



# Cambridge International AS & A Level

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**CHEMISTRY****9701/11**

Paper 1 Multiple Choice

**October/November 2020****1 hour**

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)  
Data booklet

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**INSTRUCTIONS**

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A, B, C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

**INFORMATION**

- The total mark for this paper is 40.
- Each correct answer will score one mark. A mark will not be deducted for a wrong answer.
- Any rough working should be done on this question paper.

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This document has **16** pages. Blank pages are indicated.



## Section A

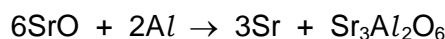
For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider to be correct.

Use of the Data Booklet may be appropriate for some questions.

1 Which statement is correct?

- A  $Cl$  has a relative isotopic mass of 35.5.
- B  $Cl_2$  has a relative molecular mass of 70.
- C  $ICl$  has a relative molecular mass of 162.4.
- D  $NaCl$  has a relative molecular mass of 58.5.

2 Strontium metal can be extracted from strontium oxide,  $SrO$ , by reduction with aluminium. One of the possible reactions is shown.



What is the maximum mass of strontium metal that can be produced from the reduction of 100 g of strontium oxide using this reaction?

- A 41.3 g      B 42.3 g      C 84.6 g      D 169.2 g

3 A single  $^{32}P$  nucleus can be produced when a single  $^{32}S$  nucleus joins with particle X. In the process a proton is emitted.

What is particle X?

- A a deuteron,  $^2_1H^+$
- B an electron
- C a neutron
- D a proton

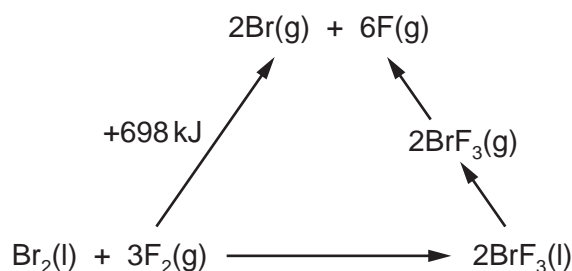
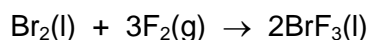
4 In which of the following, when in liquid form, are there only intermolecular forces based on temporary dipoles between the particles?

- A bromine
- B ethanol
- C hydrogen chloride
- D water

- 5 Copper has a high melting point.

What is the reason for the high melting point of copper?

- A strong attractive forces between copper atoms only  
 B strong attractive forces between copper ions and delocalised electrons  
 C strong attractive forces between copper ions only  
 D strong attractive forces between copper atoms and delocalised electrons
- 6 Which pair of standard enthalpy changes are numerically equal?
- A atomisation of  $\text{CH}_4(\text{g})$  and formation of  $\text{CH}_4(\text{g})$   
 B combustion of  $\text{CH}_3\text{OH}(\text{l})$  and combustion of graphite + 2(combustion of  $\text{H}_2(\text{g})$ )  
 C combustion of graphite and formation of  $\text{CO}_2(\text{g})$   
 D neutralisation of  $\text{HCl}(\text{aq})$  with  $\text{NaOH}(\text{aq})$  and formation of  $\text{H}_2\text{O}(\text{l})$
- 7 An energy cycle is drawn for the following reaction.



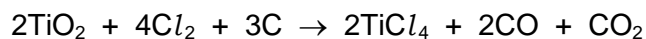
The standard enthalpy of formation of  $\text{BrF}_3(\text{l}) = -301 \text{ kJ mol}^{-1}$ .

The enthalpy change of  $\text{BrF}_3(\text{l})$  to  $\text{BrF}_3(\text{g})$  is  $+44 \text{ kJ mol}^{-1}$ .

What is the average bond energy of the Br–F bond in  $\text{BrF}_3$ ?

