**

Paper 2 (AS Structured Questions), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Cambridge is publishing the mark schemes for the May/June 2014 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

Page 2	Mark Scheme	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2014	9701	22

Question	Answers	Mark	Total
1 (a)	The (total) number of protons and neutrons (in the nucleus of an atom)	1	1
(b) (i)	<p>Mass of an atom(s) or isotope</p> <p>relative to <math>\frac{1}{12}</math> (the mass) of (an atom of) carbon-12</p> <p><b>OR</b></p> <p>relative to carbon-12 which is (exactly) 12 (units)</p> <p>allow a correct expression</p>	1 1	2
(ii)	<p><math>^{79}\text{Br}</math>                  <math>^{81}\text{Br}</math></p> <p><math>78.92x</math>                  <math>80.92(100-x)</math> where <math>x = \% \text{ abundance of } ^{79}\text{Br}</math></p> <p>so <math>\frac{78.92x + 80.92(100-x)}{100} = 79.9</math></p> <p><math>x = 51</math></p> <p>hence <math>^{79}\text{Br} : ^{81}\text{Br} = 51 : 49</math></p>	1 1 1	3

Page 3	Mark Scheme	Syllabus	Paper
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Question	Answers	Mark	Total
(c)	$\frac{A}{A_r} \quad \frac{Br}{A_r} = 1:3$ $\frac{4.31}{A_r} \quad \frac{95.69}{79.9} = 1:3$ <p>So <math>\frac{95.69/79.9}{4.31/A_r} = 3</math></p> $A_r = \frac{3 \times 4.31 \times 79.9}{95.69} = 10.796 = 10.8 \text{ to 3 s.f.}$ <p>3 sig figs</p> <p>allow alternative correct methods</p>	1 1 1	3
(d) (i)	<p>Mg: bright/white light/flame <b>OR</b> white solid/smoke</p> $Mg + \frac{1}{2} O_2 \rightarrow MgO$ <p>allow correct multiples</p> <p>S: blue flame <b>OR</b> white/steamy fumes <b>OR</b> yellow solid disappears</p> $S + O_2 \rightarrow SO_2$ <p>allow correct multiples</p>	1 1 1 1	4
(ii)	$Al_2O_3 + 2NaOH + 7H_2O \rightarrow 2NaAl(OH)_4(H_2O)_2 \text{ OR}$ $Al_2O_3 + 2NaOH + 3H_2O \rightarrow 2NaAl(OH)_4 \text{ OR}$ $Al_2O_3 + 2NaOH \rightarrow 2NaAlO_2 + H_2O \text{ OR}$ $Al_2O_3 + 2OH^- + 7H_2O \rightarrow 2[Al(OH)_4(H_2O)_2]^- \text{ OR}$ $Al_2O_3 + 2OH^- + 3H_2O \rightarrow 2[Al(OH)_4]^- \text{ OR}$ $Al_2O_3 + 2OH^- \rightarrow 2AlO_2^- + H_2O$ $Al_2O_3 + 6HCl \rightarrow 2AlCl_3 + 3H_2O$ <p>allow correct ionic equations</p>	1 1	2

Page 4	Mark Scheme	Syllabus	Paper
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Question	Answers	Mark	Total
(e)	shape of $PCl_5$ = (trigonal) bipyramid(al)	1	2
	bond angles in $PCl_5$ = $120^\circ$ and $90^\circ$	1	
			<b>17</b>
<b>2 (a) (i)</b>	(The $C_2O_4^{2-}$ ions) lose electrons owtte/ora	1	1
(ii)	$2MnO_4^-(aq) + 5C_2O_4^{2-}(aq) + 16H^+(aq) \rightarrow 2Mn^{2+}(aq) + 10CO_2(aq) + 8H_2O(l)$	1+1+1	3
<b>(b) (i)</b>	$\frac{20.0 \times 0.100}{1000} = 2(.00) \times 10^{-3}$ (mol)	1	1
(ii)	$MnO_4^- : C_2O_4^{2-} = 2 : 5$ so amount of $C_2O_4^{2-} = (5/2) \times 2.00 \times 10^{-3} = 5(.00) \times 10^{-3}$ (mol) ecf from <b>(b)(i)</b>	1	1
(iii)	$5.00 \times 10^{-3} \times 250/25 = 0.05(0)$ (mol) ecf from <b>(b)(ii)</b>	1	1
(iv)	amount = mass / $M_r$ so $M_r = \text{mass} / \text{amount} = 6.30 / 0.05 = 126$ ecf from <b>(b)(iii)</b>	1	1
(v)	$126 - 90 = 36$ $36 / 18 = 2.00$ $x = 2$ Ecf from <b>(b)(iv)</b> if suitable	1	1
			<b>9</b>

