

# OCR (B) Chemistry A-Level

## DF2 - Bonding and structure

### Flashcards

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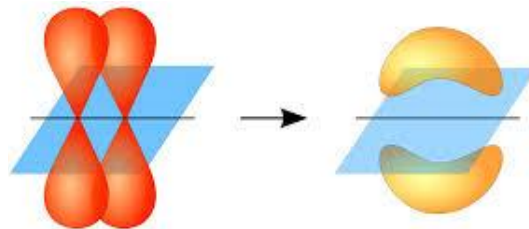


What is the difference between a  $\pi$  bond  
and a  $\sigma$  bond?



What is the difference between a  $\pi$  bond and a  $\sigma$  bond?

A  $\sigma$  bond is caused by the direct overlap of the p-orbitals of carbon atoms,  $\pi$  bonds are caused by the indirect 'leaning' overlap of p-orbitals.



# What is a double bond?



# What is a double bond?

A double bond consists of one  $\pi$  bond and one  $\sigma$  bond.



# How can you draw a molecule in 3D?



## How can you draw a molecule in 3D?

- A solid line represents a bond in the plane of the paper.
- A solid wedge represents a bond coming out of the plane of the paper.
- A dashed/dotted wedge represents a bond going into the plane of the paper.