- A  $152 \text{ kJ mol}^{-1}$     B  $202 \text{ kJ mol}^{-1}$     C  $304 \text{ kJ mol}^{-1}$     D  $404 \text{ kJ mol}^{-1}$
- 8 In which reaction does the greatest change in the oxidation number of sulfur occur?
- A  $\text{S}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{SO}_2(\text{g})$   
 B  $\text{SO}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{SO}_3(\text{g})$   
 C  $\text{SO}_3(\text{g}) + \text{H}_2\text{SO}_4(\text{l}) \rightarrow \text{H}_2\text{S}_2\text{O}_7(\text{l})$   
 D  $\text{H}_2\text{S}_2\text{O}_7(\text{l}) + \text{H}_2\text{O}(\text{l}) \rightarrow 2\text{H}_2\text{SO}_4(\text{l})$

- 9 The first stage in the chloride process for the manufacture of titanium consists of the following reaction.



What is reduced in this reaction?

- A carbon
  - B chlorine
  - C oxygen
  - D titanium
- 10 In aqueous solution, sulfuric acid dissociates as shown.

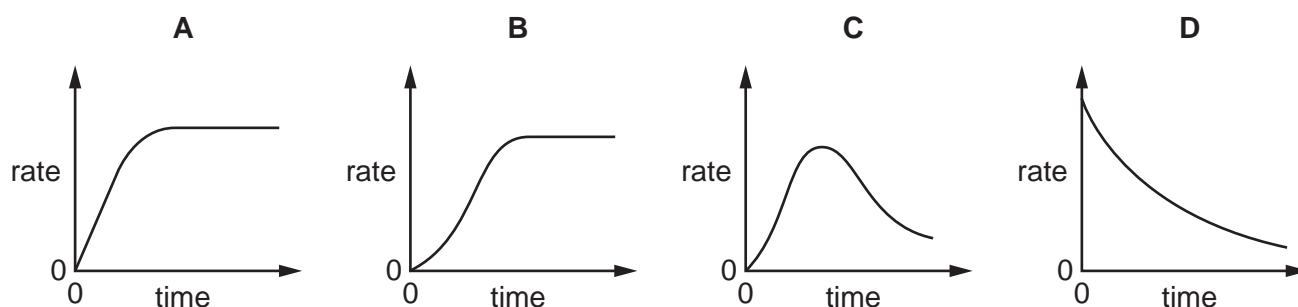


Analysis of a  $2.00 \text{ mol dm}^{-3}$  solution of  $\text{H}_2\text{SO}_4$  found the  $\text{HSO}_4^-$  concentration to be  $1.988 \text{ mol dm}^{-3}$ .

What is  $K_c$ ?

- A  $1.381 \times 10^5 \text{ dm}^3 \text{ mol}^{-1}$
  - B  $82.34 \text{ dm}^3 \text{ mol}^{-1}$
  - C  $1.214 \times 10^{-2} \text{ mol dm}^{-3}$
  - D  $7.244 \times 10^{-5} \text{ mol dm}^{-3}$
- 11 An autocatalytic reaction is a reaction in which one of the products catalyses the reaction.

Which curve would be obtained if the rate of an autocatalytic reaction is plotted against time?



12 X and Y are two elements in Period 3 of the Periodic Table. They combine to form compound Z.

X forms a soluble acidic oxide. The oxidation number of X in this oxide is +4.

Y forms an amphoteric oxide.

What is the formula of compound Z?

- A  $AlP$                       B  $Al_2S_3$                       C  $Si_2P_5$                       D  $SiS_2$

13 This question is about two elements in Group 2, Q and R.

Three of the statements shown are correct for metal Q.

The one remaining statement is correct for metal R.

Which statement applies to R?

- A A saturated solution of the hydroxide of this metal has the higher pH value.  
B This metal has a carbonate that is used in agriculture to reduce the acidity of soil.  
C This metal has the greater atomic radius.  
D This metal reacts more quickly with cold water.

14 The electronic arrangement for atoms of four elements is given.

Which element is the strongest oxidising agent?

