

A Level Physics A

H556/02 Exploring physics

Question Set 20

1(a) The structure of atoms was deduced in the early 1900s by Rutherford and his co-workers from the scattering of alpha-particles by a very thin sheet of gold.

Rutherford assumed that the scattering of the alpha-particles was due to electrostatic forces. Fig. 23 shows a detector used to record the number N of alpha-particles scattered through an angle θ .

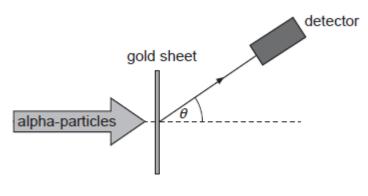


Fig. 23

At $\theta = 0^{\circ}$, N was too large to be measured. The table below summarises some of the collected data.

O I°	Ig (<i>N</i>)
150	1.5
75	2.3
60	2.7
30	3.9
15	5.1
0	N too large

(i) Show that the number of alpha-particles scattered through 15° is about 4000 times more than those scattered through 150°.

$$\theta = (5^{\circ} \Rightarrow) \log N = 5 \cdot (5) \Rightarrow N = (0^{\circ}) = 3081 = 4000$$
 [1]

[3]

(ii) Use the evidence from the table to explain the structure of the atom.

- (b) A proton with kinetic energy 0.52MeV is travelling directly towards a stationary nucleus of cobalt-59 (59 Co) in a head-on collision.
 - (i) Explain what happens to the electric potential energy of the proton-nucleus system.

 [1]
 - (ii) Calculate the **minimum** distance R between the proton and cobalt nucleus.

KE is transferred to electric potential energy:
$$E = \frac{Qq}{4\pi\epsilon_0}R$$

Q of colonia = $27 \times 1.6 \times 10^{-19}$

O of prolon = 1.6×10^{-19}
 $0.52 \times 1 \times 10^6 \times 1.6 \times 10^{-19} = \frac{27 \times \left(1.6 \times 10^{-19}\right)^2}{4\pi \times 8.85 \times 10^{-12} \times R}$

R = $\frac{7.5 \times 10^{-14}}{1.6 \times 10^{-19}}$ m [3]

Total Marks for Question Set 20: 8



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