

GCSE Chemistry B (Twenty First Century Science)
J258/01 Breadth in Chemistry (Foundation Tier)

Question Set 30

Multiple Choice Questions

1

Lithium-ion batteries are used in phones, tablets and electric cars.

(a) Lithium reacts with chlorine and with bromine.

(i) Lithium is in Group 1. Chlorine and bromine are in Group 17.

Draw lines to connect each element with **one** correct property.

Element	Property
Lithium (Group 1)	Conducts electricity
Chlorine (Group 17)	Unreactive
	Colourless gas
	Green gas

[2]

(ii) 14 g of lithium reacts with 71 g of chlorine.

What mass of chlorine reacts with 5.6 g of lithium?

Mass of chlorine =g [2]

(iii) Jack reacts lithium with chlorine. He then reacts lithium with bromine.

Describe how the rates of these two reactions are different.

[1]

(b) Lithium is made by the electrolysis of molten lithium chloride.

Which substance is formed at each electrode?

Put a ring around each correct answer.

Anode (positive electrode): **chloride** **chlorine** **hydride** **hydrogen**

Cathode (negative electrode): **oxide** **oxygen** **lithium**

[2]

(c) Lithium-ion batteries contain chemical cells.

Which statement is the correct definition for a chemical cell?

Tick (✓) **one** box.

A chemical cell produces its full potential difference but the potential difference then quickly decreases.

A chemical cell takes a long time to get to its full potential difference.

A chemical cell produces a potential difference that lasts for a short time.

A chemical cell produces a potential difference until the reactants are used up.

[1]

Total Marks for Question Set 30: 8

Resource Materials

Question Set No: 30

The Periodic Table of the Elements

(1)	(2)											(3)	(4)	(5)	(6)	(7)	(0)	
1 H hydrogen 1.0																		2 He helium 4.0
3 Li lithium 6.9	4 Be beryllium 9.0											5 B boron 10.8	6 C carbon 12.0	7 N nitrogen 14.0	8 O oxygen 16.0	9 F fluorine 19.0	10 Ne neon 20.2	
11 Na sodium 23.0	12 Mg magnesium 24.3											13 Al aluminium 27.0	14 Si silicon 28.1	15 P phosphorus 31.0	16 S sulfur 32.1	17 Cl chlorine 35.5	18 Ar argon 39.9	
19 K potassium 39.1	20 Ca calcium 40.1	21 Sc scandium 45.0	22 Ti titanium 47.9	23 V vanadium 50.9	24 Cr chromium 52.0	25 Mn manganese 54.9	26 Fe iron 55.8	27 Co cobalt 58.9	28 Ni nickel 58.7	29 Cu copper 63.5	30 Zn zinc 65.4	31 Ga gallium 69.7	32 Ge germanium 72.6	33 As arsenic 74.9	34 Se selenium 79.0	35 Br bromine 79.9	36 Kr krypton 83.8	
37 Rb rubidium 85.5	38 Sr strontium 87.6	39 Y yttrium 88.9	40 Zr zirconium 91.2	41 Nb niobium 92.9	42 Mo molybdenum 95.9	43 Tc technetium	44 Ru ruthenium 101.1	45 Rh rhodium 102.9	46 Pd palladium 106.4	47 Ag silver 107.9	48 Cd cadmium 112.4	49 In indium 114.8	50 Sn tin 118.7	51 Sb antimony 121.8	52 Te tellurium 127.6	53 I iodine 126.9	54 Xe xenon 131.3	
55 Cs caesium 132.9	56 Ba barium 137.3	57-71 lanthanoids	72 Hf hafnium 178.5	73 Ta tantalum 180.9	74 W tungsten 183.8	75 Re rhenium 186.2	76 Os osmium 190.2	77 Ir iridium 192.2	78 Pt platinum 195.1	79 Au gold 197.0	80 Hg mercury 200.6	81 Tl thallium 204.4	82 Pb lead 207.2	83 Bi bismuth 209.0	84 Po polonium	85 At astatine	86 Rn radon	
87 Fr francium	88 Ra radium	89-103 actinoids	104 Rf rutherfordium	105 Db dubnium	106 Sg seaborgium	107 Bh bohrium	108 Hs hassium	109 Mt meitnerium	110 Ds darmstadtium	111 Rg roentgenium	112 Cn copernicium		114 Fl flerovium		116 Lv livermorium			

Key atomic number Symbol name relative atomic mass
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