- A  $1s^2 2s^2 2p^5$   
B  $1s^2 2s^2 2p^6 3s^2$   
C  $1s^2 2s^2 2p^6 3s^2 3p^5$   
D  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

15 A student mixes pairs of chemicals together in separate test-tubes.

- excess calcium (s) + water (l)
- barium chloride (aq) + strontium hydroxide (aq)
- calcium carbonate (s) + excess hydrochloric acid (aq)
- magnesium sulfate (aq) + barium nitrate (aq)

How many of the mixtures produce a white, solid product?

- A 0                      B 1                      C 2                      D 3

16 With which compound does concentrated sulfuric acid react **both** as a strong acid **and** as an oxidising agent?

- A magnesium carbonate
- B potassium chloride
- C sodium bromide
- D sulfur trioxide

17 Ammonia can undergo an acid–base reaction with hydrogen chloride to form ammonium chloride.

Which statement is correct?

- A The ammonium ion is basic.
- B The hydrogen atom from  $\text{HCl}$  donates a lone pair of electrons to the nitrogen atom.
- C The  $\text{H–N–H}$  bond angle in ammonia is the same as the  $\text{H–N–H}$  bond angle in the ammonium ion.
- D The  $\text{H–N–H}$  bond angle in the ammonium ion is the same as the  $\text{H–C–H}$  bond angle in methane.

18 What are the trends in the stated properties as Group 2 is descended from magnesium to barium?

|   | decomposition temperature of the carbonate | first ionisation energy |
|---|--|-------------------------|
| A | decreases                                  | increases               |
| B | decreases                                  | decreases               |
| C | increases                                  | increases               |
| D | increases                                  | decreases               |

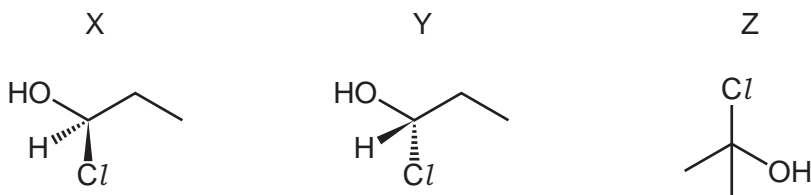
19 Sulfur dioxide,  $\text{SO}_2$ , reacts with calcium hydroxide in aqueous solution.

What is the main product that is first formed?

- A  $\text{Ca}(\text{HSO}_4)_2$     B  $\text{CaS}$     C  $\text{CaSO}_3$     D  $\text{CaSO}_4$

- 20 Structural and stereoisomerism should be considered when answering this question.

Compounds X, Y and Z are shown.



How many other isomers of  $C_3H_7ClO$  are there that are alcohols?

- A** 2                      **B** 3                      **C** 4                      **D** 5
- 21 Two students each make a statement about 2-methylbut-1-ene.
- Student 1 states that 2-methylbut-1-ene has geometrical isomers.
- Student 2 states that 2-methylbut-1-ene reacts with HBr in an addition reaction to give 1-bromo-2-methylbutane as the main product.
- Which students are correct?
- A** both 1 and 2  
**B** 1 only  
**C** 2 only  
**D** neither 1 nor 2
- 22 Which statement is correct when referring to the complete combustion of PVC?
- A** A gas is made which contributes to global warming.  
**B** Carbon dioxide and water are the only products.  
**C** If water is used to clean the exhaust gases, the water becomes alkaline.  
**D** There is no need to treat the exhaust gases as the products are non-hazardous.
- 23 Iodoethane,  $CH_3CH_2I$ , reacts with aqueous silver nitrate at  $50^\circ C$ . A precipitate forms during this reaction.