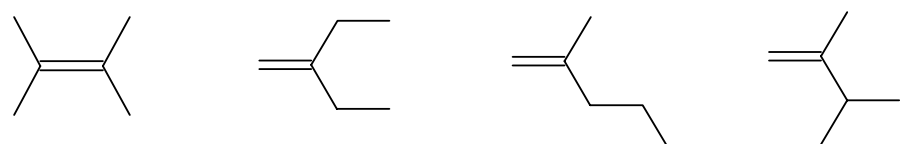
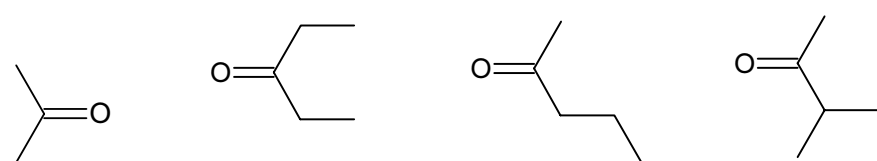
Page 5	Mark Scheme	Syllabus	Paper
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Question	Answers	Mark	Total
3 (a) (i)	metallic bonding	1	3
	strength of attraction/metallic bonding increases (Na–Al)/more energy is needed to break 'bonds'	1	
	due to increasing cation charge/charge density/increasing number of delocalised electrons/decreasing ionic radius	1	
(ii)	van der Waals' (forces)	1	3
	are greatest/more in sulfur/relative magnitude of forces $S > P > Cl > Ar$	1	
	because sulfur has the greatest number of electrons/as no. of electrons (in the molecules) decreases	1	
(iii)	<u>Covalent</u> bond(s) broken <b>OR</b> (Si has a) giant covalent (structure)	1	1
(b) (i)	Nuclear charge (in Ar) greater (than Cl) <b>AND</b> same shielding owtte	1	1
(ii)	p subshell/orbital in Al at higher energy (than s subshell in Mg) <b>ora</b> <b>OR</b> p subshell/orbital more shielded <b>ora</b>	1	1
(iii)	repulsion due to electron pair (in same/p orbital)	1	1
			<b>10</b>

Page 6	Mark Scheme	Syllabus	Paper
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Question	Answers	Mark	Total
4 (a)	physical: <u>fractional</u> distillation/fractionation	1	2
	chemical: crack(ing) (allow: reforming, isomerisation, thermal decomposition, desulfurisation)	1	
(b) (i)	Strong (C–C and C–H) bonds / high bond energies	1	2
	Non-polar / C and H have similar electronegativities	1	
(ii)	109.5° <b>AND</b> 120° (117° – 122°)	1	1
(iii)	ethane = tetrahedral ethene = trigonal planar	1	1
(iv)	4 × $\sigma$ /single bonds on Cs in ethane <b>AND</b> 3 × $\sigma$ and 1 × $\pi$ on Cs in ethene <b>OR</b> 2 × single and 1 × double on Cs in ethene allow from suitable labelled diagram	1	1
(c) (i)	$Cl_2 \rightarrow 2Cl\cdot$	1	5
	$C_2H_6 + Cl\cdot \rightarrow \cdot C_2H_5 + HCl$	1	
	$\cdot C_2H_5 + Cl_2 \rightarrow C_2H_5Cl + Cl\cdot$	1	
	$\cdot C_2H_5 + Cl\cdot \rightarrow C_2H_5Cl$	1	
	correct alternative terminations allowed		
	initiation, propagation, termination (correctly assigned)	1	
(ii)	$\cdot C_2H_5 + \cdot C_2H_5 \rightarrow C_4H_{10}$	1	1
			<b>13</b>

Page 7	Mark Scheme	Syllabus	Paper
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Question	Answers	Mark	Total
5 (a) (i)	<p>decolourisation of bromine: P is an alkene / contains C=C / is unsaturated</p> <p>hot conc. manganate(VII): breaks C=C <b>OR</b> single product implies P is symmetrical <b>OR</b> single organic product implies terminal =CH<sub>2</sub> as methanal is oxidised to CO<sub>2</sub></p> <p>2,4-DNPH confirms &gt;C=O / carbonyl / ketone in Q</p> <p>no reaction with Tollens': Q is not an aldehyde / is a ketone</p>	1 1 1 1	4
(ii)	 <p>(2,3-)dimethylbut-2-ene, 2-ethylbut(-1-)ene, 2-methylpent-1-ene, (2,3-)dimethylbut(-1-)ene</p>	1 1	2
(iii)	 <p>propan(-2-)one / acetone, pentan-3-one, pentan-2-one, 3-methylbutan(-2-)one ecf possible on <b>(a)(ii)</b></p>	1 1	2
(b)	<p>any 3 of</p> <p>CH<sub>3</sub>CH=C(CH<sub>3</sub>)C<sub>2</sub>H<sub>5</sub>            CH<sub>3</sub>CH=CHCH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub>            CH<sub>3</sub>CH=CHCH(CH<sub>3</sub>)<sub>2</sub>            C<sub>2</sub>H<sub>5</sub>CH=CHC<sub>2</sub>H<sub>5</sub></p>	1 1 1 1	max 3
			11