

GCSE Chemistry A (Gateway Science) J248/03 C1-C3 and C7 Higher (Higher Tier)

Question Set 31

1 A student adds calcium to dilute hydrochloric acid. The mixture begins to fizz.

Write a balanced symbol equation for this reaction.

 $Ca + 2HCI \longrightarrow CaCl_{2} + H_{2}$ Total Marks for Question Set 31: 2

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(0)	18 2 He ^{hellum} 4.0	10 Ne 20.2	18 Ar argon 39.9	36	Kr	krypton 83.8	54	Xenon	131.3	86	Rn	radon			
(2)	17	9 F fluorine 19.0	17 C1 chlorine 35.5	35	Br	bromine 79.9	53	I lodine	126.9	85	At	astatine			
(9)	16	8 O oxygen 16.0	16 Suffur 32.1	34	Se	selenium 79.0	52	Te tellurium	127.6	84	Ро	polonium	116	Ľ	livermorium
(5)	15	7 N nitrogen 14.0	15 Phosphorus 31.0	33	As	arsenic 74.9	51	Sb	121.8	83	Bi	bismuth 209.0			
(4)	14	6 C carbon 12.0	14 silicon 28.1	32	Ge	germanium 72.6	50	۳ ۳	118.7	82	Pb	lead 207.2	114	F۱	flerovium
(3)	13	5 Baran 10.8	13 A1 aluminium 27.0	31	Ga	gallium 69.7	49	In Indium	114.8	81	T1	thallium 204.4			
			12	30	Zn	zine 65.4	48	Cd	112.4	80	Hg	mercury 200.6	112	ы	copernicium
			5	29	cu	copper 63.5	47	Ag	107.9	79	Αu	^{gold} 197.0	111	Rg	roentgenium
			10	28	iN	nickel 58.7	46	Pd	106.4	78	ħ	platinum 195.1	110	Ds	darmsta dijum
თ						cobalt 58.9	45	Rh	102.9	77	Ir	iridium 192.2	109	Mt	meitnerium
			œ	26	Fe	lron 55.8	44	Ru	101.1	76	os	osmium 190.2	108	Hs	hassium
			2	25	Mn	manganese 54.9	1	-	-		Re	rhenium 186.2	107	旧	bohrium
	er mass		9	24	ບັ	chromium 52.0	42	Mo	95.9	74	8	tungsten 183.8	106	Sg	seaborgium
	Key atomic number Symbol ^{name} relative atomic mass		ις.	23	>	vanadium 50.9					Та	tantalum 180.9	105	Вb	dubnium
	ato relativ		4	22	Ħ	ttanium 47.9	40	Zr zireonium	91.2	72	Ħ	hafinium 178.5	104	Rf	rutherfordium
			67	21	Sc	scandium 45.0	39	th ium (88.9	i	57-71	lanthanoids	007 00	89-103	actinoids
(2)	7	4 Be ^{beryllium} 9.0	12 Mg ^{magneslum} 24.3	20	Ca	calcium 40.1	38	Sr	87.6	56	Ba	barium 137.3	88	Ra	radium
(1)	H 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1	3 Li ^{lithium} 6.9	11 sodium 23.0	19	¥	potassium 39.1	37	Rb	85.5	55	cs	caesium 132.9	87	F	francium

The Periodic Table of the Elements



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