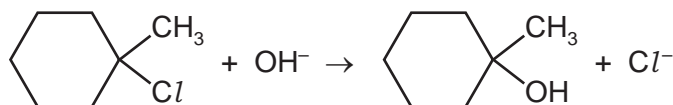
Which row of the table is correct about this reaction?

|          | type of organic reaction   | colour of precipitate |
|----------|----------------------------|-----------------------|
| <b>A</b> | electrophilic substitution | cream                 |
| <b>B</b> | electrophilic substitution | yellow                |
| <b>C</b> | nucleophilic substitution  | cream                 |
| <b>D</b> | nucleophilic substitution  | yellow                |

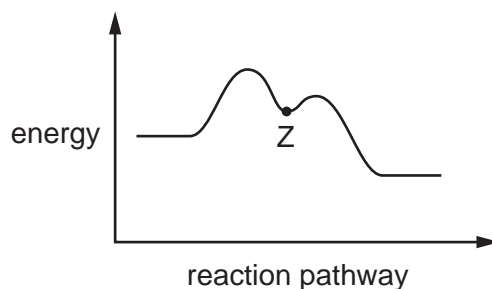
- 24 A student converts 1-iodopropane,  $C_3H_7I$ , into butanoic acid,  $C_3H_7CO_2H$ , by a two-stage chemical synthesis.

In the **first** of the two stages, which reagent is reacted with 1-iodopropane?

- A aqueous sodium hydroxide  
 B ethanolic ammonia  
 C ethanolic potassium cyanide  
 D ethanolic sodium hydroxide
- 25 1-chloro-1-methylcyclohexane is hydrolysed by heating with  $NaOH(aq)$ .



The reaction pathway is shown.



One carbon atom in 1-chloro-1-methylcyclohexane is bonded to three other carbon atoms.

What is the charge on this carbon atom at point Z?

- A  $1-$                       B  $\delta-$                       C  $\delta+$                       D  $1+$
- 26 An alcohol with the molecular formula  $C_5H_{12}O$  decolourises warm acidified potassium manganate(VII). The alcohol also gives a yellow precipitate with alkaline aqueous iodine.

What could be the identity of the alcohol?

- A 2-methylbutan-2-ol  
 B 3-methylbutan-2-ol  
 C pentan-1-ol  
 D pentan-3-ol

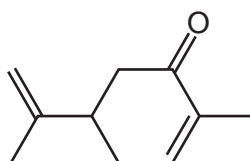


27 Which pair of test results would prove that a substance, X, is a ketone?

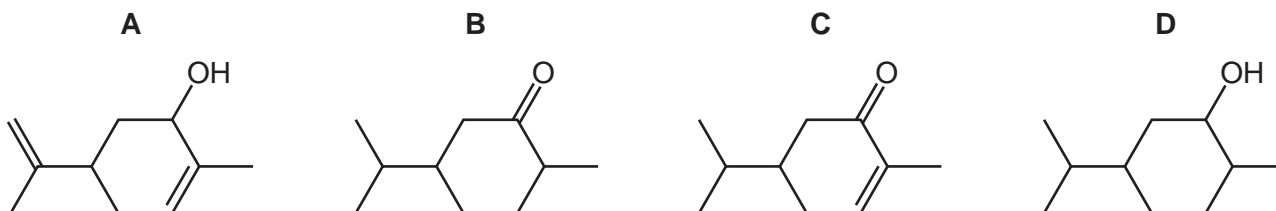
- A X has no reaction with Tollens' reagent. X reacts with alkaline aqueous iodine.
- B X is reduced by lithium aluminium hydride. X is oxidised by acidified dichromate(VI).
- C X reacts with 2,4-DNPH reagent. X has no reaction with Fehling's reagent.
- D X reacts with hydrogen cyanide. X is reduced by lithium aluminium hydride.

