

CAIE Chemistry A-level

6: Electrochemistry Definitions

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Definitions and Concepts for CAIE Chemistry A-level Electrochemistry

Disproportionation: A reaction in which a substance is simultaneously reduced and oxidised. Chlorine undergoes disproportionation in the reaction with cold, dilute, aqueous sodium hydroxide.

Oxidation: The loss of electron(s) which leads to an increase in oxidation number.

Oxidation number: The charge of an ion or a theoretical charge of an atom in a covalently bonded compound assuming the bond becomes ionic.

Oxidising agent: A species which brings about oxidation by gaining electrons. The oxidising agent is itself reduced.

Redox reaction: A reaction in which both reduction and oxidation are occurring simultaneously.

Reducing agent: A species which brings about reduction by losing electrons. The reducing agent is itself oxidised.

Reduction: The gain of electron(s) which leads to a decrease in oxidation number.

