

The block indicates the type of orbital which the outer shell electron is held in

Periodic table is split into s, p, d and f blocks

Elements are arranged in the periodic table by increasing atomic number

2.1 PERIODICITY: CLASSIFICATION

Elements with same number of electron shells

Arranged in same period (rows)

Elements with the same number of valence electrons (except Group 0)

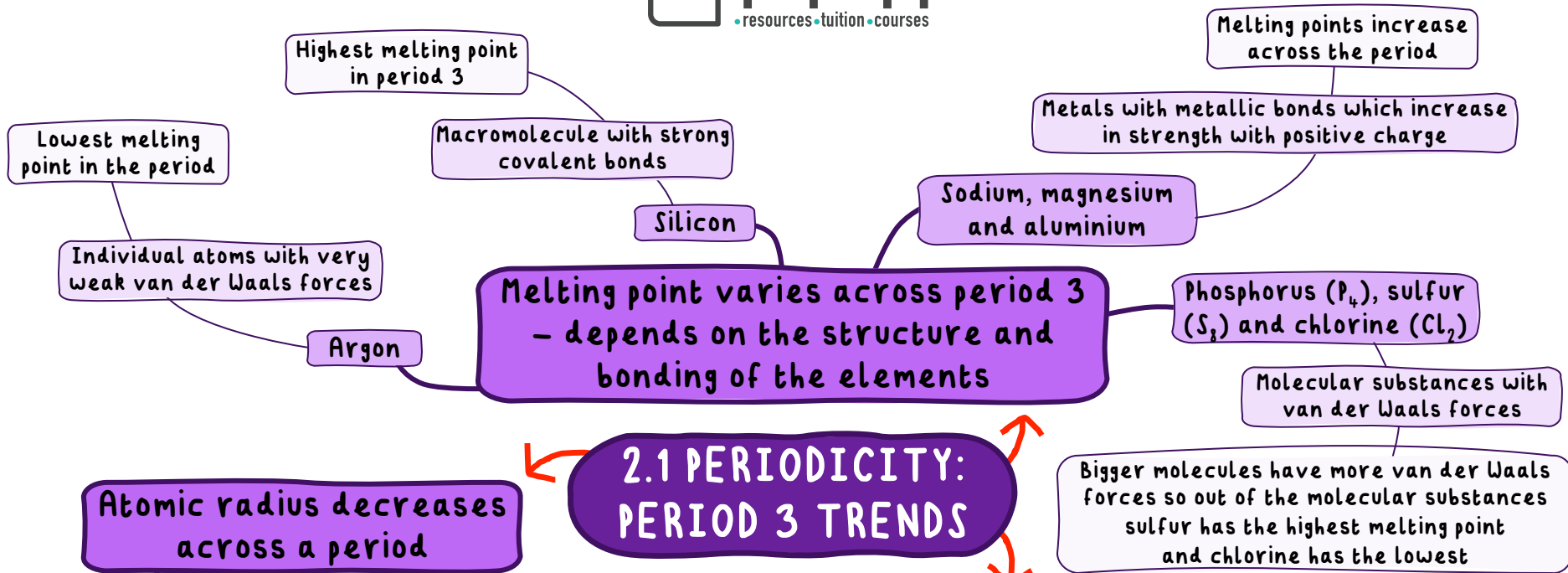
Similar chemical properties

Arranged in same group (column)

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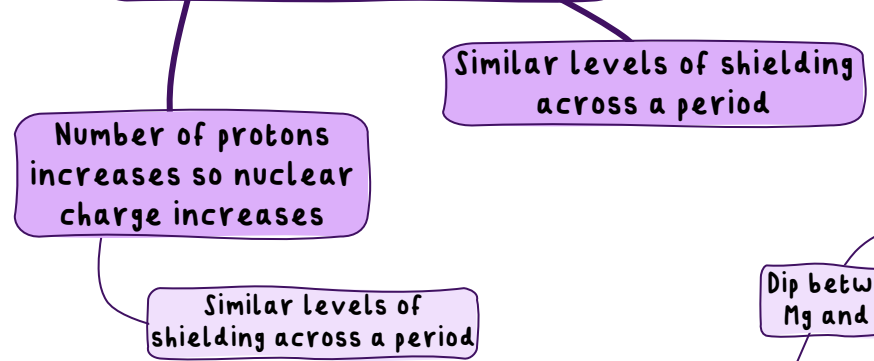


**Melting point varies across period 3
- depends on the structure and bonding of the elements**

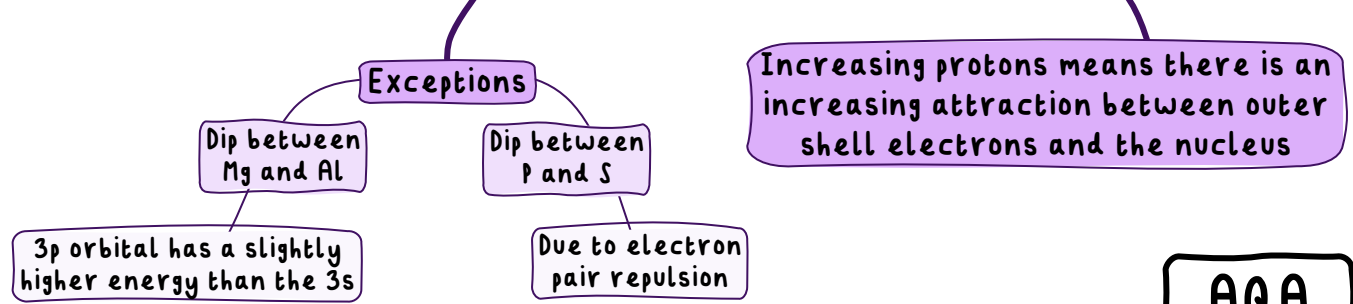


**2.1 PERIODICITY:
PERIOD 3 TRENDS**

Atomic radius decreases across a period



Ionisation energy generally increases across the period



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