28 Carvone is found in spearmint oil.

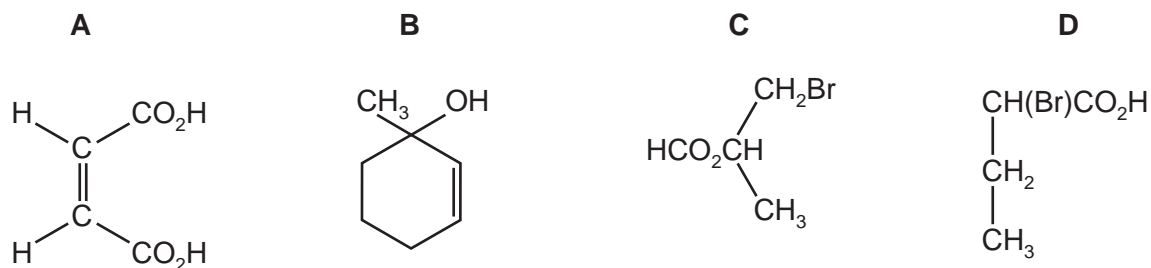
carvone



Which product is formed when carvone is reacted with  $\text{NaBH}_4$ ?

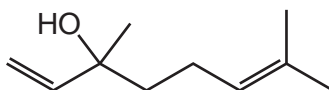


29 Which compound is chiral and reacts with  $\text{Na}_2\text{CO}_3$  to give  $\text{CO}_2$ ?



30 The skeletal formula of compound X is shown.

compound X



What is the molecular formula of compound X?

- A  $\text{C}_{10}\text{H}_{18}\text{O}$
- B  $\text{C}_{10}\text{H}_{20}\text{O}$
- C  $\text{C}_{11}\text{H}_{22}\text{O}$
- D  $\text{C}_{11}\text{H}_{24}\text{O}$

## Section B

For each of the questions in this section, one or more of the three numbered statements 1 to 3 may be correct.

Decide whether each of the statements is or is not correct (you may find it helpful to put a tick against the statements that you consider to be correct).

The responses **A** to **D** should be selected on the basis of

| <b>A</b>               | <b>B</b>                 | <b>C</b>                 | <b>D</b>          |
|------------------------|--------------------------|--------------------------|-------------------|
| 1, 2 and 3 are correct | 1 and 2 only are correct | 2 and 3 only are correct | 1 only is correct |

No other combination of statements is used as a correct response.

Use of the Data Booklet may be appropriate for some questions.

**31** Nitrogen forms a number of oxides. Their enthalpies of formation are given.

$$\Delta H_f^\ominus[\text{NO}(\text{g})] = +90 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\ominus[\text{N}_2\text{O}(\text{g})] = +82 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\ominus[\text{NO}_2(\text{g})] = +33 \text{ kJ mol}^{-1}$$

Which statements are correct?

- 1 If  $\text{N}_2\text{O}(\text{g})$  is oxidised by  $\text{O}_2(\text{g})$  to  $\text{NO}_2(\text{g})$ , 16 kJ is released per mole of  $\text{N}_2\text{O}$ .
- 2 The decomposition of  $\text{N}_2\text{O}(\text{g})$  to  $\text{N}_2(\text{g})$  and  $\text{O}_2(\text{g})$  is exothermic.
- 3 The reaction between NO and oxygen is exothermic.

**32** Which statements are correct?

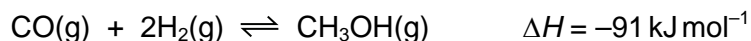
- 1 enthalpy of combustion of  $\text{H}_2$  = enthalpy of formation of  $\text{H}_2\text{O}$
- 2 enthalpy of formation of  $\text{H}_2$  =  $-(\text{enthalpy of atomisation of } \text{H}_2)$
- 3 enthalpy of solution of  $\text{HCl}$  = enthalpy of hydration of  $\text{H}^+$  + enthalpy of hydration of  $\text{Cl}^-$

**33** The units of  $K_c$  for an equilibrium reaction are  $\text{mol}^{-1} \text{ dm}^3$ .

What could be the equation for the equilibrium?

- 1  $\text{A}(\text{aq}) + \text{B}(\text{aq}) \rightleftharpoons \text{C}(\text{s}) + \text{D}(\text{aq})$
- 2  $\text{P}(\text{aq}) + \text{Q}(\text{aq}) \rightleftharpoons \text{R}(\text{aq})$
- 3  $\text{W}(\text{aq}) + 2\text{X}(\text{aq}) \rightleftharpoons \text{Y}(\text{aq}) + \text{Z}(\text{aq})$

34 Methanol,  $\text{CH}_3\text{OH}$ , can be produced industrially by reacting  $\text{CO}$  with  $\text{H}_2$ .



The process can be carried out at  $4 \times 10^3 \text{ kPa}$  and  $1150 \text{ K}$ .

Which statements about this reaction are correct?

- 1 Increasing the temperature will increase the rate of reaction because more effective collisions will occur.
- 2 Lowering the temperature will reduce the rate of reaction because the forward reaction is exothermic.
- 3 Increasing the pressure will reduce the rate of reaction because there are a larger number of moles on the left-hand side of the equation.

35 Which rows correctly show the relative electrical conductivities of the sets of three Period 3 elements?

|   | greatest conductivity | →         | least conductivity |
|---|-----------------------|-----------|--------------------|
| 1 | sodium                | silicon   | chlorine           |
| 2 | aluminium             | magnesium | phosphorus         |
| 3 | sulfur                | silicon   | phosphorus         |

36 Three test-tubes, X, Y and Z, each contain water.

- A small amount of  $\text{NaCl}$  is added to test-tube X.
- A small amount of  $\text{SiCl}_4$  is added to test-tube Y.
- A small amount of  $\text{AlCl}_3$  is added to test-tube Z.

After a short time, two drops of universal indicator solution are added to each test-tube.

Which statements can be correct?

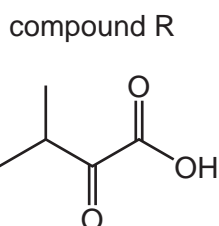
- 1 The pH in test-tube X is 7.
- 2 The pH in test-tube Y is 2.
- 3 The pH in test-tube Z is 2.

The responses **A** to **D** should be selected on the basis of

| <b>A</b>               | <b>B</b>                 | <b>C</b>                 | <b>D</b>          |
|------------------------|--------------------------|--------------------------|-------------------|
| 1, 2 and 3 are correct | 1 and 2 only are correct | 2 and 3 only are correct | 1 only is correct |

No other combination of statements is used as a correct response.

**37** The structure of compound R is shown.



Which statements about compound R are correct?

- 1 It has an  $M_r$  of 116.
  - 2 It contains two groups that show strong absorptions between  $1640$  and  $1740\text{ cm}^{-1}$  in its infrared spectrum.
  - 3 Its only infrared absorption between  $2500$  and  $3000\text{ cm}^{-1}$  is sharp and strong.
- 38** During the bromination of methane, the free radical  $\text{CH}_3\bullet$  is generated. A possible termination step of this reaction is the formation of  $\text{C}_2\text{H}_6$  by the combination of two free radicals.

What could be produced in a termination step during the bromination of **propane**?

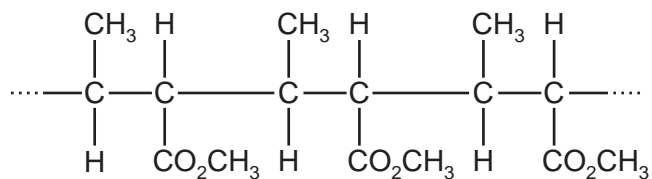
- 1  $\text{CH}_3\text{CH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_3$
- 2  $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}(\text{CH}_3)_2$
- 3  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{CH}_3)_2$

39 Three reactions of primary alcohols are listed.

Which reactions give water as one of the products?

- 1 reaction with ethanoic acid
- 2 reaction with concentrated HBr
- 3 passing the alcohol vapour over heated  $Al_2O_3$

40 The diagram shows part of the structure of polymer X.



Which reagents react with polymer X?

- 1 aqueous sulfuric acid
- 2 aqueous sodium hydroxide
- 3 sodium

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